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USING THIS BOOK

These books are intended as a user friendly resource for teachers to encourage “hands on science”. The experiments can easily be incorporated when developing new programs, responding to a new syllabus or to enrich current programs.

While handing photocopy experiments to students may seem too easy, there are advantages beyond time saved to teachers. I have found that writing up lengthy procedures merely fragments student topic notes and the time consumed in writing directions often means the activity is rushed with conclusions poorly addressed.

TRAINING

Students are often unused to following written practical instructions. It is wise to start with a simple experiment, read and demonstrate each procedure step, carefully explain what you expect for records of results and conclusion, then read out any warnings in the Risk Assessment on your Teacher Copy. Warnings do not appear on Student Copies since some parents might be unduly alarmed.

Make sure you collect and mark the completed Student Copies in the first instance and randomly thereafter. You will find the students adapt rapidly to this approach to practical work. You will be able to allow accelerated progression, different work stations and cooperative learning approaches in the laboratory.

EXPERIMENTS

The experiments are listed alphabetically by name to make them easy to find, however I draw your attention to the INDEX BY TOPIC at the end of the book. The topic index covers all volumes and lists experiment names under topics to which they are related. The purpose of the topic index is so you can quickly find experiments relating to a particular area of study. Rather than complicate the index by duplicating junior and senior topics, common topics appear only once with both junior and senior experiments appearing below.

EQUIPMENT

I have tried to include all the equipment needed in each experiment. Concentrations are given in percentages so you are not constantly stopping to calculate molarities.

The following guides might help:

- 1/ The equipment list is based on items required by one group.
- 2/ Any chemical listed with a concentration is a stock solution that must be prepared. In the case of concentrated acids with density and strength corrections the following applies;

| | |
|--|---------------|
| Hydrochloric acid, 370g/litre | 1Molar = 10% |
| Sulfuric Acid, 98%, 1.84g/ml density, | 1Molar = 5.4% |
| Nitric Acid, 70%, 1.42g/ml density, | 1Molar = 6.3% |
| Phosphoric acid, 85%, 1.69g/ml density | 1Molar = 6.8% |
| Ethanoic Acid (Glacial Acetic) 99%, | 1Molar = 6.6% |

3/ Any chemicals without a concentration means simply a class supply.

4/ Please read the risk assessment for your own protection during preparation and DISPOSAL.

5/ I recommend that you photocopy the Teacher Copies and place them in plastic sleeves in a ring folder.

Please feel free to write to me with any suggested improvements and any new experiments would be most welcome.

RISK ASSESSMENT

Every experiment has certain risks, not just from chemicals and equipment but from the unpredictable nature of students. In my years of teaching I have seen some remarkably stupid things such as a student attempting to “snort” citric acid or another trying a sucking contest with a vacuum cleaner. With this in mind my classification of risk is based on chemical toxicity and exposure (following the new lists), except where the “student factor” seems a greater hazard. Of course professional judgment is needed. Some junior classes can be trusted with delicate equipment while others cannot be trusted with a pair of scissors. However as a general guide:

Low Hazard - Junior Classes

Mild Hazard - Junior classes with close supervision.

Moderate Hazard - Senior classes

HAZARDOUS - Teacher demonstration only.

Remember, familiarity often breeds contempt. Chemicals that are used often may be more toxic than you realise. For example cobalt chloride is a suspected carcinogen with an LD50 of 80mg/kg and has been deleted from junior experiments in these books. By comparison, copper sulfate, a very commonly used laboratory chemical, has an LD50 of only 300mg/kg. Phenol has the same toxicity yet I am sure you are much more cautious of phenol than you are of copper sulfate. By contrast, lead nitrate is not overly toxic but is dangerous due to its accumulation from repeated small exposures.

PRACTICAL ASSESSMENT SUGGESTIONS

1/ A list of controlled experiments appears in the topic index. Ask your students to identify the appropriate control in each of these experiments.

2/ Collect student work sheets at random and apply a standard marking scale eg. records (4marks), observations (2marks), results (2marks), and conclusion (2marks). This should make the students take practical work seriously, encouraging participation, accurate records and a deductive conclusion (too often neglected).

3/ Record anecdotal marks as the students perform the experiment, focusing on equipment recognition, reading instructions and complete notes.

STUDENT: _____

201

A Scientists Eyes

Aim: To observe closely and see beyond the obvious.

Equipment

A house candle

Alfoil squares, 10X10cm, 3

Procedure

Late last century, Michael Faraday revolutionised the world by discovering how to make electricity using magnets. His success was owed much to careful observation. He once made 53 observations of a candle flame.

1/ Light a candle and fix it upright on the square of alfoil with some melted wax.

2/ Closely observe the flame. On a page in your book record what you notice about the shape, location, structure, colour and behaviour of the flame.

3/ Closely observe the wick. Record what you notice about its colour, position and rate of burning.

4/ Closely observe the wax. Record what you notice about the state and flow of the wax.

5/ Closely observe the air above the candle flame. Record what you notice about the movement and smoke above the flame. Hold a peice of alfoil 10cm above the flame and record any changes.

6/ Place an alfoil tube around the candle flame and note any change.

- Predict what would happen if the candle were narrow like a birthday candle.
- Predict what would happen if the candle was very broad.
- Predict what might happen when a broad candle is moved.
- Predict what would happen if the wick ran up one side instead of through the centre.
- Given the same amount of wax, which would burn longest, a very tall and thin candle or a short and fat candle?
- What is the advantage of a candle of medium thickness?

Results: _____

Conclusion: _____

A Scientists Eyes

Topics: Scientific Method

Aim: To observe closely and see beyond the obvious.

Equipment

A house candle

Alfoil squares, 10X10cm, 3

Procedure

Late last century, Michael Faraday revolutionised the world by discovering how to make electricity using magnets. His success was owed much to careful observation. He once made 53 observations of a candle flame.

1/ Light a candle and fix it upright on the square of alfoil with some melted wax.

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3/ Closely observe the wick. Record what you notice about its colour, position and rate of burning.

4/ Closely observe the wax. Record what you notice about the state and flow of the wax.

5/ Closely observe the air above the candle flame. Record what you notice about the movement and smoke above the flame. Hold a piece of alfoil 10cm above the flame and record any changes.

6/ Place an alfoil tube around the candle flame and note any change.

- Narrow candles burn quickly as most of the wax is lost down the sides.
- Broad candles develop a sunken pool of melted wax and burn slowly.
- When moved, broad candles go out as the pool of wax floods the flame.
- A wick at the side would rapidly burn down.
- Thick candles last longer than thin candles.
- Medium candles optimise portability and burning time.

Result: The flame burns around the wick a short distance above the melted wax. The wick does not appear to burn. Convection currents of warm air above the flame carry soot which condenses onto objects.

Conclusion: The flame is burning wax vapours drawn up the wick from the pool of melted wax below. NOTE: In zero gravity there are no convection currents. The flame would be spherical, the candle would not drip but melt rapidly, the liquid wax tending to flow up the wick and flooding the flame. All circuits in zero gravity must be cooled by fans or radiation.

Risk Level: Low Hazard: Some minor burns are to be expected from hot wax .

Aim: _____

Equipment

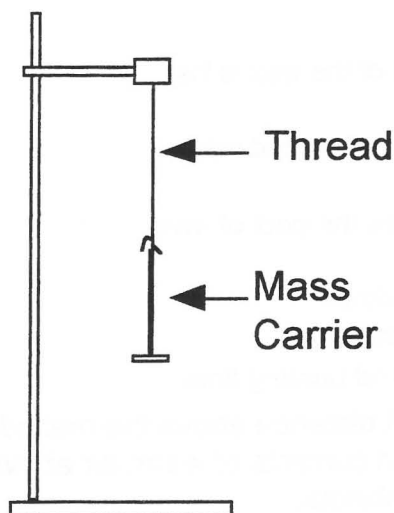
Retort Stand
Boss head and clamp
Cotton thread, gauge 1,
Rubber Band, Gauge 11
Mass Carrier
Masses, 50g (4), 250g,
500g, 1kg.

Procedure

A single experiment can only answer a restricted "Aim".

Perform the following procedure:

1. Tie a thread of cotton to hang from a clamp on a retort stand.
2. Tie a mass carrier to the other end of the thread
3. Add masses to the carrier and record the total mass which breaks the thread. _____g
4. Replace the cotton thread with a rubber band (gauge 11).
5. Add masses and record the total mass which breaks the rubber band. _____g



Thread - Next to the heading "Aim"
at the top of this page, write
a sentence which accurately
states what this experiment
might prove.

Results: _____

Conclusion: _____

Topics: Scientific Method

Forces

Aim: _____**Equipment**

Retort Stand

Boss head and clamp

Cotton thread, gauge 1,

Rubber Band, Gauge 11

Mass Carrier

Masses, 50g (4), 250g,

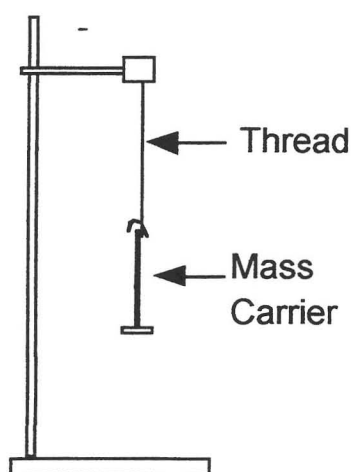
500g, 1kg.

Procedure

A single experiment can only answer a restricted "Aim".

Perform the following procedure:

1. Tie a thread of cotton to hang from a clamp on a retort stand.
2. Tie a mass carrier to the other end of the thread
3. Add masses to the carrier and record the total mass which breaks the thread. _____g
4. Replace the cotton thread with a rubber band (gauge 11).
5. Add masses and record the total mass which breaks the rubber band. _____g



- Next to the heading "Aim" at the top of this page, write a sentence which accurately states what this experiment might prove.

Result: Less mass was required to break the cotton thread than the rubber band.

Conclusion: A Gauge 11 rubber band can usually support more weight than a gauge 1 cotton thread.

Risk Level: Low Hazard; Caution the students that the masses must be suspended over the retort base or it may topple. Allow plenty of room for the rubber band to stretch.

STUDENT: _____

203

Procedure

Aim: To find the flaws in the procedure below designed to compare a laundry powder and a liquid detergent.

Equipment

Laundry powder
Liquid Detergent
Beakers, 250ml, two
Mixed rags
Water supply, hot or cold
stopwatch
stirring rod
Stains: Tomato sauce
Soy sauce
Sump oil

Procedure

- 1/ Choose two rags and put on some stain.
- 2/ Add some water to two beakers.
- 3/ Add some laundry powder to one beaker.
- 4/ Add some liquid detergent to the other beaker.
- 5/ Place a stained rag in each beaker.
- 6/ After a while, remove the rags and see which is cleanest.

- Should the rags be identical? _____
- Should the cloth be white? _____
- Should the same stain be tested on each rag? _____
- Should the same amount of water be in each beaker? _____
- Did you use hot or cold water? _____
- List three other factors that should be the same and should have been mentioned.
a) _____
b) _____
c) _____

Results: _____

Conclusion: _____

Procedure

Topics: Scientific Method Consumer Science

Aim: To find the flaws in the procedure below designed to compare a laundry powder and a liquid detergent.

Equipment

Laundry powder
Liquid Detergent
Beakers, 250ml, two
Mixed rags
Water supply, hot or cold
stopwatch
stirring rod
Stains: Tomato sauce
Soy sauce
Sump oil

Procedure

- 1/ Choose two rags and put on some stain.
- 2/ Add some water to two beakers.
- 3/ Add some laundry powder to one beaker.
- 4/ Add some liquid detergent to the other beaker.
- 5/ Place a stained rag in each beaker.
- 6/ After a while, remove the rags and see which is cleanest.

- Should the rags be identical? _____
- Should the cloth be white? _____
- Should the same stain be tested on each rag? _____
- Should the same amount of water be in each beaker? _____
- Did you use hot or cold water? _____
- List three other factors that should be the same and should have been mentioned.

- a) _____
- b) _____
- c) _____

Result: The result will vary depending on the stain and water temperature.

Conclusion: Many factors must be specified and equal in this experiment. The cloth should be white and the same material. The same stain should be tested and the water temperature must be the same. In addition, the amount used of each detergent should be stated, how long the washing proceeds and how the the cloth is agitated in the water.

Risk Level: Low Hazard; Some laundry powders are caustic and contact with the eyes should be treated by prolonged irrigation. Keep a close eye on the stains lest some smart student decide to "test" the uniform of another student.

STUDENT: _____

204

Results

Aim: To correctly record the results of an experiment measuring the temperature changes in a reaction.

Equipment

Magnesium ribbon, 2cm
Test tube rack
Test tube, medium
Thermometer, 0-100
Graph paper
Hydrochloric acid, 1M(10%)

Procedure

The results of an experiment should include any observations you have made and ideally some measurements.

Measurements should be recorded in a table and then put into a graph which compares the figures with a variable such as time.

1/ In the space below draw up a table that will record a temperature every ten seconds from 0 to 120 seconds.

2/ Add 2cm of hydrochloric acid to a test tube in a test tube rack.

3/ Add a thermometer to the test tube and record the temperature.

4/ Add a 2cm length of magnesium ribbon to the tube.

5/ Record the temperature every 10 seconds for two minutes

6/ Make notes of any observations beside your table.

7/ Use the graph paper to draw up a line graph of the temperature readings versus time.

8/ In the "Results" space below, write a brief summary of your observations and the pattern of temperature change.

Results: _____

Conclusion: _____

Results

Topics: Scientific Method

Exothermic

Aim: To correctly record the results of an experiment measuring the temperature changes in a reaction.

Equipment

Magnesium ribbon, 2cm
Test tube rack
Test tube, medium
Thermometer, 0-100
Graph paper
Hydrochloric acid, 1M(10%)

Procedure

The results of an experiment should include any observations you have made and ideally some measurements. Measurements should be recorded in a table and then put into a graph which compares the figures with a variable such as time.

- 1/ In the space below draw up a table that will record a temperature every ten seconds from 0 to 120 seconds.
- 2/ Add 2cm of hydrochloric acid to a test tube in a test tube rack.
- 3/ Add a thermometer to the test tube and record the temperature.
- 4/ Add a 2cm length of magnesium ribbon to the tube.
- 5/ Record the temperature every 10 seconds for two minutes
- 6/ Make notes of any observations beside your table.
- 7/ Use the graph paper to draw up a line graph of the temperature readings versus time.
- 8/ In the "Results" space below, write a brief summary of your observations and the pattern of temperature change.

Result: Magnesium reacts with hydrochloric acid, gradually dissolving and liberating bubbles of colourless gas. The temperature of the acid rose rapidly, gradually reached a plateau as the reaction ended, and then fell slowly.

Conclusion: Heat energy liberated in the reaction increased the temperature of the acid at a relatively constant rate until the supply of reactants became limiting. In this case the surface area of the magnesium ribbon decreased as it dissolved and this slowed the reaction rate. When the reaction ceased, heat loss to the surroundings caused a slow decline in temperature.

Risk Level: Low Hazard: Hydrochloric acid 1M is mildly corrosive and any contact with skin should be washed with water. The reaction produces hydrogen gas which is explosive if allowed to accumulate.

Conclusion

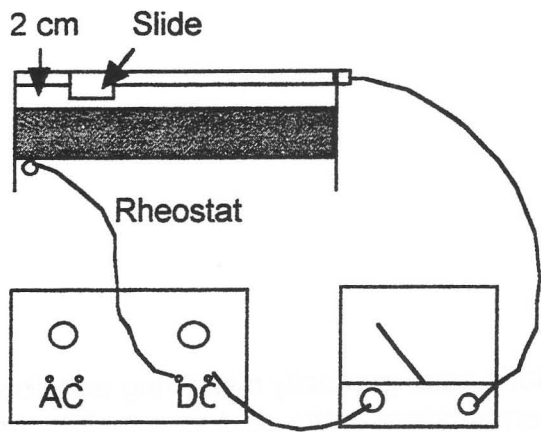
Aim: To write a conclusion for an experiment investigating the relationship between resistance and current in a circuit.

Equipment
Rheostat, 10 ohm
Power supply, 0-12V DC.
Ammeter, 0-5A
Connecting wires, three

Note: The rheostat is a coil of resistance wire. As the slide is moved, resistance increases.

Procedure
The conclusion of an experiment has several roles:
- It must state whether the aim of the experiment was achieved.
- It will attempt to explain what happened in the experiment.
- It may suggest improvements or new experiments to continue the investigation.

- 1/ Set the power supply to 4 Volts DC.
- 2/ Connect the negative DC terminal to the rheostat.
- 3/ Connect the slide bar terminal of the rheostat to the negative terminal of the ammeter.
- 4/ Connect the positive terminal of the ammeter to the positive DC terminal of the power supply.
- 5/ Move the rheostat slide to opposite end from the connecting wire, then move it back 2cm.
- 6/ Turn on the power and record the meter reading.
- 7/ Move the slide another 2cm and record the meter reading.
- 8/ Continue until the slide reaches the other end.
- 9/ Draw a graph of Resistance (cm.) versus Current (amps).



| Resist. (cm) | Current (amps) |
|--------------|----------------|
| 2cm | |
| 4cm | |
| 6cm | |
| 8cm | |
| 10cm | |
| 12cm | |
| 14cm | |

Results: _____

Conclusion: _____

Conclusion

Topics: Scientific Method

Electricity

Aim: To write a conclusion for an experiment investigating the relationship between resistance and current in a circuit.

Equipment

Rheostat, 10 ohm

Power supply, 0-12V DC.

Ammeter, 0-5A

Connecting wires, three

Note: The rheostat is a coil of resistance wire. As the slide is moved, resistance increases.

Procedure

The conclusion of an experiment has several roles:

- It must state whether the aim of the experiment was achieved.

- It will attempt to explain what happened in the experiment.
- It may suggest improvements or new experiments to continue the investigation.

1/ Set the power supply to 4 Volts DC.

2/ Connect the negative DC terminal to the rheostat.

3/ Connect the slide bar terminal of the rheostat to the negative terminal of the ammeter.

4/ Connect the positive terminal of the ammeter to the positive DC terminal of the power supply.

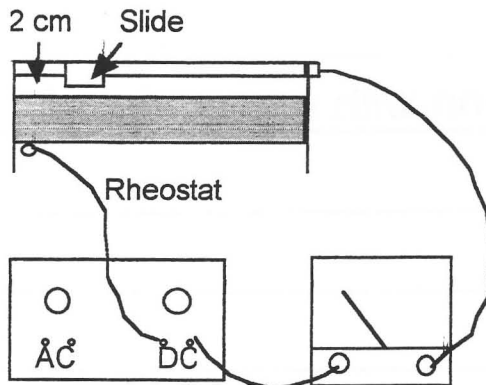
5/ Move the rheostat slide to opposite end from the connecting wire, then move it back 2cm.

6/ Turn on the power and record the meter reading.

7/ Move the slide another 2cm and record the meter reading.

8/ Continue until the slide reaches the other end.

9/ Draw a graph of Resistance (cm.) versus Current (amps).



| Resist. (cm) | Current (amps) |
|-----------------|-------------------|
| 2cm | |
| 4cm | |
| 6cm | |
| 8cm | |
| 10cm | |
| 12cm | |
| 14cm | |

Result: As resistance increased the current in the circuit decreased. A graph of resistance versus current was found to be a curve.

Conclusion: Resistance and Current appear to be inversely related. Increasing the resistance in a circuit restricts the current that can flow at a given voltage. The experiment should be repeated for different voltages and graphs drawn relating voltage and amperage at particular resistances.

Risk Level: Low Hazard



Controls

Aim: To identify the importance of a control in an experiment testing lead contamination in water samples.

Equipment

- Test Tube Rack
- Test tubes, medium, five
- Sodium Iodide, 0.1% in a Dropper bottle.
- Lead Nitrate, 0.1%
- Soil Water (Filtrate of 10g soil & 200ml water).
- Pond Water
- Road Water (from a bitumen puddle)

Procedure

- In an experiment all the factors which might affect the experiment are kept constant except for the one variable being tested. Many factors are difficult or impossible to keep constant eg temperature, humidity, wind, sunshine. The effect of these variables can be eliminated by using a “control”, that is, a part of your experiment identical to the rest except for the test variable.
- 1/ Place five test tubes in a test tube rack.
 - 2/ Add 2cm of tap water to the first tube, 2cm of soil water to the second tube, 2cm of pond water to the third tube, 2cm of road water to the fourth tube and 2cm of Lead solution to the last tube.
 - 3/ Add 10 drops of sodium iodide solution to each tube.
 - 4/ Record your results in the table below.
 - Which tube is the negative control (showing the that no reaction occurs if lead is not present).
 - Which tube is the positive control (showing a particular reaction when lead is known to be present).

| Test Tube | Reaction with sodium iodide |
|-----------|-----------------------------|
| 1 | |
| 2 | |
| 3 | |
| 4 | |
| 5 | |

Results: _____

Conclusion: _____

Topics: Scientific Method Water Chem Tests

Aim: To identify the importance of a control in an experiment testing lead contamination in water samples.

Equipment

- Test Tube Rack
- Test tubes, medium, five
- Sodium Iodide, 0.1% in a Dropper bottle.
- Lead Nitrate, 0.1%
- Soil Water (Filtrate of 10g soil & 200ml water).
- Pond Water
- Road Water (from a bitumen puddle)

Water samples may be simulated with a pinch of clay, and road water simulated with a 1:10 dilution of lead nitrate soln.

Procedure

- In an experiment all the factors which might affect the experiment are kept constant except for the one variable being tested. Many factors are difficult or impossible to keep constant eg temperature, humidity, wind, sunshine. The effect of these variables can be eliminated by using a “control”, that is, a part of your experiment identical to the rest except for the test variable.
- 1/ Place five test tubes in a test tube rack.
 - 2/ Add 2cm of tap water to the first tube, 2cm of soil water to the second tube, 2cm of pond water to the third tube, 2cm of road water to the fourth tube and 2cm of Lead solution to the last tube.
 - 3/ Add 10 drops of sodium iodide solution to each tube.
 - 4/ Record your results in the table below.
 - Which tube is the negative control (showing the that no reaction occurs if lead is not present).
 - Which tube is the positive control (showing a particular reaction when lead is known to be present)

| Test Tube | Reaction with sodium iodide |
|-----------|-----------------------------|
| 1 | |
| 2 | |
| 3 | |
| 4 | |
| 5 | |

Result: The fourth and fifth test tubes produced a bright yellow precipitate.

Conclusion: Sodium iodide produces a bright yellow precipitate with lead ions. The road water sample produces a precipitate due to residues from leaded petrol. The first test tube is a negative control and the last test tube a positive control.

Risk Level: Moderate Hazard: Lead nitrate is a cumulative toxin and contact should be avoided. Sodium iodide may be irritating to skin and is harmful if ingested in quantity. Reagents may be disposed down the sink with flushing.

Variables & Constants

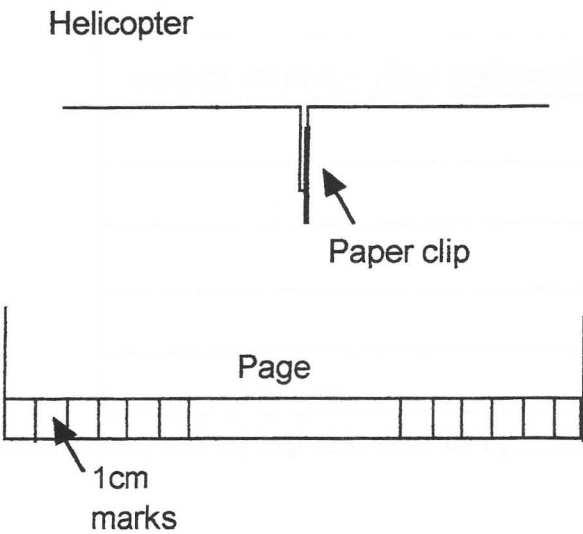
Aim: To identify the factors which are constants and the factors which are variables in an experiment comparing paper helicopters.

Equipment

Stop watch
Scissors
Paper clip

Procedure

- 1/ Rule a line across the bottom of this page, 1.5 cm from the bottom edge.
 - 2./ Draw vertical lines on this strip marking 6cm from each side edge. Continue making vertical lines for each centimetre to the edge on each side.
 - 3/ Cut the strip from the bottom of the page.
 - 4/ Fold the strip in half.
 - 5/ Fold 8cm of each half back to make a broad "T"
 - 6/ Attach a paper clip to the vertical centre piece.
 - 7/ Stand on a seat holding your helicopter at the top of your reach.
 - 8/ Use the stop watch to time how long the helicopter takes to fall to the floor. Record the result in the table.
 - 9/ Cut 1 cm off each wing and repeat step 8.
 - 10/ Repeat step 9, five times.
- From the following list, which have been kept constant?
Width of the strip, height above the floor, wing length, weight, type of paper.
- Which item in the list was the variable being tested?



| Wing Length | Fall Time |
|-------------|-----------|
| 8cm | |
| 7cm | |
| 6cm | |
| 5cm | |
| 4cm | |
| 3cm | |
| 2cm | |

Results: _____

Conclusion: _____

Variables & Constants

Topics: Scientific Method Flight

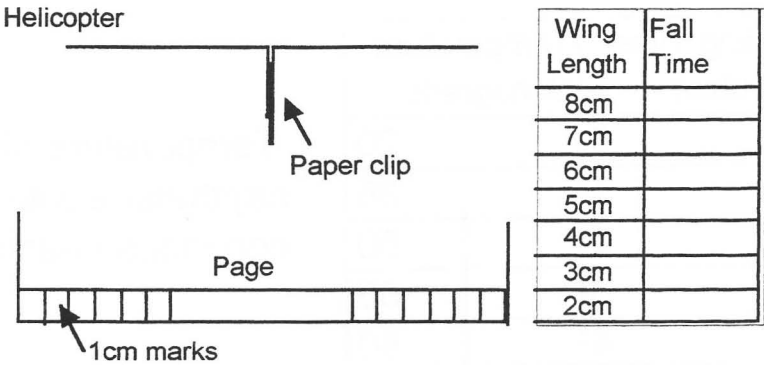
Aim: To identify the factors which are constants and the factors which are variables in an experiment comparing paper helicopters.

Equipment

- Stop watch
- Scissors
- Paper clip

Procedure

- 1/ Rule a line across the bottom of this page, 1.5 cm from the bottom edge.
 - 2./ Draw vertical lines on this strip marking 6cm from each side edge. Continue making vertical lines for each centimetre to the edge on each side.
 - 3/ Cut the strip from the bottom of the page.
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 - 5/ Fold 8cm of each half back to make a broad “T”
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 - 7/ Stand on a seat holding your helicopter at the top of your reach.
 - 8/ Use the stop watch to time how long the helicopter takes to fall to the floor. Record the result in the table.
 - 9/ Cut 1 cm off each wing and repeat step 8.
 - 10/ Repeat step 9, five times.
- From the following list , which have been kept constant?
Width of the strip, height above the floor, wing length, weight, type of paper.
- Which item in the list was the variable being tested?



Result: Time to fall decreases with wing length.

Conclusion: Factors which were held constant include: Wing width, paper type, height above the floor, and weight (since the paper clip greatly exceeds the weight of the paper).
The variable factor being tested in this experiment was wing length.

Risk Level: Low Hazard:

Graphs 1

Aim: To convert tabulated data into a graph from which trends may be recognised.

Equipment

Procedure

- 1/ Subtract the lowest figure from the highest figure in the first column. The difference between these two figures must fit on the horizontal axis of your graph. If the difference is 12, 120 or 12000 draw a line 12cm long on the bottom of the graph paper. If the difference is a number like 3, 30 or 3000 draw a line 3 or 4 times longer than 3cm so your graph is not too small.
- 2/ Mark the left hand end of the line with the lowest figure and the right hand end of the line with the highest figure.
- 3/ Divide the line into equal lengths which correspond to ones, tens or thousands depending on what you need.
- 4/ Label the divisions you have made and write below what the numbers stand for eg. Temperature (centigrade).
- 5/ Repeat steps 1 to 4 for the next column of the table, this time making a vertical line at the left of the page.
- 6/ Look at the first number in column one. Find the position of this number on the bottom line of the graph. Place your ruler vertically marking this position.
- 7/ Look at the first number in column two. Find the position of this number on your line at the left of the graph. Mark a dot on the graph, next to your ruler and which is level with the position you found.
- 8/ Repeat steps 6 and 7 for each of the numbers in the table.
- 9/ Join the dots on your graph.
- 10/ Describe the graph in the "Results" and what it means in the "Conclusion".

| Heating Time (minutes) | Temperature Centigrade |
|---------------------------|---------------------------|
| 0 | 20 |
| 1 | 35 |
| 2 | 50 |
| 3 | 65 |
| 4 | 80 |
| 5 | 80 |
| 6 | 80 |

Temperature of
naphthalene with
continous heating.

Results: _____

Conclusion: _____

Graphs 1

Topics: Scientific Method

Aim: To convert tabulated data into a graph from which trends may be recognised.

Equipment

Procedure

- 1/ Subtract the lowest figure from the highest figure in the first column. The difference between these two figures must fit on the horizontal axis of your graph. If the difference is 12, 120 or 12000 draw a line 12cm long on the bottom of the graph paper. If the difference is a number like 3, 30 or 3000 draw a line 3 or 4 times longer than 3cm so your graph is not too small.
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- 7/ Look at the first number in column two. Find the position of this number on your line at the left of the graph. Mark a dot on the graph, next to your ruler and which is level with the position you found.
- 8/ Repeat steps 6 and 7 for each of the numbers in the table.
- 9/ Join the dots on your graph.
- 10/ Describe the graph in the "Results" and what it means in the "Conclusion".

Result: Temperature increases with heating time until 80 degree centigrade after which the temperature remains constant.

Conclusion: The melting point of naphthalene is 80 degrees centigrade.

Risk Level: Low Hazard.

Graphs 2

Aim: To convert tabulated data into a graph from which trends may be recognised.

Equipment

Procedure

- 1/ Subtract the lowest figure from the highest figure in the first column. The difference between these two figures must fit on the horizontal axis of your graph. If the difference is 12, 120 or 12000 draw a line 12cm long on the bottom of the graph paper. If the difference is a number like 3, 30 or 3000 draw a line 3 or 4 times longer than 3cm so your graph is not too small.
- 2/ Mark the left hand end of the line with the lowest figure and the right hand end of the line with the highest figure.
- 3/ Divide the line into equal lengths which correspond to ones, tens or thousands depending on what you need.
- 4/ Label the divisions you have made and write below what the numbers stand for eg. Temperature (centigrade).
- 5/ Repeat steps 1 to 4 for the next column of the table, this time making a vertical line at the left of the page.
- 6/ Look at the first number in column one. Find the position of this number on the bottom line of the graph. Place your ruler vertically marking this position.
- 7/ Look at the first number in column two. Find the position of this number on your line at the left of the graph. Mark a dot on the graph, next to your ruler and which is level with the position you found.
- 8/ Repeat steps 6 and 7 for each of the numbers in the table.
- 9/ Join the dots on your graph.
- 10/ Describe the graph in the "Results" and what it means in the "Conclusion".

| Temperature Centigrade | Height Kilometres |
|---------------------------|----------------------|
| -100 | 80 |
| 0 | 0 |
| -50 | 10 |
| 200 | 130 |
| -80 | 100 |
| 0 | 50 |
| -50 | 70 |

Variation of temperature
in the Atmosphere with
height above the ground.

Results: _____

Conclusion: _____

Graphs 2

Topics: Scientific Method Atmosphere

Aim: To convert tabulated data into a graph from which trends may be recognised.

Equipment

Procedure

- 1/ Subtract the lowest figure from the highest figure in the first column. The difference between these two figures must fit on the horizontal axis of your graph. If the difference is 12, 120 or 12000 draw a line 12cm long on the bottom of the graph paper. If the difference is a number like 3, 30 or 3000 draw a line 3 or 4 times longer than 3cm so your graph is not too small.
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- 8/ Repeat steps 6 and 7 for each of the numbers in the table.
- 9/ Join the dots on your graph.
- 10/ Describe the graph in the "Results" and what it means in the "Conclusion".

Result: Temperature falls with increasing height to 10km then rises to 50km. Temperature falls again with increasing height to 80km then increases to 130km.

Conclusion: Temperature decreases with increasing height in two regions of the atmosphere, 0 - 10km and 50 - 80km. Elsewhere the temperature in the atmosphere increases with height. There is no simple relationship between height and temperature.

Risk Level: Low Hazard.

STUDENT: _____

210

Hypothesis

Aim: To write a good hypothesis.

Equipment

Procedure

An "Aim" for an experiment is a statement of what you hope an experiment will determine. The most common flaw in an "Aim" is that it is too general eg. "To determine whether manure is the best fertiliser." Such an aim will require not only comparing all possible fertilisers but also measuring their effect on all known plants in all possible conditions. An "Aim" actually possible in a single experiment would be "To determine whether manure produces better growth in lima beans compared to the same weight of 'grow plus' fertiliser."

An Hypothesis is an aim stated in a form which predicts an outcome eg. "Manure produces better growth in lima beans than the same weight of 'grow plus' fertiliser."

Rewrite each of the following "Aims" as a "Hypothesis" .

Aim: To determine if the North poles of magnets repel each other.

Hypothesis: _____

Aim: At what temperature does water boil?

Hypothesis: _____

Aim : To find which ball point pen writes longest.

Hypothesis: _____

Aim: To determine whether soap is better than detergent.

Hypothesis: _____

Results: _____

Conclusion: _____

Hypothesis

Topics: Scientific Method

Aim: To write a good hypothesis.

Equipment

Procedure

An "Aim " for an experiment is a statement of what you hope an experiment will determine. The most common flaw in an "Aim" is that it is too general eg. " To determine whether manure is the best fertiliser." Such an aim will require not only comparing all possible fertilisers but also measuring their effect on all known plants in all possible conditions. An "Aim" actually possible in a single experiment would be " To determine whether manure produces better growth in lima beans compared to the same weight of 'grow plus' fertiliser."

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Rewrite each of the following "Aims" as a "Hypothesis" .

Aim: To determine if the North poles of magnets repel each other.

Hypothesis: _____

Aim: At what temperature does water boil?

Hypothesis: _____

Aim : To find which ball point pen writes longest.

Hypothesis: _____

Aim: To determine whether soap is better than detergent.

Hypothesis: _____

Result: To be left blank.

Conclusion: The North Poles of magnets repel each other. Water boils at 100 degrees centigrade. Brand "X" pen writes longer than other ball point pens. Brand "Y" soap is better for washing grease from dishes in hot water than brand "Z" detergent.

Risk Level: Low Hazard

Parameters

Aim: To select the correct parameter to measure in an experiment examining short term memory.

Equipment
Photocopy of a page from the telephone directory.

Procedure
A Parameter is something which can be measured. In an experiment you should choose a parameter which is clearly related to the subject of your experiment. For example; When measuring growth in clover, height is not a good parameter, leaf width however, closely matches growth of the plant. Perform the following experiment in pairs.
1/ The student acting as experimenter reads the last six digits of a telephone number to the other student.
2/ The student acting as the subject must count five seconds then repeat the number. Record the result in the table below.
3/ Another six digit number is read but this time the subject must count to ten before repeating the number.
4/ Repeat the procedure increasing the count by five each time until the wait time reaches 50.

- Which parameter was a better measure of short term memory, all correct or number of digits correct.

| Wait Time (secs) | Correct (yes or no) | Number of Digits Correct |
|---------------------|------------------------|-----------------------------|
| 5 | | |
| 10 | | |
| 15 | | |
| 20 | | |
| 25 | | |
| 30 | | |
| 35 | | |
| 40 | | |
| 45 | | |
| 50 | | |

Results: _____

Conclusion: _____

Parameters

Topics: Scientific Method

Aim: To select the correct parameter to measure in an experiment examining short term memory.

Equipment

Photocopy of a page from the telephone directory.

Procedure

A Parameter is something which can be measured. In an experiment you should choose a parameter which is clearly related to the subject of your experiment. For example; When measuring growth in clover, height is not a good parameter , leaf width however, closely matches growth of the plant.

Perform the following experiment in pairs.

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3/ Another six digit number is read but this time the subject must count to ten before repeating the number.

4/ Repeat the procedure increasing the count by five each time until the wait time reaches 50.

- Which parameter was a better measure of short term memory, all correct or number of digits correct.

| Wait Time (secs) | Correct (yes or no) | Number of Digits Correct |
|---------------------|------------------------|-----------------------------|
| 5 | | |
| 10 | | |
| 15 | | |
| 20 | | |
| 25 | | |
| 30 | | |
| 35 | | |
| 40 | | |
| 45 | | |
| 50 | | |

Result: The subject could not repeat the whole number correct beyond 5 seconds but could remember some of the digits. The number of digits remembered decreased as the time interval lengthened.

Conclusion: Short term memory lasts only a few seconds before it begins to break down. In this experiment the “number of digits correct” is a better parameter to measure than “correct or incorrect”. “Number of digits correct” actually shows some decay of memory over time whereas “correct or incorrect” shows no change beyond the shortest time intervals.

Risk Level: Low Hazard:

STUDENT: _____

212

Replicates

Aim: To examine the importance of replicates in an experiment to determine the population density of ants.

Equipment

Wire square 20cm X 20cm

Procedure

In many experiments a variety of unknown factors may effect the measurements taken. Replicates means taking several measurements or to performing the experiment several times to be sure the results are reliable.

1/ Go to the school oval with your teacher.

2/ Drop the wire square on the ground.

3/ Carefully examine the area within the coat hanger and count the number of ants seen. Record the result in the table below.

4/ Move ten metres from your first position and repeat step 3.

5/ Repeat step 4 another three times.

6/ Calculate the average number of ants found and estimate the number of ants per square metre.

- List some reasons why your result may be different from other groups.

- What factors may influence the result if the counts were done at a different time or in a different place?

- How might a reliable figure be determined?

| Area | Ants |
|---------|------|
| 1 | |
| 2 | |
| 3 | |
| 4 | |
| 5 | |
| Total | |
| Average | |

Area of Wire Square
= 0.04 sq metre

Ants per sq. metre
= Average X 25

= _____

Results: _____

Conclusion: _____

Replicates

Topics: Scientific Method Invertebrates

Aim: To examine the importance of replicates in an experiment to determine the population density of ants.

Equipment

Wire square 20cm X 20cm

Procedure

In many experiments a variety of unknown factors may effect the measurements taken. Replicates means taking several measurements or to performing the experiment several times to be sure the results are reliable.

- 1/ Go to the school oval with your teacher.
- 2/ Drop the wire square on the ground.
- 3/ Carefully examine the area within the coat hanger and count the number of ants seen. Record the result in the table below.
- 4/ Move ten metres from your first position and repeat step 3.
- 5/ Repeat step 4 another three times.
- 6/ Calculate the average number of ants found and estimate the number of ants per square metre.
- List some reasons why your result may be different from other groups.
- What factors may influence the result if the counts were done at a different time or in a different place?
- How might a reliable figure be determined?

- Coat hangers are easily made into wire square.
- The mass of ants on the Earth exceeds the mass of humans .

| Area | Ants |
|---------|------|
| 1 | |
| 2 | |
| 3 | |
| 4 | |
| 5 | |
| Total | |
| Average | |

Area of Wire Square
= 0.04 sq metre

Ants per sq. metre
= Average X 25

= _____

Result: The counts by each group will vary depending on whether any of their test areas were close to a nest.

Conclusion: Individual groups will have different results depending on their observation skills and nearness to ant nests. Ant numbers seen will vary with the seasons, weather, time of day, soil type and food supply in particular locations. A much larger survey including many measurements at many locations at a variety of times would be needed for a reliable average.

Risk Level: Low Hazard: Avoid known locations of aggressive ants.

Sampling

Aim: To use proper sampling of a population to examine the hypothesis that teenage boys are generally taller than teenage girls.

Equipment

Tape measure
Note to show other teachers.

Procedure

To determine whether teenage boys are taller than teenage girls it will not be enough to compare the average height of both groups in your class. Firstly the sample must be representative eg. include all classes from year 7 to year 12. Secondly the sample must be random ie. students must be selected by a method not related to height.

1/ The teacher will give each group a note which explains your task to other teachers.

2/ The teacher will assign a classroom you are to check. If that room is empty try the room next door.

3/ Go to your assigned room, knock and wait.

4/ Show the teacher in that room your note and ask to measure the height of the first girl on the right side of the class and the first boy on the left side of the class.

5/ Record the results in the table below, return to your class and complete the table with the results from other groups.

- What might be wrong with your sample?
- Devise a plan to determine the number of stars in the sky.

| Group | Class | Girls Height | Boys Height |
|-------|---------|--------------|-------------|
| 1 | | | |
| 2 | | | |
| 3 | | | |
| 4 | | | |
| 5 | | | |
| 6 | | | |
| | Average | | |

Results: _____

Conclusion: _____

Sampling

Topics: Scientific Method

Aim: To use proper sampling of a population to examine the hypothesis that teenage boys are generally taller than teenage girls.

Equipment
Tape measure
Note to show other teachers.

Procedure
To determine whether teenage boys are taller than teenage girls it will not be enough to compare the average height of both groups in your class. Firstly the sample must be representative eg. include all classes from year 7 to year 12. Secondly the sample must be random ie. students must be selected by a method not related to height.
1/ The teacher will give each group a note which explains your task to other teachers.
2/ The teacher will assign a classroom you are to check. If that room is empty try the room next door.
3/ Go to your assigned room, knock and wait.
4/ Show the teacher in that room your note and ask to measure the height of the first girl on the right side of the class and the first boy on the left side of the class.
5/ Record the results in the table below, return to your class and complete the table with the results from other groups.
- What might be wrong with your sample?
- Devise a plan to determine the number of stars in the sky.

Allocate only two elective classrooms as it is likely these will be occupied by one year group.

| Group | Class | Girls Height | Boys Height |
|-------|---------|--------------|-------------|
| 1 | | | |
| 2 | | | |
| 3 | | | |
| 4 | | | |
| 5 | | | |
| 6 | | | |
| | Average | | |

Result: Older teenage boys are generally taller than girls but younger teenage boys are often shorter than girls the same age.

Conclusion: The sample size is too small to positively determine whether teenage boys are taller than teenage girls. To estimate the number of stars in the sky will require counting the stars in a telescope field then repeating this count in other fields **randomly** distributed but including **representative** samples in and out of the Milky Way. The average is then applied to the whole sky.

Risk Level: Low Hazard:

STUDENT: _____

Blind Trial

Aim: To examine the effect of expectations on the result of an experiment.

Equipment

Dropper Bottles:
Sugar 0.1%
Sugar 0.01%
Sugar 0.001%
Salt 0.1%
Salt 0.01%
Salt 0.001%
Water

Blind Fold

Procedure

- 1/ One student in the group is chosen as a subject.
- 2/ Another student in the group will place a drop of one solution on the tongue of the subject.
- 3/ The subject reports whether the solution tastes sweet, salty or nothing. The response is recorded in the table below.
- 4/ Continue until all the solutions have been tasted.
- 5/ Blindfold the subject
- 6/ Repeat the taste test making sure the solutions are tasted in random order, that is without a pattern, not strong to weak or sweet to salt. You may present some solutions twice.
- 7/ Compare the results of test with and without the blindfold.

| Solution | Taste (no blindfold) | Taste (Blindfold) |
|--------------|-----------------------|-------------------|
| Sugar 0.1% | | |
| Sugar 0.01% | | |
| Sugar 0.001% | | |
| Water | | |
| Salt 0.1% | | |
| Salt 0.01% | | |
| Salt 0.001% | | |

Results: _____

Conclusion: _____

Blind Trial

Topics: Scientific Method Coordination

Aim: To examine the effect of expectations on the result of an experiment.

Equipment

Dropper Bottles:
Sugar 0.1%
Sugar 0.01%
Sugar 0.001%
Salt 0.1%
Salt 0.01%
Salt 0.001%
Water

Blind Fold

Procedure

- 1/ One student in the group is chosen as a subject.
- 2/ Another student in the group will place a drop of one solution on the tongue of the subject.
- 3/ The subject reports whether the solution tastes sweet, salty or nothing. The response is recorded in the table below.
- 4/ Continue until all the solutions have been tasted.
- 5/ Blindfold the subject
- 6/ Repeat the taste test making sure the solutions are tasted in random order, that is without a pattern, not strong to weak or sweet to salt. You may present some solutions twice.
- 7/ Compare the results of test with and without the blindfold

NB: This experiment is a great way to compare “Coke” and “Pepsi”.

| Solution | Taste (no blindfold) | Taste (Blindfold) |
|--------------|-----------------------|-------------------|
| Sugar 0.1% | | |
| Sugar 0.01% | | |
| Sugar 0.001% | | |
| Water | | |
| Salt 0.1% | | |
| Salt 0.01% | | |
| Salt 0.001% | | |

Result: The subject was less accurate when blindfolded.

Conclusion: Without a blindfold the subject expected a certain taste from each solution and reported the correct response. This is called experiment bias or “placebo effect”. With the blind fold and with solutions presented in random order the subject did not know what to expect and could not always correctly identify the weaker solutions.

Risk Level: Low Hazard: Beware of students fooling around when blind folded.

Double Blind Trial

Aim:

Equipment

- Dropper Bottles:
- Soln A
- Soln B
- Soln C
- Soln D
- Soln E
- Soln F
- Soln G

Blind fold

Procedure

- 1/ One student in the group is chosen as a subject.
- 2/ Another student in the group will place a drop of one solution on the tongue of the subject.
- 3/ The subject reports whether the solution tastes sweet, salty or nothing. The response is recorded in the table below. Do not taste the solution yourself until after the subject has reported.
- 4/ Continue until all the solutions have been tasted.
- 5/ Blindfold the subject.
- 6/ When the class has finished the taste test the teacher will write the identify the solutions on the board.
- 7/ Repeat the taste test.
- 8/ The teacher will tally the correct responses from each group for each solution, with and without the blindfold.
- 9/ Compare the tallies of correct responses.

| Solution | Taste (blindfold) | Class Tally |
|----------|--------------------|-------------|
| A | | |
| B | | |
| C | | |
| D | | |
| E | | |
| F | | |
| G | | |

Results: _____

Conclusion: _____

Double Blind Trial

Topics: Scientific Method

Coordination

Aim:
Equipment

Dropper Bottles:

Soln A

Soln B

Soln C

Soln D

Soln E

Soln F

Soln G

Blind fold

Procedure

1/ One student in the group is chosen as a subject.

2/ Another student in the group will place a drop of one solution on the tongue of the subject.

3/ The subject reports whether the solution tastes sweet, salty or nothing. The response is recorded in the table below. Do not taste the solution yourself until after the subject has reported.

4/ Continue until all the solutions have been tasted.

5/ Blindfold the subject.

6/ When the class has finished the taste test the teacher will

write the identify the solutions on the board.

7/ Repeat the taste test.

8/ The teacher will tally the correct responses from each group for each solution, with and without the blindfold.

9/ Compare the tallies of correct responses.

C/Sug 0.1%, A/Sug 0.01%

E/ Sug 0.001%, F/ Sal0.1%

G/ Sal0.01%, B/Sal 0.001%

D/ Water

| Solution | Taste (blindfold) | Class Tally |
|----------|--------------------|-------------|
| A | | |
| B | | |
| C | | |
| D | | |
| E | | |
| F | | |
| G | | |

Result: More correct responses were recorded for the weaker solutions when the subjects were blindfolded.

Conclusion: In some experiments the tester sometimes purposely or accidentally gives clues to the subject about the expected result. In a double blind trial neither the tester or the subject knows what to expect. Double blind trials are needed when testing medications or treatments. This experiment makes good use of replication.

Risk Level: Low Hazard: Beware of students clowning around when blind folded.

STUDENT: _____

Residual Error

Aim: To determine residual error as a measure of accuracy in an experiment to determine the acceleration due to gravity.

Equipment

Stop watch
Tennis ball

Procedure

- 1/ One student proceeds with the tennis ball to a first floor walkway and holds the ball level with the top of the railing.
- 2/ A second student remains below with the stop watch.
- 3/ The student with the ball gives a countdown, "3,2,1, go", releasing the ball on "go".
- 4/ The student below times the fall of the ball to the ground.
- 5/ Repeat the procedure six times, recording each result in the table below.
- 6/ For each result calculate "g", the acceleration due to gravity.
- 7/ Calculate the average value for "g"
- 8/ Find the "residual error", the difference between the average and the result most different from the average.
- 9/ Express the result as: average \pm residual error m/sec/sec.

| Trial | Fall time (t) | Height (h) | g (2h/tt) |
|-------|---------------|------------|------------|
| 1 | | | |
| 2 | | | |
| 3 | | | |
| 4 | | | |
| 5 | | | |
| 6 | | | |
| | | Average | |

Results: _____

Conclusion: _____

Residual Error

Topics: Scientific Method Gravity

Aim: To determine residual error as a measure of accuracy in an experiment to determine the acceleration due to gravity.

Equipment
Stop watch
Tennis ball

Procedure
1/ One student proceeds with the tennis ball to a first floor walkway and holds the ball level with the top of the railing.
2/ A second student remains below with the stop watch.
3/ The student with the ball gives a countdown, "3,2,1, go", releasing the ball on "go".
4/ The student below times the fall of the ball to the ground.
5/ Repeat the procedure six times, recording each result in the table below.
6/ For each result calculate "g", the acceleration due to gravity.
7/ Calculate the average value for "g"
8/ Find the "residual error", the difference between the average and the result most different from the average.
9/ Express the result as: average± residual error m/sec/sec.

A tape measure is needed to determine the fall from the railing to the ground.

| Trial | Fall time (t) | Height (h) | g (2h/tt) |
|-------|---------------|------------|------------|
| 1 | | | |
| 2 | | | |
| 3 | | | |
| 4 | | | |
| 5 | | | |
| 6 | | | |
| | | Average | |

Result: The acceleration due to gravity should be 9.8

Conclusion: Residual error gives a measure of the accuracy of the results by showing how much the results may deviate from the average. The main source of error in this experiment is the reaction time of the stop watch operator. Best results are obtained by those operators who anticipate the release and the impact.

Risk Level: Low Hazard:

STUDENT: _____

Statistical Analysis

Aim: To compare results in an experiment involving natural variation eg. Does corn grow faster with urea.

Equipment

Seedling Trays, two,
(30cmX 30cmX 7cm)
40 corn seeds
Potting mix
Urea
Beaker, 800ml

Procedure

- 1/ Fill both seedling trays with moistened potting mix.
- 2/ Plant 20 seeds in each tray , 4 rows of 5 seeds 1cm deep.
- 3/ Label one tray "control" and one tray "urea".
- 4/ Pour 800ml of water evenly over the "control" tray
- 5/ Dissolve 2g of urea in 800ml of water and apply evenly to the " urea" tray.
- 6/ Place both trays in a warm sunny position. Cover the trays with fly screen for 4 days to prevent mice eating the seeds.
- 7/ Water the control tray and apply urea to the test tray every three or four days for two weeks.

SAMPLING- Measure the height of one plant in each tray. Is this a reliable comparison between the control and test group? If not, explain.

AVERAGING- Record the height of each plant in the table below then compare the average height of each group. Compare your "control" average with the "control" averages from other groups. Does this mean your control is better or worse than the controls in other groups?

| Control Group Heights (cm) | | | | Urea Group Heights (cm) | | | |
|----------------------------|--|--|--|-------------------------|--|--|--|
| | | | | | | | |
| | | | | | | | |
| | | | | | | | |
| | | | | | | | |
| | | | | | | | |
| Average (Sum/ number)= | | | | Average (Sum/ number)= | | | |

Results: _____

Conclusion: _____

Statistical Analysis

Topics: Scientific Method

Plants

Aim: To compare results in an experiment involving natural variation eg. Does corn grow faster with urea.

Equipment

Seedling Trays, two,
(30cmX 30cmX 7cm)

40 corn seeds

Potting mix

Urea

Beaker, 800ml

Procedure

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3/ Label one tray "control" and one tray "urea".

4/ Pour 800ml of water evenly over the "control" tray

5/ Dissolve 2g of urea in 800ml of water and apply evenly to the "urea" tray.

6/ Place both trays in a warm sunny position. Cover the trays with fly screen for 4 days to prevent mice eating the seeds.

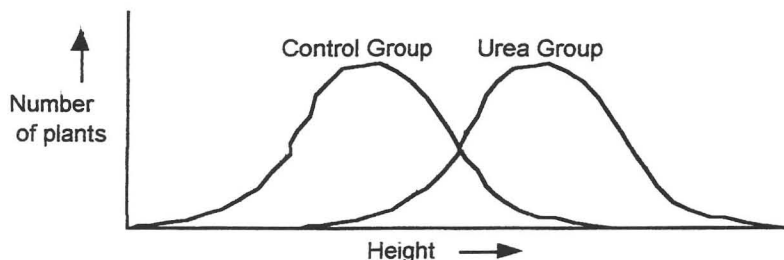
7/ Water the control tray and apply urea to the test tray every three or four days for two weeks.

SAMPLING- Measure the height of one plant in each tray. Is this a reliable comparison between the control and test group? If not, explain.

AVERAGING- Record the height of each plant in the table below then compare the average height of each group.

Compare your "control" average with the "control" averages from other groups. Does this mean your control is better or worse than the controls in other groups?

Teacher - do a Standard Error Analysis between two control groups and between a urea and control group.



Explain to the students that all the measurements for the control groups will fall under the "normal curve" above. While the average for each group will be different , they are not "significantly different". Standard Error Analysis will show the urea group is significantly higher.

Result: Two weeks after planting, the average height of the corn treated with 0.25% urea was higher than the average height of corn given identical watering, soil and sunlight.

Conclusion: Sampling is unreliable because plants vary in height in each group. Sampling needs to be repeated many times. Comparing average heights of the groups showed a difference but the averages of control group were also different due to natural variation in the plants. Standard Error Analysis showed control groups to be the same and urea groups significantly better.

Risk Level: Low Hazard: Urea can irritate the eyes and skin.

STUDENT: _____

218

The Fizz

Aim: To use the reaction of acids and carbonates to make a sweet with a fizzing bite.

Equipment

Cream of Tartar
(ground Tartaric acid)
Baking Soda
(sodium bicarbonate)
Icing sugar
(ground sucrose)
Teaspoon (Food Tech)
Paper Bag, small (canteen)

Procedure

- 1/ Place one third of a teaspoon of creme of tartar in the paper bag.
- 2/ Add one third of a teaspoon of baking soda.
- 3/ Crush any lumps.
- 4/ Shake the paper bag to mix the acid and carbonate.
- 5/ Add two heaped teaspoons of icing sugar.
- 6/ Shake the paper bag for 30 seconds.
- 7/ Taste the product.

- Explain what happens in your mouth to create an unusual sensation.

Results: _____

Conclusion: _____

The Fizz

Topics: Acids & Bases

Aim: To use the reaction of acids and carbonates to make a sweet with a fizzing bite.

Equipment

Cream of Tartar
(ground Tartaric acid)
Baking Soda
(sodium bicarbonate)
Icing sugar
(ground sucrose)
Teaspoon (Food Tech)
Paper Bag, small (canteen)

Procedure

- 1/ Place one third of a teaspoon of creme of tartar in the paper bag.
- 2/ Add one third of a teaspoon of baking soda.
- 3/ Crush any lumps.
- 4/ Shake the paper bag to mix the acid and carbonate.
- 5/ Add two heaped teaspoons of icing sugar.
- 6/ Shake the paper bag for 30 seconds.
- 7/ Taste the product.

- Explain what happens in your mouth to create an unusual sensation.

This practical is very popular with students and often related to parents.

Result: The mixture tastes sweet but also tingles in a biting fizz.

Conclusion: Acids and carbonates react to produce bubbles of carbon dioxide, a salt and water. Tartaric acid and baking soda cannot react until dissolved. When the mixture is placed in your mouth, saliva provides a liquid to dissolve the ingredients and allow the reaction to take place. Your tongue is assaulted by bubbles of carbon dioxide and brief tastes of bitter and sour.

Risk Level: Low Hazard: Students always think more is better. They are likely to double or triple the quantities unless you are vigilant. Students are also prone to eat more than is good for themselves. The sugar boost is also likely to make their behaviour a problem. This practical is best performed in the last 20 minutes before a break.

STUDENT: _____

219

Amphoteric salts

Aim: To demonstrate that some salts have both acidic and basic properties.

Equipment

Test Tubes, 2
Test Tube Rack
Dropper Bottles of:
Hydrochloric Acid
Sodium Hydroxide (1%)
Suspensions of:
Zinc Hydroxide(1%)
Aluminium Hydroxide (1%)
Copper Hydroxide (1%)
Calcium Hydroxide (1%)

Procedure

1/ Add about 2cm of the copper hydroxide suspension to each of the test tubes.
2/ To one tube add drops of hydrochloric acid and record any change.
3/ To the second tube add drops of sodium hydroxide and record any change.
4/ Thoroughly rinse the tubes and repeat the experiment for zinc hydroxide.
5/ Repeat the steps for calcium hydroxide and then aluminium hydroxide.

| Hydroxide | Reaction with Acid | Reaction with Base |
|-----------|--------------------|--------------------|
| Copper | | |
| Zinc | | |
| Calcium | | |
| Aluminium | | |

Results: _____

Conclusion: _____

Amphoteric salts

Topics: Acids and Bases

Aim: To demonstrate that some salts have both acidic and basic properties.

Equipment

Test Tubes, 2
Test Tube Rack
Dropper Bottles of:
Hydrochloric Acid
Sodium Hydroxide (1%)
Suspensions of:
Zinc Hydroxide (1%)
Aluminium Hydroxide (1%)
Copper Hydroxide (1%)
Calcium Hydroxide (1%)

Procedure

1/ Add about 2cm of the copper hydroxide suspension to each of the test tubes.
2/ To one tube add drops of hydrochloric acid and record any change.
3/ To the second tube add drops of sodium hydroxide and record any change.
4/ Thoroughly rinse the tubes and repeat the experiment for zinc hydroxide.
5/ Repeat the steps for calcium hydroxide and then aluminium hydroxide.

If some hydroxides are unavailable, dilute NaOH is added drop wise to 1% nitrate solution until precipitation.

| Hydroxide | Reaction with Acid | Reaction with Base |
|-----------|--------------------|--------------------|
| Copper | | |
| Zinc | | |
| Calcium | | |
| Aluminium | | |

Result: Copper and calcium hydroxide react only with acid. Zinc and aluminium hydroxide react with both acid and with a base.

Conclusion: Zinc and aluminium hydroxide are amphoteric, that is, reacting as both acids and bases eg.



Risk Level: Mild Hazard: Zinc and copper salts are harmful if ingested. Calcium hydroxide is caustic and contact with the skin should be avoided. Sodium hydroxide is very caustic and any contact with the skin treated with prolonged rinsing with water. Hydrochloric acid is corrosive however at 1% concentration poses little hazard.

STUDENT: _____

220

Metallic Order

Aim: To determine a rank order or chemical activity among a selection of metals.

Equipment

Nitric Acid, 0.1M (0.6%, in a
dropper bottle)
Test tube rack
Test Tubes, medium, six
Copper strip
Iron pieces
Tin granules
Lead shot
Magnesium Ribbon
Zinc granules
Universal Indicator (in a
dropper bottle)
Steel wool

Procedure

- 1/ Add 20 drops of nitric acid to each test tube.
- 2/ Add two drops of universal indicator to each test tube.
- 3/ Clean the copper samples with steel wool and rinse in water.
- 4/ Add a sample of copper to the first tube, iron to the second and so on until each tube contains a metal.
- 5/ After a while the solutions will begin to change colour from red to orange indicating reaction with the acid.
- 6/ Record the order in which the metals change the the indicator colour.

| Reactivity order | Metal |
|---------------------|-------|
| 1st | |
| 2nd | |
| 3rd | |
| 4th | |
| 5th | |
| 6th | |

Results: _____

Conclusion: _____

Metallic Order

Topics: Acids and Bases

Elements

Aim: To determine a rank order or chemical activity among a selection of metals.**Equipment**

Nitric Acid, 0.1M (0.6%, in a
dropper bottle)
Test tube rack
Test Tubes, medium, six
Copper strip
Iron pieces
Tin granules
Lead shot
Magnesium Ribbon
Zinc granules
Universal Indicator (in a
dropper bottle)
Steel wool

Procedure

1/ Add 20 drops nitric acid to each test tube.
2/ Add two drops of universal indicator to each test tube.
3/ Clean the copper samples with steel wool and rinse in water.
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5/ After a while the solutions will begin to change colour from red to orange indicating reaction with the acid.
6/ Record the order in which the metals change the the indicator colour.

Plan a second activity as
this experiment may take an
hour.

| Reactivity order | Metal |
|------------------|-------|
| 1st | |
| 2nd | |
| 3rd | |
| 4th | |
| 5th | |
| 6th | |

Result: The order of colour change is Magnesium first, then zinc, iron, tin, copper and finally lead.

Conclusion: The order of reactivity of the metals (most reactive to least reactive) is Magnesium, zinc, iron, tin, copper and lead.
Surface area affects reaction rate but in this case the difference in reactivity should more than compensate.

Risk Level: Low Hazard: Nitric acid is corrosive but poses negligible hazard at 0.1M concentration.

STUDENT: _____

221

Plant Indicators

Aim: To make a dye solution which indicates the presence of acids or bases

Equipment

Red Cabbage (1/2 cup)

Beaker, 250ml

Wire gauze

Tripod

Bunsen

Test Tube Rack

Test tubes, medium, two

Hydrochloric acid, 0.1M(1%)

Sodium Hydroxide,
0.1M(4%)

Dropper bottle

Procedure

1/ Place the chopped plant material in the beaker.

2/ Half fill the beaker with water.

3/ Place the beaker on the wire gauze and tripod.

4/ Heat the beaker with a Bunsen until the water has boiled for five minutes.

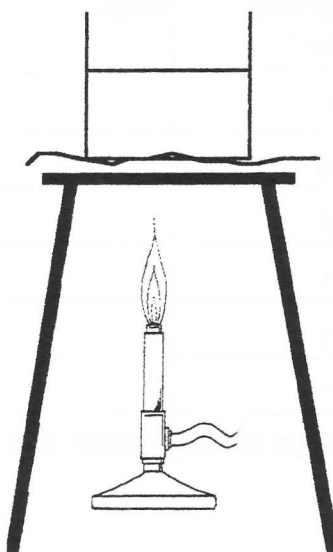
5/ Pour some of the liquid into a dropper bottle.

6/ Add 2cm of acid to one test tube.

7/ Add a few drops of your indicator liquid.

8/ Add 2cm of sodium hydroxide to the second test tube.

9/ Add a few drops of your indicator liquid.



Results: _____

Conclusion: _____

Plant Indicators

Topics: Acids and Bases

Plants

Aim: To make a dye solution which indicates the presence of acids or bases**Equipment**

Red Cabbage (1/2 cup)

Beaker, 250ml

Wire gauze

Tripod

Bunsen

Test Tube Rack

Test tubes, medium, two

Hydrochloric acid,

0.1M(1%)

Sodium Hydroxide,

0.1M(4%)

Dropper bottle

Procedure

1/ Place the chopped plant material in the beaker.

2/ Half fill the beaker with water.

3/ Place the beaker on the wire gauze and tripod.

4/ Heat the beaker with a Bunsen until the water has boiled for five minutes.

5/ Pour some of the liquid into a dropper bottle.

6/ Add 2cm of acid to one test tube.

7/ Add a few drops of your indicator liquid.

8/ Add 2cm of sodium hydroxide to the second test tube.

9/ Add a few drops of your indicator liquid.

Grated beetroot or red
hibiscus flowers are
acceptable alternatives to
red cabbage.

Result: The purple dye extracted from the cabbage turns red in acid and blue in a base.**Conclusion:** The natural dye from cabbage will indicate the presence of acids or bases.**Risk Level:** Low Hazard: At 0.1M concentrations, hydrochloric acid and sodium hydroxide pose very little hazard.

STUDENT: _____

222

Titration 1

Aim: To determine the concentration of an acid by reaction with a carbonate.

Equipment

Burette and burette clamp
Retort stand
Measuring cylinder, 100ml
Conical flask, 250ml
Methyl Red, 1% (in dropper bottle)
Sodium Hydrogen Carbonate, 0.1M (8.4 g dissolved in 1 litre water in a volumetric flask.)
Hydrochloric Acid
Filter funnel, small

Procedure

- 1/ Carefully rinse the burette with the acid.
- 2/ Fill the burette with acid to the top then clamp the burette to the retort stand so there is about 15cm clearance to the bench.
- 3/ Place the conical flask beneath the burette and release the burette tap until the acid reaches the 0.0ml mark.
- 4/ Thoroughly rinse the conical flask and then add 50ml of the sodium hydrogen carbonate solution.
- 5/ Add four drops of Methyl Red solution to the flask.
- 6/ Place a sheet of white paper under the flask .
- 7/ Adjust the burette tap to slowly add acid to the flask.
- 8/ Swirl the contents of the conical flask with your left hand while grasping the burette tap with your right hand.
- 9/ Splashes of red will appear in the flask as the acid enters but then disappear as the solution is swirled.
- 10/ When the splashes of red begin to linger, slow the burette drop rate to about 1 every 2 seconds.
- 11/ When the red colour no longer clears, turn off the burette tap and record how much acid was used.

Results: _____

Conclusion: _____

Titration 1

Topics: Acids and Bases

Aim: To determine the concentration of an acid by reaction with a carbonate.

Equipment

Burette and burette clamp
Retort stand
Measuring cylinder, 100ml
Conical flask, 250ml
Methyl Red, 1% (in
dropper bottle)
Sodium Hydrogen
Carbonate, 0.1M (8.4 g
dissolved in 1 litre water in
a volumetric flask.)
Hydrochloric Acid
Filter funnel, small

Hydrochloric acid 0.2M, 2%

Procedure

- 1/ Carefully rinse the burette with the acid.
- 2/ Fill the burette with acid to the top then clamp the burette to the retort stand so there is about 15cm clearance to the bench.
- 3/ Place the conical flask beneath the burette and release the burette tap until the acid reaches the 0.0ml mark.
- 4/ Thoroughly rinse the conical flask and then add 50ml of the sodium hydrogen carbonate solution.
- 5/ Add four drops of Methyl Red solution to the flask.
- 6/ Place a sheet of white paper under the flask .
- 7/ Adjust the burette tap to slowly add acid to the flask.
- 8/ Swirl the contents of the conical flask with your left hand while grasping the burette tap with your right hand.
- 9/ Splashes of red will appear in the flask as the acid enters but then disappear as the solution is swirled.
- 10/ When the splashes of red begin to linger, slow the burette drop rate to about 1 every 2 seconds.
- 11/ When the red colour no longer clears, turn off the burette tap and record how much acid was used.

Result: 25ml of acid was needed to neutralise 50ml of the sodium hydrogen carbonate.

Conclusion: The acid is twice as concentrated as the 1M sodium hydrogen carbonate solution.

Risk Level: Moderate Hazard: Hydrochloric acid is corrosive and any contact with the skin should be vigorously washed in water. While the chemicals pose only a moderate hazard, junior students are likely to break expensive burettes.

STUDENT: _____

223

Titration 2

Aim: To determine the molarity of a base solution

Equipment

Burette and burette clamp
Retort stand
Measuring cylinder, 100ml
Conical flask, 250ml
Methyl Red, 1% (in dropper bottle)
Hydrochloric Acid, 0.20M
Filter funnel, small
Sodium Hydroxide solution

Procedure

- 1/ Carefully rinse the burette with the acid.
- 2/ Fill the burette with acid to the top then clamp the burette to the retort stand so there is about 15cm clearance to the bench.
- 3/ Place the conical flask beneath the burette and release the burette tap until the acid reaches the 0.0ml mark.
- 4/ Thoroughly rinse the conical flask and then add 30ml of the sodium hydroxide solution.
- 5/ Add four drops of Methyl Red solution to the flask.
- 6/ Place a sheet of white paper under the flask .
- 7/ Adjust the burette tap to slowly add acid to the flask.
- 8/ Swirl the contents of the conical flask with your left hand while grasping the burette tap with your right hand.
- 9/ Splashes of red will appear in the flask as the acid enters but then disappear as the solution is swirled.
- 10/ When the splashes of red begin to linger, slow the burette drop rate to about 1 every 2 seconds.
- 11/ When the red colour no longer clears, turn off the burette tap and record how much acid was used.



Moles of Acid used = Moles of Base neutralised

Moles of acid used = Molarity X volume

$$= 0.20 \times \text{_____} (\text{litres})$$

$$= \text{_____}$$

Molarity of Base = Moles / volume

$$= \text{_____} / \text{_____}$$

$$= \text{_____} \text{M}$$

Results: _____

Conclusion: _____

Titration 2

Topics: Acids and Bases

Molarity
Aim: To determine the molarity of a base solution

Equipment

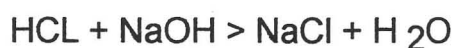
Burette and burette clamp
Retort stand
Measuring cylinder, 100ml
Conical flask, 250ml
Methyl Red, 1% (in
dropper bottle)
Hydrochloric Acid, 0.20M
Filter funnel, small
Sodium Hydroxide solution

Sodium Hydroxide, 0.15M,
6%

Use the standardised acid
from Titration 1 adjusted to
0.20M.

Procedure

- 1/ Carefully rinse the burette with the acid.
- 2/ Fill the burette with acid to the top then clamp the burette to the retort stand so there is about 15cm clearance to the bench.
- 3/ Place the conical flask beneath the burette and release the burette tap until the acid reaches the 0.0ml mark.
- 4/ Thoroughly rinse the conical flask and then add 30ml of the sodium hydroxide solution.
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- 9/ Splashes of red will appear in the flask as the acid enters but then disappear as the solution is swirled.
- 10/ When the splashes of red begin to linger, slow the burette drop rate to about 1 every 2 seconds.
- 11/ When the red colour no longer clears, turn off the burette tap and record how much acid was used.



Moles of Acid used = Moles of Base neutralised

Moles of acid used = Molarity X volume

$$= 0.20 \times 0.0225(\text{litres})$$

$$= 0.0045$$

Molarity of Base = Moles / volume

$$= 0.045 / 0.030$$

$$= 1.50 \text{ M}$$

Result: 22.5mls of acid was required to neutralise 30mls of sodium hydroxide solution.

Conclusion: The concentration of the sodium hydroxide solution is 0.15M.

Risk Level: Moderate Hazard: Sodium hydroxide is caustic and hydrochloric acid is corrosive. If either of these reagents should contact skin the area should be treated with prolonged washing with water. While the chemicals are only moderately hazardous, junior students are very likely to break expensive burettes.

STUDENT: _____

224

Tin Canometer

Aim: To measure the height of clouds with a tin can and a thermometer.

Equipment

Tin can
Thermometer, 0-100

Procedure

- Dew Point is the temperature at which dew begins to form.
- 1/ Fill a tin can with tepid water.
 - 2/ Place a thermometer in the can and hold it against the metal side.
 - 3/ Add a half a cup of crushed ice to the can.
 - 4/ Record the temperature at which moisture just begins to form on the outside of the can.
 - 5/ Remove and dry the thermometer then hold it in the air, out of the sun. Record the temperature after 3 minutes.

Air temperature decreases with height from the ground, about one degree for every 120 metres. When the Dew Point temperature is reached, water vapour in the air begins to condense to form cloud. The height of cumulus clouds on a particular day can be calculated as follows:

$$H = (T - DP) \times 120 , H = \text{height in metres}$$

T = air temperature
DP = Dew Point temperature

Results: _____

Conclusion: _____

Tin Canometer

Topics: Atmosphere

Matter

Aim: To measure the height of clouds with a tin can and a thermometer.**Equipment**

Tin can

Thermometer, 0-100

Procedure

Dew Point is the temperature at which dew begins to form.

1/ Fill a tin can with tepid water.

2/ Place a thermometer in the can and hold it against the metal side.

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4/ Record the temperature at which moisture just begins to form on the outside of the can.

5/ Remove and dry the thermometer then hold it in the air, out of the sun. Record the temperature after 3 minutes.

Air temperature decreases with height from the ground, about one degree for every 120 metres. When the Dew Point temperature is reached, water vapour in the air begins to condense to form cloud. The height of cumulus clouds on a particular day can be calculated as follows:

$$H = (T - DP) \times 120, H = \text{height in metres}$$

T = air temperature

DP = Dew Point temperature

Result: Typically Cumulus cloud forms at around 1000 metres above the ground.**Conclusion:** The height is in metres above your position. The Dew Point temperature will vary depending on the humidity in the air and air temperature.**Risk Level:** Low Hazard:

STUDENT: _____

225

Conservation of Mass

Aim: To demonstrate that the total mass of the products of a reaction must equal the total mass of the reactants.

Equipment

Glass Jar (wide mouth and screw cap lid)
Specimen Tube (to fit inside the glass jar)
Balance, 0.01g
Forceps
Hydrochloric Acid 1M(10%)
Copper Carbonate

Procedure

- 1/ Accurately weight the glass jar (with lid).
- 2/ Weigh 0.5g of copper carbonate into the jar.
- 3/ Accurately weigh the specimen tube.
- 4/ Weigh 5.0g of hydrochloric acid (1M) into the specimen tube.
- 5/ Using forceps, carefully lower the specimen tube into the jar.
- 6/ Seal the jar
- 7/ Briefly up end the jar to mix the reactants.
- 8/ Reweigh the Jar.
- 9/ Release the lid of the jar, allowing the gasses to escape.
- 10/ Reseal and reweigh the jar.

- Write a balanced equation for the reaction.
- Name the gas and the salt produced.
- Why is it not possible to calculate the weight of salt produced?

| Object | Weight | Calculated Weight |
|-----------------------------|--------|-------------------|
| Glass Jar with lid | | X |
| Copper Carbonate added | | X |
| Specimen Tube | | X |
| Hydrochloric Acid added | | X |
| Weight of Reactants | X | |
| Total Weight after Reaction | | X |
| Weight of Products | X | |
| Weight after Gas released | | X |
| Weight of gas produced | X | |

Results: _____

Conclusion: _____

Conservation of Mass

Topics: Chem Rns

Aim: To demonstrate that the total mass of the products of a reaction must equal the total mass of the reactants.

Equipment

Glass Jar (wide mouth and screw cap lid)
Specimen Tube (to fit inside the glass jar)
Balance, 0.01g
Forceps
Hydrochloric Acid 1M(10%)
Copper Carbonate

Procedure

- 1/ Accurately weight the glass jar (with lid).
- 2/ Weigh 0.5g of copper carbonate into the jar.
- 3/ Accurately weigh the specimen tube.
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- 7/ Briefly up end the jar to mix the reactants.
- 8/ Reweigh the Jar.
- 9/ Release the lid of the jar, allowing the gasses to escape.
- 10/ Reseal and reweigh the jar.

- Write a balanced equation for the reaction.
- Name gas and the salt produced.
- Why is it not possible to calculate the weight of salt produced?

| Object | Weight | Calculated Weight |
|-----------------------------|--------|-------------------|
| Glass Jar with lid | | X |
| Copper Carbonate added | | X |
| Specimen Tube | | X |
| Hydrochloric Acid added | | X |
| Weight of Reactants | X | |
| Total Weight after Reaction | | X |
| Weight of Products | X | |
| Weight after Gas released | | X |
| Weight of gas produced | X | |

Result: The copper carbonate reacts with the acid producing bubbles of gas and a dark green solution.

Conclusion: The mass of the reactants equalled the mass of the products.

$\text{CuCO}_3 + 2\text{HCl} \rightarrow \text{CuCl}_2 + \text{H}_2\text{O} + \text{CO}_2$. The gas produced is carbon dioxide.

The salt produced is copper chloride. Since water is also produced in the reaction only the weight of the salt plus the water can be deduced.

Risk Level: Moderate Hazard: Hydrochloric acid is corrosive and contact with the skin should be treated with prolonged washing. Copper salts are harmful by ingestion and skin contact should be avoided.

STUDENT: _____

Reaction Rate

Aim: To examine the factors affecting reaction rate.

Equipment

Test Tube ,20ml, 5
Test Tube Rack
Test Tube clamp
Thermometer, 0 - 100C
Stick of chalk
Hydrochloric Acid 1M(10%)
Measuring Cylinder, 10ml
Filter paper
Hot water or boiling beaker
at teacher bench.

Procedure

- 1/ Place four test tubes in the rack.
- 3/ Break off a 1cm piece of chalk and add it to tube 1.
- 4/ Use a ruler to crush the remaining chalk on a piece paper and divide the powder into three equal piles.
- 5/ Scrape one pile onto a filter paper and pour the powder into test tube 2. Repeat to add powder to tubes 3 and 4.
- 6/ Measure 5mls of water into tube 2.
- 7/ Measure 5mls of 1M Hydrochloric acid into test tube 1 and record your observations in the table below.
- 8/ Measure 5mls of acid into test tube 2. Record whether the reaction is faster or slower than tube 1.
- 9/ Measure 5mls of acid into test tube 3. Record whether the reaction is faster or slower than tube 2.
- 10/ Add 5mls of acid to a clean test tube. Add a thermometer to the tube and heat to 50 C by placing in hot water.
- 11/ Add the heated acid to tube 4. Record whether the reaction is faster or slower than tube 2.

| Trial | Observations |
|------------------------------|--------------|
| 1. Chalk piece, acid | |
| 2. Chalk powder,diluted acid | |
| 3. Chalk powder, acid | |
| 4. Chalk powder, hot acid | |

Results: _____

Conclusion: _____

Reaction Rate

Topics: Chemical Rns

Equilibrium

Aim: To examine the factors affecting reaction rate.**Equipment**

Test Tube ,20ml, 5

Test Tube Rack

Test Tube clamp

Thermometer, 0 - 100C

Stick of chalk

Hydrochloric Acid 1M(10%)

Measuring Cylinder, 10ml

Filter paper

Hot water or boiling beaker
at teacher bench.**Procedure**

1/ Place four test tubes in the rack.

3/ Break off a 1cm piece of chalk and add it to tube 1.

4/ Use a ruler to crush the remaining chalk on a piece paper
and divide the powder into three equal piles.5/ Scrape one pile onto a filter paper and pour the powder into
test tube 2. Repeat to add powder to tubes 3 and 4.

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reaction is faster or slower than tube 1.9/ Measure 5mls of acid into test tube 3. Record whether the
reaction is faster or slower than tube 2.10/ Add 5mls of acid to a clean test tube. Add a thermometer
to the tube and heat to 50 C by placing in hot water.11/ Add the heated acid to tube 4. Record whether the
reaction is faster or slower than tube 2.

| Trial | Observations |
|------------------------------|--------------|
| 1. Chalk piece, acid | |
| 2. Chalk powder,diluted acid | |
| 3. Chalk powder, acid | |
| 4. Chalk powder, hot acid | |

Result: Chalk reacts with acid to produce bubbles and heat. The reaction was slowest in tube 1, faster in tube 2, faster again in tube 3 and very fast in tube 4.**Conclusion:** The rate of a reaction is increased by : increasing the surface area for reaction (crushing the chalk) , increasing the concentration of acid and by increasing the temperature.**Risk Level:** Moderate Hazard: 1M Hydrochloric acid is mildly corrosive. Skin contact should be avoided and treated with thorough washing. Heating the acid is an additional hazard.

Ignition

Aim: To safely ignite a reactive mixture.

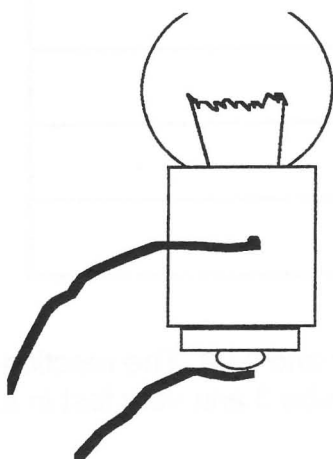
Equipment

Torch globe, 3V
Triangular file, small
Insulated wire, 30cm, two
Soldering Iron, electric
Heat Mat
Electrical solder
Chemical powder
5m of twin lead wires
Power supply, 12V DC
Wire stripping tool

Procedure

- 1/ File the tip of the globe until a hole is made in the glass.
 - 2/ Place the globe and soldering iron on the heat tile.
 - 3/ Turn on the soldering iron and wait 3 minutes for heating.
 - 4/ Strip both ends of the two insulated wires.
 - 5/ Rest the end of one wire on the metal sleeve of the globe.
 - 6/ Press the soldering iron to the wire end and count to five.
 - 7/ Touch some electrical solder to the soldering iron tip until a drop of solder melts onto the wire and globe.
 - 8/ Place the end of the second wire against the round metal base of the globe.
 - 9/ Repeat steps 6 and 7.
- IN THE FUME HOOD AND UNDER SUPERVISION**
- 10/ Gently add a little chemical powder into the globe.
 - 11/ Place a teaspoon of chemical mix on a heat tile and nestle the globe into the mix.
 - 12/ Connect the globe wires to the long lead wires.
 - 13/ At a bench away from the fume hood, connect the lead wires to the power supply DC terminals.
 - 14/ Set the supply to 12V and turn on the power.

Torch Globe Ignition Fuse



Under no circumstances should the fuse wires be connected to the lead wires unless the lead wires are disconnected from the power supply

Results: _____

Conclusion: _____

Ignition

Topics: Chemical Runs

Aim: To safely ignite a reactive mixture.

Equipment

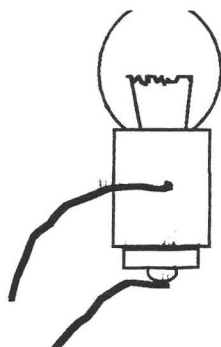
Torch globe, 3V
 Triangular file, small
 Insulated wire, 30cm, two
 Soldering Iron, electric
 Heat Mat
 Electrical solder
 Chemical powder
 5m of twin lead wires
 Power supply, 12V DC
 Wire stripping tool

Chemical Powder: finely
 ground sodium nitrate
 mixed with an equal volume
 of pure icing sugar.

Procedure

- 1/ File the tip of the globe until a hole is made in the glass.
 - 2/ Place the globe and soldering iron on the heat tile.
 - 3/ Turn on the soldering iron and wait 3 minutes for heating.
 - 4/ Strip both ends of the two insulated wires.
 - 5/ Rest the end of one wire on the metal sleeve of the globe.
 - 6/ Press the soldering iron to the wire end and count to five.
 - 7/ Touch some electrical solder to the soldering iron tip until a drop of solder melts onto the wire and globe.
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Torch Globe Ignition Fuse



Under no circumstances should the fuse wires be connected to the lead wires unless the lead wires are disconnected from the power supply

Result: The torch globe fuse instantly ignites the chemical powder.

Conclusion: Advantages: Nitrate or fuel soaked wicks can fail half way or have delays inviting close approach. Electrical fuses from hobby shops have a chemical coating which can cause fires. The globe size prevents effective confinement which can transform a fast burn into an explosion. Up to 100m of lead wire can be used with a 12V battery.

Risk Level: Moderate Hazard: Most boys will at some stage experiment with explosive mixtures. Accidents from these adventures are most often a result of attempts at ignition with matches or wicks. Showing the students how to create simple, cheap and safe electrical fuses may help avoid some of the accidents which occur every year. **CLOSELY SUPERVISE THE CHEMICAL MIX AND THE IGNITION TESTS.**

STUDENT: _____

228

Quantitative Assay

Aim: To determine the phosphate concentration in a water sample using colourmetric assay.

Equipment

Molybdate Reagent: Add 33ml of sulfuric acid to 500ml of water. Add water to 1 litre final volume. Add 2g ammonium molybdate.
Test tubes, medium,
Test tube racks, two
Laundry powder
Potassium Dihydrogen Phosphate, 0.00001M (0.136g in 1litre water)
Measuring cylinder, 100ml
Colourimeter (optional)
Pipette, 10ml
Pipette, 1ml, two
Beaker, 250ml, two
Sulfuric acid, 0.1M, 5%

Procedure

- 1/ Rinse all glassware in 0.1M sulfuric acid then in water.
- 2/ Add 1g of the laundry powder to 100ml of water in a beaker and dissolve.
- 3/ Dilution 1: Pipette 0.5ml of the solution into a test tube then add 4.5ml of water.
- 4/ Dilution 2: Pipette 0.5ml of Dilution 1 into a test tube and then add 4.5ml of water.
- 5/ Add 1ml of the potassium dihydrogen phosphate standard to a beaker and add 99ml of water. (1/100 standard)
- 6/ Dilute 1ml of this solution with 4mls of water in a test tube.
- 7/ Dilute 1ml of this solution with 4mls of water in a test tube.
- 8/ In a separate test tube rack add the following to separate test tubes: 0.5ml of laundry powder solution, 0.5ml of Dilution 1, 0.5ml of Dilution 2, 0.5ml of diluted standard 1/100, 0.5ml of diluted standard 1/500, 0.5ml of diluted standard 1/2500 and finally 0.5ml of water.
- 9/ Add 4.5ml of molybdate solution to the tubes in step 8.
- 10/ After 15 minutes the tubes may be compared or read in a colorimeter.

Results: _____

Conclusion: _____

Quantitative Assay

Topics: Chemical Runs

Water

Aim: To determine the phosphate concentration in a water sample using colourmetric assay.

Equipment

Molybdate Reagent: Add 33ml of sulfuric acid to 500ml of water. Add water to 1 litre final volume. Add 2g ammonium molybdate.
 Test tubes, medium,
 Test tube racks, two
 Laundry powder
 Potassium Dihydrogen Phosphate, 0.00001M (0.136g in 1litre water)
 Measuring cylinder, 100ml
 Colourimeter (optional)
 Pipette, 10ml
 Pipette, 1ml, two
 Beaker, 250ml, two
 Sulfuric acid, 0.1M, 5%

Procedure

- 1/ Rinse all glassware in 0.1M sulfuric acid then in water.
- 2/ Add 1g of the laundry powder to 100ml of water in a beaker and dissolve.
- 3/ Dilution 1: Pipette 0.5ml of the solution into a test tube then add 4.5ml of water.
- 4/ Dilution 2: Pipette 0.5ml of Dilution 1 into a test tube and then add 4.5ml of water.
- 5/ Add 1ml of the potassium dihydrogen phosphate standard to a beaker and add 99ml of water. (1/100 standard)
- 6/ Dilute 1ml of this solution with 4mls of water in a test tube.
- 7/ Dilute 1ml of this solution with 4mls of water in a test tube.
- 8/ In a separate test tube rack add the following to separate test tubes: 0.5ml of laundry powder solution, 0.5ml of Dilution 1, 0.5ml of Dilution 2, 0.5ml of diluted standard 1/100, 0.5ml of diluted standard 1/500, 0.5ml of diluted standard 1/2500 and finally 0.5ml of water.
- 9/ Add 4.5ml of molybdate solution to the tubes in step 8.
- 10/ After 15minutes the tubes may be compared or read in a colorimeter.

Many detergents contain phosphates from their method of manufacture. These phosphates are washed down sinks and drains to eventually enter creeks and rivers. Phosphate powerfully stimulates plant growth and in particular the growth of algae. Blooms of poisonous blue/green algae are increasingly affecting waterways during the warm summer months.

Result: The test tubes develop a blue colour, the intensity of which is dependent on the concentration of phosphate in the samples.

Conclusion: It is possible to assess the phosphate contained in the laundry powder by comparing the colour developed in the diluted samples with the colour in the range of standard phosphate dilutions. If a colorimeter was used, absorbance can be graphed against standard concentration allowing a more accurate figure for the samples to be read from the graph.

Risk Level: Low Hazard: The acids, molybdate and phosphate pose no hazard at these concentrations (the reagents having been prepared by a laboratory assistant). The experiment itself requires good pipette technique and high levels of concentration from the students. The value of this experiment lies in its ability to accurately determine very low concentrations of a contaminant.

STUDENT: _____

229

Smoke Bomb

Aim: To make a chemical reaction which produces a lot of smoke.

Equipment

Cotton rags
Sodium Nitrate
Tin can, 300ml approx
Beaker, 250ml
Retort stand
Boss head and clamp
Screw driver and hammer
Heat mat

Procedure

NOT TO BE CARRIED OUT BY STUDENTS

Results: _____

Conclusion: _____

Smoke Bomb

Topics: Chemical Rns

Aim: To make a chemical reaction which produces a lot of smoke.

Equipment

Cotton rags
Sodium Nitrate
Tin can, 300ml approx
Beaker, 250ml
Retort stand
Boss head and clamp
Screw driver and hammer
Heat mat

Procedure

- 1/ Add 100ml of water to the beaker.
- 2/ Dissolve 30g of sodium nitrate in the water.
- 3/ Tear the cotton cloth into long strips about 2cm wide
- 4/ Dip the cotton strips into the solution, squeeze dry and hang on a retort stand in the fume hood.
- 5/ Allow the strips to dry for two days.
- 6/ Punch breather holes in the sides of the can with the screw driver.
- 7/ Stuff the cloth strips into the can leaving 3cm of the last strip protruding.
- 8/ Take the can to a large open area eg. the oval.
- 9/ Stand the can on a heat mat.
- 10/ Ensure the prevailing breeze will not take the smoke into the school or nearby houses.
- 11/ Light the protruding cloth wick with a match and stand back.

Result: Copious volumes of smoke issue from the can for about a minute.

Conclusion: Sodium nitrate chemically oxidises cellulose in cotton however the oxidation is incomplete. Complete combustion of an organic substance produces invisible gases such as carbon dioxide and water vapour. Incomplete combustion produces small particles and droplets of organic compounds which refract and reflect light causing an opaque haze in the air.

Risk Level: EXTREMELY HAZARDOUS: Teacher demonstration only. The Principal and other staff should be warned. The smoke contains toxic organic compounds and should not be inhaled. Sodium nitrate is a powerful oxidising agent and can be explosive in mixtures containing reducing agents such as cyanides, aluminium powder or sodium thiosulfate. Sodium nitrate is harmful if ingested in quantity.

STUDENT: _____

230

Consumer Science

Aim: To make use of the technical information provided on product labels.

Equipment

Various empty product containers eg.

Caustic Soda

Window Cleaner

Oven Cleaner

Baking powder

Soldering Flux

Hair Shampoo

Methylated Spirits

Bleach

Floor cleaner

Laundry detergent

Mayonnaise

Lemonade

Procedure

Examine the label of a product. In the space below record the name of the product, whether it carries a warning, the active ingredient or the main ingredient, and one additive if any.

Repeat this procedure for another four products.

1/ _____

2/ _____

3/ _____

4/ _____

Results: _____

Conclusion: _____

Consumer Science

Topics: Consumer Science

Aim: To make use of the technical information provided on product labels.

Equipment

Various empty product containers eg.

Caustic Soda

Window Cleaner

Oven Cleaner

Baking powder

Soldering Flux

Hair Shampoo

Methylated Spirits

Bleach

Floor cleaner

Laundry detergent

Mayonnaise

Lemonade

Procedure

1/Give the students notes on or have them research the following terms:

Caustic, Corrosive, Toxic, Inflammable, Volatile, Wetting agent, Brightening Agent, Antioxidant, Softening Agent, Preservative, Emulsifier, Active Ingredient.

2/ Explain the labelling laws. Ingredients being listed in order of concentration. Distribute a copy of the additive number codes.

Result: Product labels contain much useful information besides its name and usage directions. Some carry directions in case of poisoning.

Conclusion: Warnings on products must be read for their safe use. The active ingredient quantity can be used to compare brands before you buy. Some products contain similar ingredients eg bleach and window cleaner, caustic soda and oven cleaner, yet the prices vary enormously. Some products contain additives which harm people with allergies.

Risk Level: Low Hazard

STUDENT: _____

231

Cosmetics 1

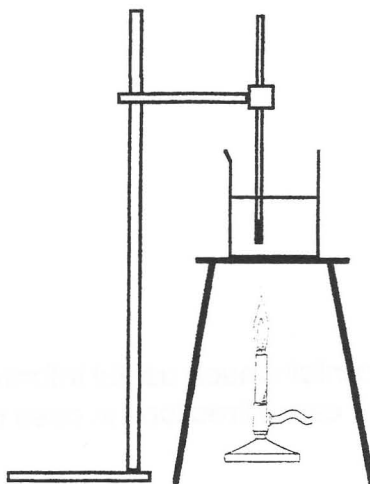
Aim: To make a simple cold cream.

Equipment

Candle wax
Liquid Paraffin
Borax
Thermometer, 0-100
Beakers, two, 250ml
Bunsen
Glass stirring rod
Tripod and gauze square
Flower essence
Retort stand
Boss head and clamp
Measuring Cylinder, 100ml
Balance and filter paper

Procedure

- 1/ Measure 16g of candle wax and 50ml of liquid paraffin into a beaker.
- 2/ Support the beaker on a tripod and gauze square.
- 3/ Arrange the retort stand, boss head and clamp to support a thermometer in the paraffin.
- 4/ Use a Bunsen to carefully heat the wax and paraffin to 75 degrees Celsius. **Paraffin and candle wax are flammable.**
- 5/ Add 33mls of water and 1g of borax to a second beaker. Heat this mixture to 75 degrees, stirring until the borax is dissolved.
- 6/ Pour the borax solution into paraffin/wax mixture.
- 7/ Stir the combined mixtures continuously as they cool until a cream forms.
- 8/ Add two drops of flower essence and stir for another few minutes.
- 9/ Transfer the cream to a small jar.
- 10/ Clean the beakers, stirring rod and thermometer in hot water and detergent.



Results: _____

Conclusion: _____

Cosmetics 1

Topics: Consumer Science Organic Chem

Aim: To make a simple cold cream.

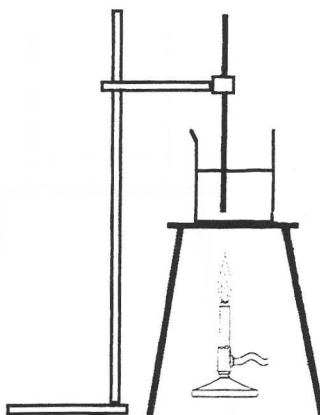
Equipment

Candle wax
Liquid Paraffin
Borax
Thermometer, 0-100
Beakers, two, 250ml
Bunsen
Glass stirring rod
Tripod and gauze square
Flower essence
Retort stand
Boss head and clamp
Measuring Cylinder, 100ml
Balance and filter paper

Students should bring their own clean jar to receive the cold cream.

Procedure

- 1/ Measure 16g of candle wax and 50ml of liquid paraffin into a beaker.
- 2/ Support the beaker on a tripod and gauze square.
- 3/ Arrange the retort stand, boss head and clamp to support a thermometer in the paraffin.
- 4/ Use a Bunsen to carefully heat the wax and paraffin to 75 degrees Celsius. **Paraffin and candle wax are flammable.**
- 5/ Add 33mls of water and 1g of borax to a second beaker. Heat this mixture to 75 degrees, stirring until the borax is dissolved.
- 6/ Pour the borax solution into paraffin/wax mixture.
- 7/ Stir the combined mixtures continuously as they cool until a cream forms.
- 8/ Add two drops of flower essence and stir for another few minutes.
- 9/ Transfer the cream to a small jar.
- 10/ Clean the beakers, stirring rod and thermometer in hot water and detergent.



Result: A soft, scented, white cream is produced.

Conclusion: Cold cream is an emulsion of water, candle wax and paraffin oil which do not normally form a stable mixture.

Risk Level: Moderate Hazard: Candle wax and paraffin are both flammable.

STUDENT: _____

232

Cosmetics 2

Aim: To produce a type of mascara.

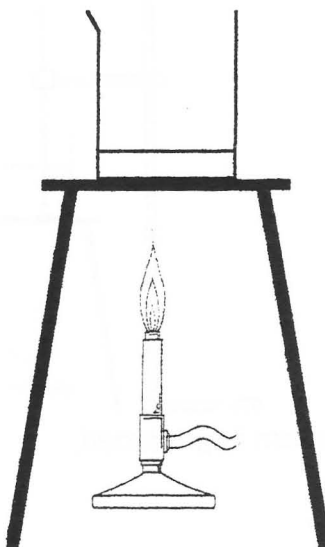
Equipment

Candle wax
Beeswax
Triethanolamine
Lanolin
Charcoal Powder
Tripod and gauze square
Bunsen
Glass stirring rod
Beaker, 100ml
Watch glass
Balance and filter paper

Procedure

- 1/ Add the following to a beaker:

| | |
|-----------------|------|
| Candle wax | 4g |
| Beeswax | 1.2g |
| Charcoal Powder | 1g |
- 2/ Weigh 0.8g of lanolin onto a tared watch glass. Add 3g of Triethanolamine then pour both into the beaker.
- 3/ Gently heat the beaker on a tripod until the components just melt.
- 4/ Stir slowly with a glass rod as the mixture cools until a thick cream forms.
- 5/ Transfer the mixture to a small jar.
- 6/ Clean the beaker, watch glass and stirring rod with hot water and detergent.



Results: _____

Conclusion: _____

Cosmetics 2

Topics: Consumer Science Organic Chem

Aim: To produce a type of mascara.

Equipment

Candle wax
Beeswax
Triethanolamine
Lanolin
Charcoal Powder
Tripod and gauze square
Bunsen
Glass stirring rod
Beaker, 100ml
Watch glass
Balance and filter paper

Procedure

1/ Add the following to a beaker:

| | |
|-----------------|------|
| Candle wax | 4g |
| Beeswax | 1.2g |
| Charcoal Powder | 1g |

2/ Weigh 0.8g of lanolin onto a tared watch glass. Add 3g of Triethanolamine then pour both into the beaker.

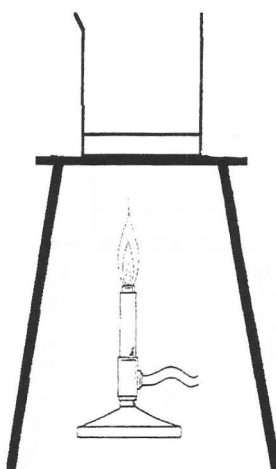
3/ Gently heat the beaker on a tripod until the components just melt.

4/ Stir slowly with a glass rod as the mixture cools until a thick cream forms.

5/ Transfer the mixture to a small jar.

6/ Clean the beaker, watch glass and stirring rod with hot water and detergent.

Students should bring their own small jar to receive the mascara.



Result: A thick, dark cream is formed.

Conclusion: Mascara is a suspension of dark charcoal powder in a semi-liquid mixture non-polar substances. Non polar liquids do not mix with water and so not easily washed away.

Risk Level: Mild Hazard: Charcoal powder is not hazardous but is often messy. Students should be cautioned when heating small quantities.

STUDENT: _____

233

Cosmetics 3

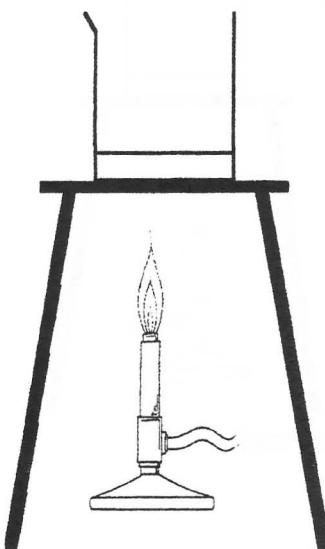
Aim: To produce a simple form of eye shadow or sun screen.

Equipment

Liquid Paraffin
Beeswax
Stearyl Alcohol
Zinc Oxide
Beaker, 100ml
Tripod and gauze square
Bunsen
Glass Stirring rod
Measuring Cylinder, 10ml

Procedure

- 1/ Measure 5ml of liquid paraffin and 2ml of stearyl alcohol into the beaker.
- 2/ Add 1.5g of beeswax.
- 3/ Heat the beaker gently until the wax melts.
- 4/ Stir a spatula of zinc oxide into the mixture.
- 5/ Transfer the mixture into a small container.
- 6/ Clean the beaker, stirring rod and measuring cylinder in hot water and detergent.



Results: _____

Conclusion: _____

Cosmetics 3

Topics: Consumer Science Organic Chem

Aim: To produce a simple form of eye shadow or sun screen.

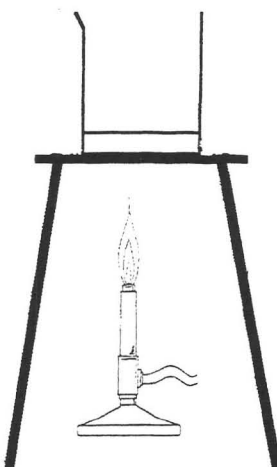
Equipment

Liquid Paraffin
Beeswax
Stearyl Alcohol
Zinc Oxide
Beaker, 100ml
Tripod and gauze square
Bunsen
Glass Stirring rod
Measuring Cylinder, 10ml

Procedure

- 1/ Measure 5ml of liquid paraffin and 2ml of stearyl alcohol into the beaker.
- 2/ Add 1.5g of beeswax.
- 3/ Heat the beaker gently until the wax melts.
- 4/ Stir a spatula of zinc oxide into the mixture.
- 5/ Transfer the mixture into a small container.
- 6/ Clean the beaker, stirring rod and measuring cylinder in hot water and detergent.

Students should bring their own small jar to receive the cream.



Result: A thick, white cream is produced.

Conclusion: Sun screen and eye shadow are similar products, both being a suspension in a semi-liquid mixture of non-polar compounds. Non polar liquids do not mix with water and so are not easily washed away. Zinc Oxide absorbs ultra violet light protecting skin covered by the cream. Body sunscreen uses titanium oxide in a lanolin base.

Risk Level: Moderate Hazard: Liquid paraffin and beeswax are both flammable. Students should be cautioned when heating small quantities.

STUDENT: _____

234

Laundry Detergents

Aim: To determine which of three common laundry detergents is best at removing stains in cold water.

Equipment

Laundry Detergents (3)
250ml beakers (4)
Plain cloth 30cm X 30cm
Stop watch
Thermometer
Stirring rod
Stains: tomato sauce
 beetroot juice
 muddy water
 sump oil
 black marker pen

Procedure

Devise a procedure to fairly compare the detergents.

Results: _____

Conclusion: _____

Laundry Detergents

Topics: Consumer Science Scientific Method

Aim: To determine which of three common laundry detergents is best at removing stains in cold water.

Equipment

- Laundry Detergents (3)
- 250ml beakers (4)
- Plain cloth 30cm X 30cm
- Stop watch
- Thermometer
- Stirring rod
- Stains: tomato sauce
 beetroot juice
 muddy water
 sump oil
 black marker pen

Procedure

- Suggested procedure:
- 1/ Tear the cloth into four strips
 - 2/ Mark each cloth strip with **equal** amounts of each stain.
 - 3/ Place five grams of each detergent into a separate beaker
 - 4/ Add 200mls of water to clean beaker. Measure and record the temperature of the water
 - 5/ Add 200mls of water to the beakers with detergent.
 - 6/ Stir for 10 seconds to dissolve the detergents.
 - 7/ Add a strip of stained cloth to each beaker and start the stop watch.
 - 8/ Stir each beaker once every ten seconds.
 - 9/ Remove the cloth strips after 5 minutes
 - 10/ Draw up a table to compare the effectiveness of each detergent in removing each stain.
 - 11/ Score each detergent according to a scale.
0- poor , 1- fair , 2- good, 3- very good,
 - 12/ Total the scores

| Beaker | Sauce | Oil | Pen | Beet | Mud | Total Score |
|--------|-------|-----|-----|------|-----|-------------|
| 1 | | | | | | |
| 2 | | | | | | |
| 3 | | | | | | |
| 4 | | | | | | |

Result: Non of the detergents were very effective on axle sump oil. Water alone is partly effective with mud and sauce. All the detergents were better than water alone, one detergent being slightly better than the others.

Conclusion: While the test is fair it does not compare the effectiveness of the detergents in warm water, with different materials or the harshness of the detergent on the material. Some detergents contain sodium hydroxide making them effective on oils in cold water but this chemical also damages clothes and skin.

Risk Level: Low Hazard.: Some laundry detergents contain sodium hydroxide which is damaging to skin and eyes.

STUDENT: _____

235

Window Cleaners

Aim: To produce and compare various window cleaning solutions.

Equipment

Hand sprayer
Dish washing detergent (in a
dropper bottle)
Ammonia ,10%, (in a
dropper bottle)
Methanol
Measuring Cylinder,100ml
Newspaper

Procedure

1/ Prepare the following three solutions:

Solution A

100ml water

10 drops detergent

Solution B

100ml water

10 drops detergent

20 drops ammonia

Solution C

50ml water

10 drops detergent

50ml methanol

2/ Place 100ml of water in the hand sprayer. Spray one laboratory window and wipe dry with newspaper.

3/ Replace the water in the sprayer with solution A and clean another laboratory window.

4/ Replace solution A with solution B and clean another window.

5/ Replace solution B with solution C and clean one more window.

6/ Which solution cleans windows best?

- Which cleaning agent was the control in this experiment.
- List four things wrong with the procedure as a method for comparing window cleaners.

Results: _____

Conclusion: _____

Window Cleaners

Topics: Consumer Science Scientific Method

Aim: To produce and compare various window cleaning solutions.

Equipment

Hand sprayer
Dish washing detergent (in a dropper bottle)
Ammonia, 10%, (in a dropper bottle)
Methanol
Measuring Cylinder, 100ml
Newspaper

Procedure

1/ Prepare the following three solutions:

Solution A

100ml water

10 drops detergent

Solution B

100ml water

10 drops detergent

20 drops ammonia

Solution C

50ml water

10 drops detergent

50ml methanol

2/ Place 100ml of water in the hand sprayer. Spray one laboratory window and wipe dry with newspaper.

3/ Replace the water in the sprayer with solution A and clean another laboratory window.

4/ Replace solution A with solution B and clean another window.

5/ Replace solution B with solution C and clean one more window.

6/ Which solution cleans windows best?

- Which cleaning agent was the control in this experiment.
- List four things wrong with the procedure as a method for comparing window cleaners.

-Water is the control liquid

Problems:

- The windows are not identically dirty
- The windows were not cleaned identically well
- Cleaness of windows is a "null result", that is success gives an absence of data.
- Cleaness of windows is subjective, that is it cannot be objectively measured
- There were no duplicates.

Result: Generally solution C gives the best result.

Conclusion: All the solutions will clean windows. It is not possible to determine which solution is best since judging clean windows is subjective and the windows were not identically dirty.

Risk Level: Low Hazard: Methanol is a flammable liquid. Ammonia is caustic affecting eyes and skin. Fumes from ammonia are pungent and can irritate eyes and lungs.

STUDENT: _____

236

Forensic Science

Aim: To analyse the variation of thumb prints in the class

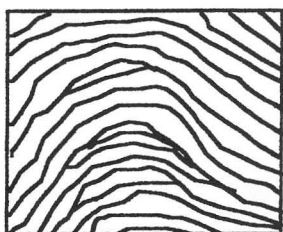
Equipment

Ink pad

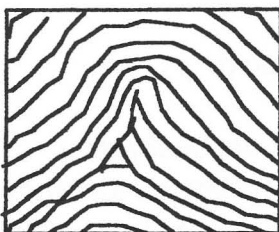
Paper squares, 10X10cm

Procedure

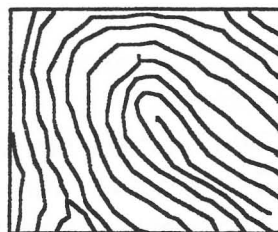
- 1/ Write your name on the back of a peice of paper.
- 2/ Make a clear print on the front of the paper by inking your right thumb and then rolling it on the paper from left to right.
- 3/ Identify the "type" of your thumb print from the diagrams.
How many students in the class have plain arches? _____
How many students in the class have tented arches? _____
How many students in the class have whorls? _____
How many students in the class have simple loops? _____
How many students in the class have compoun loops? _____
- 4/ Produce a bar graph in the space below which represents the proportions thumb print types in the class. Use a key.
- 5/ Divide into five goups depending on your thumb print type.
- 6/ Examine the prints and determine if any are identical.



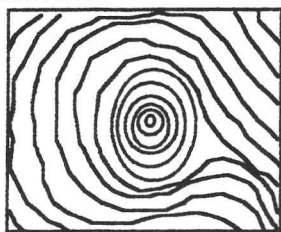
Plain Arch



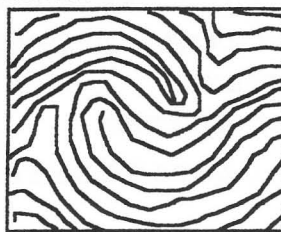
Tented Arch



Whorl



Simple Loop



Compound Loop

Results: _____

Conclusion: _____

Forensic Science

Topics: Scientific Method

Genetics

Aim: To analyse the variation of thumb prints in the class

Equipment

Ink pad

Paper squares, 10X10cm

Procedure

1/ Write your name on the back of a piece of paper.

2/ Make a clear thumb print on the front of the paper by inking your right thumb and then rolling it on the paper from left to right.

3/ Identify the "type" of your thumb print from the diagrams below.

How many students in the class have plain arches? _____

How many students in the class have tented arches? _____

How many students in the class have whorls? _____

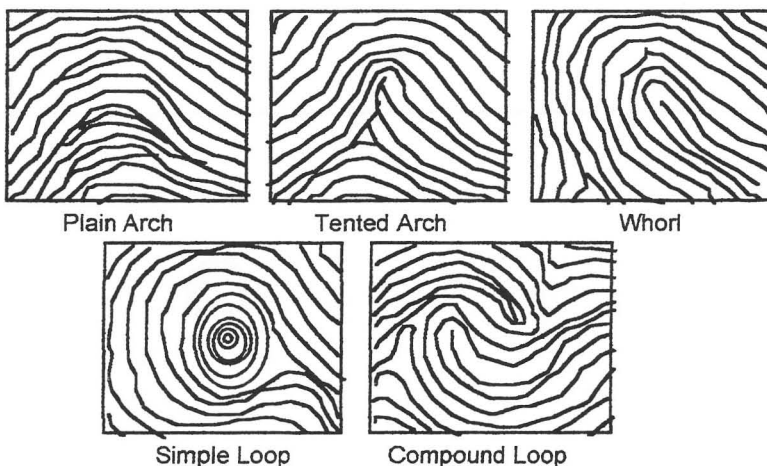
How many students in the class have simple loops? _____

How many students in the class have compound loops? _____

4/ Produce a bar graph in the space below which represents the proportions thumb print types in the class. Use a key.

5/ Divide into five groups depending on your thumb print type.

6/ Examine the prints and determine if any are identical.



Result: No two thumb prints in the class were identical (identical twins excepted).

Conclusion: While various types of thumb prints are apparent there is sufficient variation within types to say each print is unique.

Risk Level: Low Hazard:



STUDENT: _____

237

Blind Spot

Aim: To verify that human eyes have a blind spot

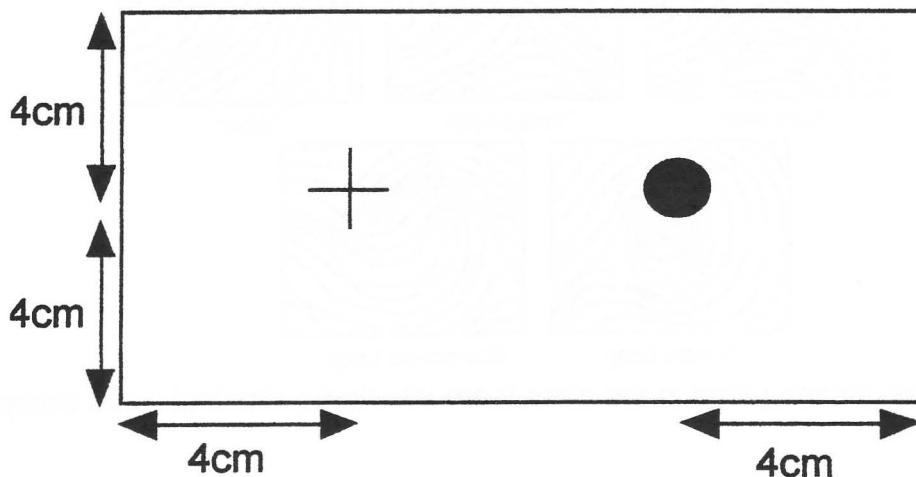
Equipment

Paper , 12cm X 8cm

Procedure

- 1/ Draw a 1cm black cross 4cm from the left , top, and bottom edges.
- 2/ Draw a 1cm black dot 4cm from the right, top and bottom edges.
- 3/ Hold your right eye closed with your right hand.
- 4/ Hold the paper in your left hand and about 30cm in front of your face.
- 5/ Concentrating on the black dot, slowly move the paper toward your face.

- Explain what happens to the black cross in terms of the physical structure of the human eye.
- Try to explain why this effect is not visible normally.



Results: _____

Conclusion: _____

Blind Spot

Topics: Coordination

Light

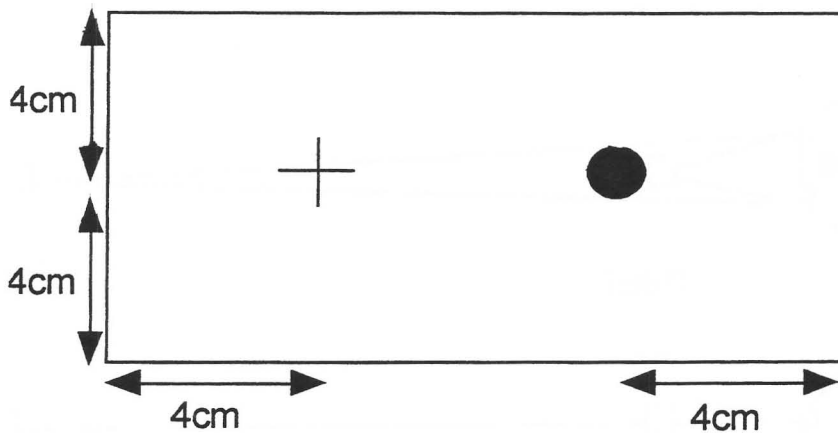
Aim: To verify that human eyes have a blind spot**Equipment**

Paper , 12cm X 8cm

Procedure

- 1/ Draw a 1cm black cross 4cm from the left , top, and bottom edges.
- 2/ Draw a 1cm black dot 4cm from the right, top and bottom edges.
- 3/ Hold your right eye closed with your right hand.
- 4/ Hold the paper in your left hand and about 30cm in front of your face.
- 5/ Concentrating on the black dot, slowly move the paper toward your face.

- Explain what happens to the black cross in terms of the physical structure of the human eye.
- Try to explain why this effect is not visible normally.

**Result:** The black dot disappears.

Conclusion: Where the optic nerve connects to the back of the eye there is a small circular gap in the light sensitive retina. An object focused on this "Bind Spot" in the retina is apparently invisible. Images from the eyes are processed by the brain which fills in details seen by one eye but not the other.

Risk Level: Low Hazard:

STUDENT: _____

238

Depth Perception

Aim: To investigate the role of binocular vision in perceiving distance.

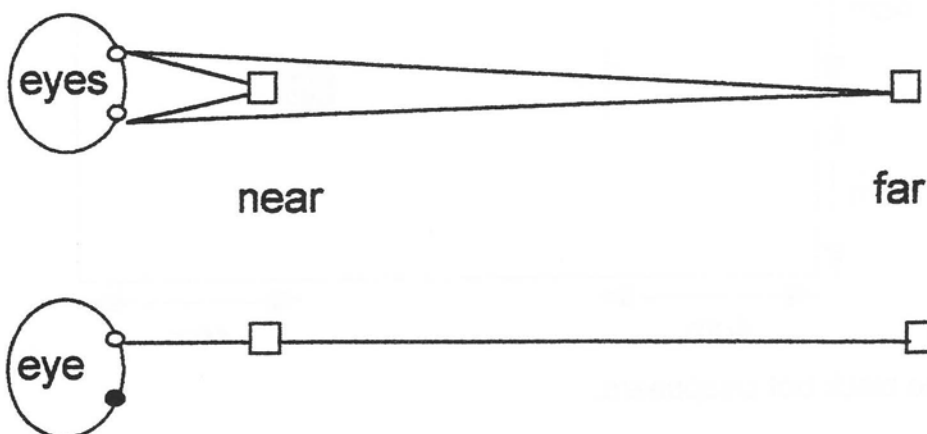
Equipment

None

Procedure

- 1/ Divide into pairs.
- 2/ One person, the subject, holds their right arm straight out to their side.
- 3/ The other person holds a ruler vertically about half a metre in front of the subject.
- 4/ The subject tries to swing their right hand in a smooth arc to touch the side of the ruler with their index finger.
- 5/ The experiment is repeated with the subject holding their left eye closed with their left hand.
- 6/ Swap roles of subject and experimenter then repeat the experiment.

- Examine the diagram below and then try to explain the result of your experiment.
- Why are one eyed people not allowed to drive a vehicle or pilot a plane.



Results: _____

Conclusion: _____

Depth Perception

Topics: Coordination

Light

Aim: To investigate the role of binocular vision in perceiving distance.

Equipment

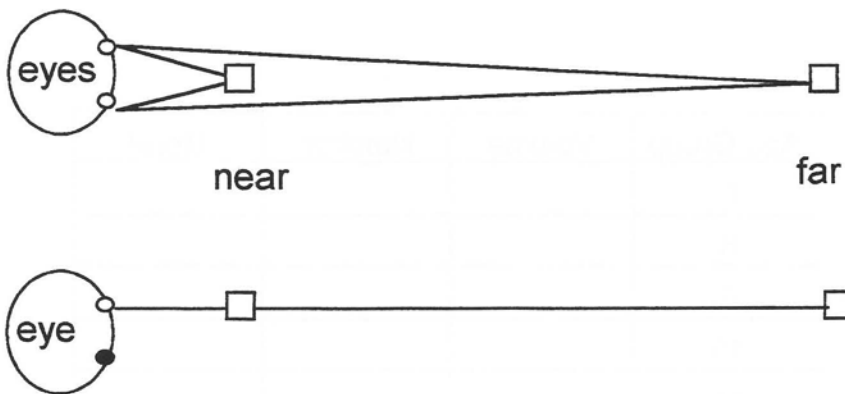
None

Procedure

- 1/ Divide into pairs.
- 2/ One person, the subject, holds their right arm straight out to their side.
- 3/ The other person holds a ruler vertically about half a metre in front of the subject.
- 4/ The subject tries to swing their right hand in a smooth arc to touch the side of the ruler with their index finger.
- 5/ The experiment is repeated with the subject holding their left eye closed with their left hand.
- 6/ Swap roles of subject and experimenter then repeat the experiment.

- Examine the diagram below and then try to explain the result of your experiment.

- Why are one eyed people not allowed to drive a vehicle or pilot a plane.



Result: With both eyes open the subject could easily touch the side of the ruler. With one eye closed the subject missed the ruler, their finger passing in front or behind by one or two centimetres.

Conclusion: When both eyes focus on a near object a triangle is formed with smaller base angles than for a distant object. The eyes **triangulate** the distance to the object by their angle relative to each other. One eye alone sees no difference in angle between near and far objects. One eyed people are not allowed to drive since they cannot accurately judge stopping distance.

Risk Level: Low hazard:

Learning

Aim: To observe Piaget comprehension stages in young students.

Equipment

Arrangement to visit a class of students aged 6 and a class aged 11.
Teams 1 and 2:
Measuring cylinder, 100ml
Beaker, 400ml
Teams 3 and 4:
match boxes, empty, seven
matches, seven
Teams 5 and 6:
Moral Conundrum

Procedure

Piaget was an educationalist who recognised particular cognitive (mental processing) stages in children.
1/ The teacher divides the class into six teams.
2/ Teams 1 and 2 will each test a different group of three, 6 year old children for "Volume Cognition": Fill the beaker and measuring cylinder with water and ask which contains the most water.
3/ Teams 3 and 4 will test separate groups of three, 6 year old students for " Number Conservation" : The students are shown each match box contains one match. The matches are removed and placed in a pile while the match boxes are placed in a separate pile. The students are asked "Which pile has more things in it?"
4/ Teams 4 and 6 ask test separate groups of six year olds for "Moral Cognition" with a question: "If one boy breaks six dishes when washing up and another boy breaks one dish while being naughty, which boy is baddest?"
5/ Teams rotate twice between the test groups.
6/ Rotate tasks and then test 11 year old students.

| Age Group | Volume | Number | Moral |
|-----------|--------|--------|-------|
| 6 | | | |
| 6 | | | |
| 6 | | | |
| 11 | | | |
| 11 | | | |
| 11 | | | |

Use a tick for each correct reponse and a cross for each incorrect response.

Results: _____

Conclusion: _____

Learning

Topics: Coordination

Aim: To observe Piaget comprehension stages in young students.

Equipment

Arrangement to visit a class of students aged 6 and a class aged 11.
Teams 1 and 2:
Measuring cylinder, 100ml
Beaker, 400ml
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4/ Teams 4 and 6 ask test separate groups of six year olds for “Moral Cognition” with a question: “If one boy breaks six dishes when washing up and another boy breaks one dish while being naughty, which boy is worse?”
5/ Teams rotate twice between the test groups.
6/ Rotate tasks and then test 11 year old students.

Test your students: One boy tells another “ Break that window or I will break your face”. Who is to blame? (eg. war crimes).

| Age Group | Volume | Number | Moral |
|-----------|--------|--------|-------|
| 6 | | | |
| 6 | | | |
| 6 | | | |
| 11 | | | |
| 11 | | | |
| 11 | | | |

Use a tick for each correct reponse and a cross for each incorrect response.

Result: Six year old children will mostly fail all three tests unless guessing. Eleven year old children have volume cognition and number conservation but are still developing moral cognition.

Conclusion: Young children can estimate only one dimension at a time, height being dominant. Size estimate overrides the abstract concept of number. Moral cognition is difficult and is overridden by number sense. Cognition of space, number, and rightness are achieved at different stages of mental development.

Risk Level: Low hazard: Students must be primed to be sensitive to the younger students avoiding any comment on their answers, especially words such as dumb, stupid etc.

STUDENT: _____

240

Antigravity

Aim: To demonstrate that air pressure can support a column of water.

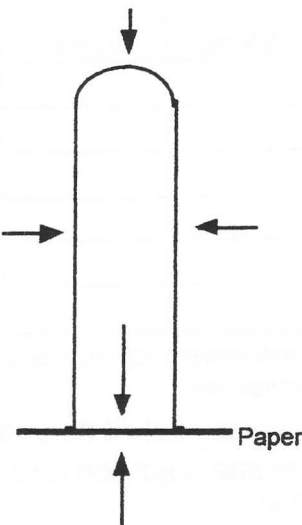
Equipment

Test tube
Writing paper, 5cmX5cm

Procedure

- 1/ Fill a test tube with water to the very top.
- 2/ Slide the paper over the top of the test tube, carefully trying to exclude any air bubbles between the water and the paper.
- 3/ Slowly invert the test tube.

- On the diagram below arrows have been drawn indicating the forces in action. Label the arrows (the same label may be used more than once).



Results: _____

Conclusion: _____

Antigravity

Topics: Density & Pressure

Aim: To demonstrate that air pressure can support a column of water.

Equipment

Test tube

Writing paper, 5cmX5cm

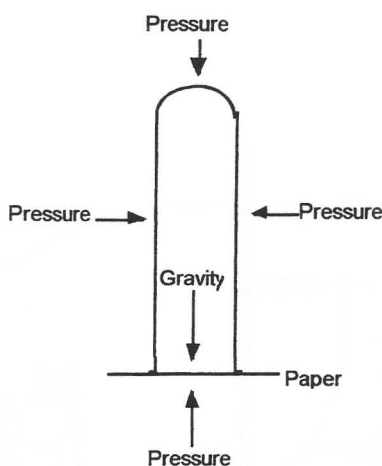
Procedure

1/ Fill a test tube with water to the very top.

2/ Slide the paper over the top of the test tube, carefully trying to exclude any air bubbles between the water and the paper.

3/ Slowly invert the test tube.

- On the diagram below arrows have been drawn indicating the forces in action. Label the arrows (the same label may be used more than once).



Result: The water does not fall out of the test tube.

Conclusion: On three sides air pressure is acting on the test tube. On the final side, air pressure is acting on the water surface. The air pressure acting on the surface is sufficient to counter the gravitational force on the water thus preventing its fall. Without paper the water is free to flow out by deforming its surface.

Risk Level: Low Hazard:

STUDENT: _____

241

Fire Alarm

Aim: To make an alarm that automatically turns on in the presence of fire.

Equipment

Power Supply, 0-12V

Electric Bell

Bimetallic strip

Connecting Wires, three

Retort stands, two

Boss head & clamps, two

Bunsen

Alligator clip

Procedure

1/ Using connecting wire link a DC terminal of the power supply to a terminal of the electric bell.

2/ Using an alligator clip and connecting wire, link the bimetallic strip to the other terminal on the bell.

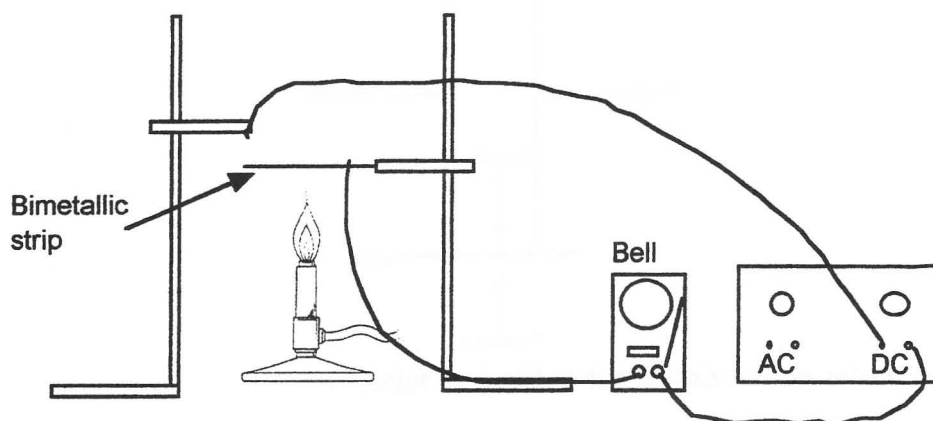
3/ Using a retort stand and clamp, support the bimetallic strip about 5cm above the Bunsen. The strip should be parallel to the desk, broad side facing the Bunsen.

4/ Using the second retort stand, support one end of the final connecting lead about 2cm above the end of the bimetallic strip.

5/ Link the free end of the last connecting lead to the other DC terminal of the power supply.

6/ Set the power supply to 12 DC and light the Bunsen.

7/ If the bimetallic strip begins to bend downward, remove the Bunsen and turn over the bimetallic strip.



Results: _____

Conclusion: _____

Fire Alarm

Topics: electricity

Kinetic Theory

Matter

Aim: To make an alarm that automatically turns on in the presence of fire.

Equipment

Power Supply, 0-12V

Electric Bell

Bimetallic strip

Connecting Wires, three

Retort stands, two

Boss head & clamps, two

Bunsen

Alligator clip

Procedure

1/ Using connecting wire link a DC terminal of the power supply to a terminal of the electric bell.

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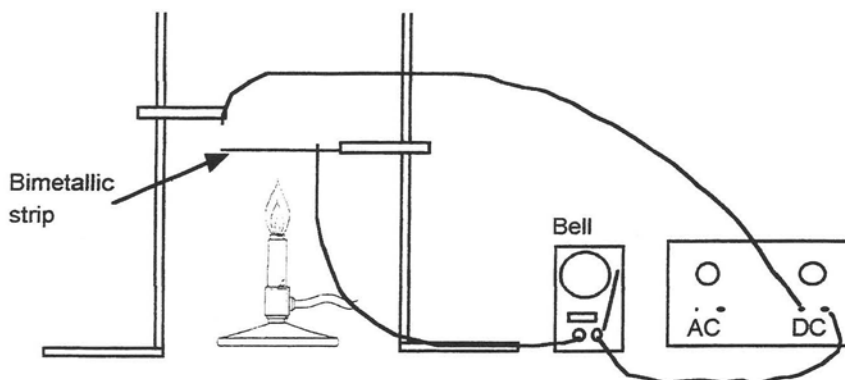
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5/ Link the free end of the last connecting lead to the other DC terminal of the power supply.

6/ Set the power supply to 12 DC and light the Bunsen.

7/ If the bimetallic strip begins to bend downward, remove the Bunsen and turn over the bimetallic strip.



Result: The bell rings when the Bunsen flame heats the bimetallic strip.

Conclusion: The bimetallic strip has two different metal layers that expand at different rates. When heated on one side the bimetallic strip bends until its tip touches the connecting lead above and so completes the circuit. The alarm rings in the presence of strong heat.

Risk Level: Low Hazard:

STUDENT: _____

242

Fuel Cell

Aim: To make and test a simple fuel cell.

Equipment

Test Tubes, 4ml, two
Carbon Rods, two
Retort Stand
Ring Clamp
Alligator clips, two
Silicone glue
Connecting wires, two
Power supply, 0 -12V
Lamp, 2V
Polystyrene cup
Sulfuric Acid, 0.1M, 0.5%

Procedure

- Draw the apparatus set up by the teacher.
- If one tube contains oxygen and the other hydrogen write the half cell reactions you would expect to be taking place.
- Write the overall equation for the reaction.
- How is a fuel cell different from a battery (think in terms of supply of the two gasses).

Results: _____

Conclusion: _____

Fuel Cell

Topics: Electricity

Ions

Aim: To make and test a simple fuel cell.

Equipment

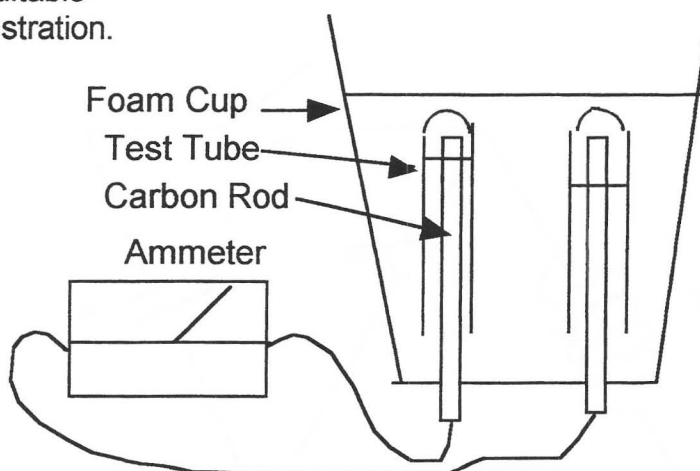
Test Tubes, 4ml, two
Carbon Rods, two
Retort Stand
Ring Clamp
Alligator clips, two
Silicone glue
Connecting wires, two
Power supply, 0 -12V
Ammeter 0-500mA
Polystyrene cup
Sulfuric Acid, 0.1M, 0.5%

Procedure

A fuel cell is a device which produces an electric current from ion reactions similar to a battery (galvanic cell) however the reactants can be continuously supplied to **fuel** the reaction.

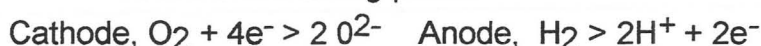
- 1/ Make two holes in the base of the foam cup.
- 2/ Place the cup in the ring clamp on the retort stand.
- 3/ Insert the carbon rods into the holes so they deeply protrude into the cup.
- 4/ Seal the carbon rods in place with silicone glue.
- 5/ When the glue is set, three quarter fill the cup with dilute acid so the electrodes are covered.
- 6/ Fill the two semi-micro test tubes with dilute acid then invert

The construction time requirements make this experiment more suitable as a teacher demonstration.



Result: A strong electric current was produced by the fuel cell.

Conclusion: Half cell reactions are taking place:



In a battery the reactions cease when the reactants have been consumed. In a fuel cell the reactants can be continually supplied maintaining the current.

Risk Level: Low Hazard:

STUDENT: _____

243

Physics Prac

Aim: To test practical skills involved in electrical circuits.

Equipment

Breadboard, plastic
Battery, 9V
Phillips screw driver
Screws and washers, 4
Battery clip
Resistor,
(Brown/Black/Black)
Solenoid
Switch
Light emitting diode
Multimetre
Connecting wires, 2 , 10cm

Procedure

1/ From the circuit diagram below, name the components labelled A,B,C,D,E.

2/ On the circuit diagram draw the following:

- The positive (+) and negative (-) sides of the battery.
- The colour of the wires leading to the battery.
- The direction of electron flow (e^-)
- The direction of current flow (I)
- The direction of magnetic field of the solenoid (B)
- A meter measuring voltage across C
- A meter measuring current flow in C

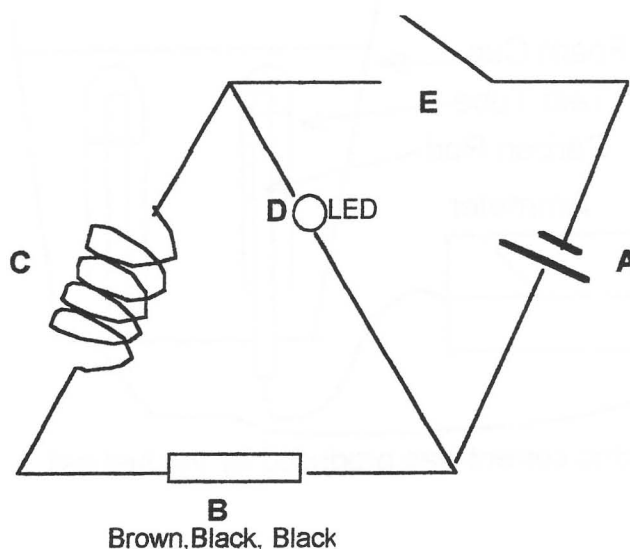
3/ Determine whether the following component pairs are in series or parallel:

- Components C and D.
- Components C and B .
- Components B and D.
- Components E and B .

4/ BUILD THE CIRCUIT:

8/ Using a multimetre, measure the following:

- The resistance of B.
- The Voltage Drop across B.
- The current flowing through C.



Results: _____

Conclusion: _____

Topics: Electricity

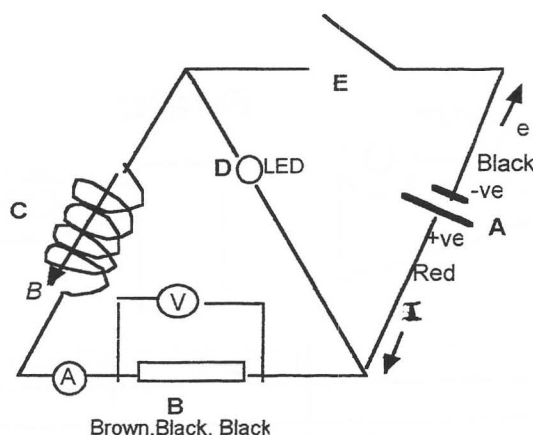
Aim: To test practical skills involved in electrical circuits.

Equipment

Breadboard, plastic
 Battery, 9V
 Phillips screw driver
 Screws and washers, 4
 Battery clip
 Resistor,
 (Brown/Black/Black)
 Solenoid
 Switch
 Light emitting diode
 Multimetre
 Connecting wires, 2 , 10cm

Procedure

- 1/ From the circuit diagram below, name the components labelled A, B, C, D, E.
- 2/ On the circuit diagram draw the following:
 - a) The positive (+) and negative (-) sides of the battery.
 - b) The colour of the wires leading to the battery.
 - c) The direction of electron flow (e^-)
 - d) The direction of current flow (I)
 - e) The direction of magnetic field of the solenoid (B)
 - f) A meter measuring voltage across C
 - g) A meter measuring current flow in C
- 3/ Determine whether the following component pairs are in series or parallel:
 - a) Components C and D. b) Components C and B .
 - c) Components B and D. d) Components E and B .
- 4/ BUILD THE CIRCUIT:
- 8/ Using a multimeter, measure the following:
 - a) The resistance of B. b) The Voltage Drop across B. c) The current flowing through C.



Result: A= Battery, B= Resistor, C= solenoid, D= Light Emitting Diode, E= switch.
 C and D are in parallel. C and B are in series. B and D are in parallel. E and B are in series.

Conclusion: Resistance of B is about 10 ohms, Voltage drop across B is about 2V, Current flowing in C is about 200 mA.

Risk Level: Low Hazard: Only use fresh batteries and caution the students to leave the switch off when not making measurements.

244

Series and Parallel

Aim: To examine the differences between series and parallel circuits.

Equipment

Power Supply, 0-12V, DC
Connecting wires, five
Lamp bases, two

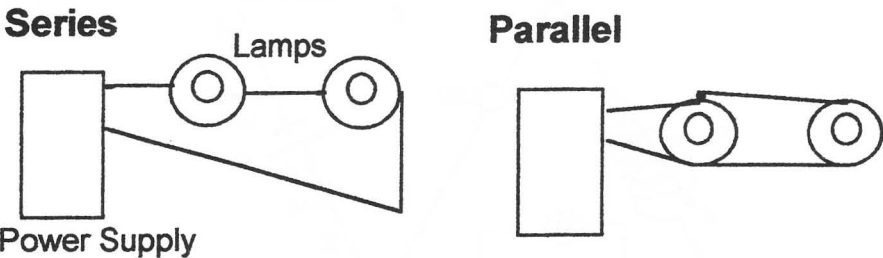
Optional: Multimeter to
record voltage and current
across each lamp.

Procedure

- 1/ Connect the lamps and power supply in series as shown in the diagram below.
- 2/ Set the power supply to 12V and record whether the lamps glow brightly, dimly or not at all.
- 3/ Unscrew one lamp and record the effect on the other lamp.
- 4/ Rearrange the connecting wires so the lamps are in parallel.
- 5/ Record how the lamps are glowing.
- 6/ Unscrew one lamp and record the effect on the other lamp.

- What is a disadvantage of series circuits?
- What is an advantage of series circuits?

Option: Measure the voltage and current across each lamp when in series and when in parallel. Measure the resistance in each circuit.



| | Lamp 1 | Lamp2 |
|---------------------------|--------|-------|
| Series | | |
| Series (one unscrewed) | | |
| Parallel | | |
| Parallel (one unscrewed) | | |

Results: _____

Conclusion: _____

Series and Parallel

Topics: Electricity

Aim: To examine the differences between series and parallel circuits.

Equipment

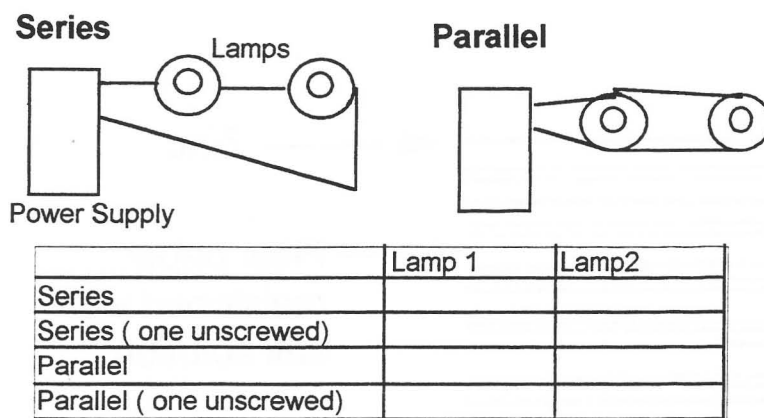
Power Supply, 0-12V, DC
Connecting wires, five
Lamp bases, two

Optional: Multimeter to record voltage and current across each lamp.

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 - 4/ Rearrange the connecting wires so the lamps are in parallel.
 - 5/ Record how the lamps are glowing.
 - 6/ Unscrew one lamp and record the effect on the other lamp.
- What is a disadvantage of series circuits?
 - What is an advantage of series circuits?

Option: Measure the voltage and current across each lamp when in series and when in parallel. Measure the resistance in each circuit.



Result: The lamps in series glowed dimly and when one globe was unscrewed the other went out. In parallel the lamps glowed brightly and when one lamp was unscrewed the other only dimmed slightly.

Conclusion: Series circuits require less wiring however if one component fails the whole circuit fails. In parallel circuits the current travels in each parallel path so when one component fails, current still flows to the other components. (Voltage across parallel components is higher, the current flow is higher and the total resistance is lower.)

Risk Level: Low Hazard:

STUDENT: _____

245

Shock Stack

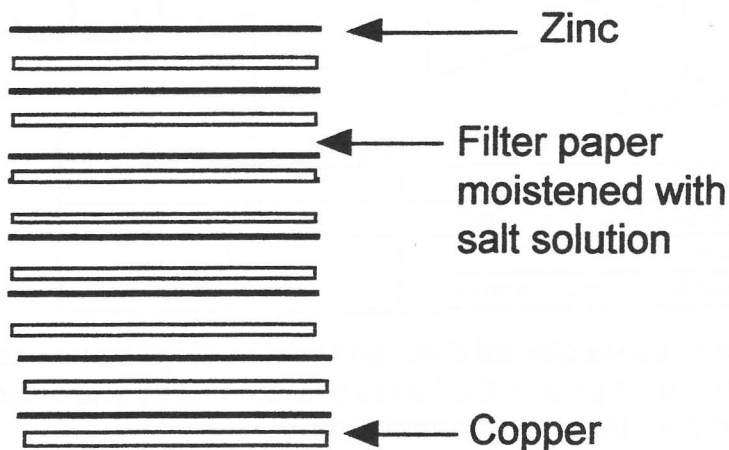
Aim: To create a simple, compact power source

Equipment

Copper sheet, 2X2cm, eight
Filter Paper, 2X2cm, twenty
Zinc sheet (galvanised iron)
, 2X2cm, eight
Sodium Chloride
Beaker, 250ml
Dropper bottle
Balance, 0.1g sensitivity
connecting wire
Measuring cylinder, 100ml
stirring rod

Procedure

- 1/ Weigh 3g of sodium chloride into a beaker.
- 2/ Add 100ml of water and dissolve the salt.
- 3/ Fill the dropper bottle with salt solution
- 4/ Place a copper square on the bench.
- 5/ Place a filter paper square on the copper square.
- 6/ Add two drops of salt solution to the filter paper.
- 7/ Place a zinc square on the moist paper.
- 8/ Place another piece of filter paper on the stack and moisten it with salt solution.
- 9/ Continue building the stack with alternate pieces of copper and zinc sheet separated by moistened filter paper.
- 10/ Turn the stack on its side holding it together with a finger of one hand on the copper base and a finger of the opposite hand on the zinc top.
- 11/ Do you feel anything?



Results: _____

Conclusion: _____

Shock Stack

Topics: Electricity

Ions

Aim: To create a simple, compact power source

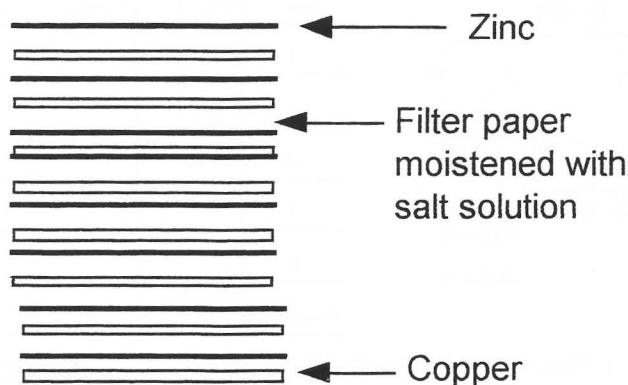
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Filter Paper, 2X2cm, twenty
Zinc sheet (galvanised iron)
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Sodium Chloride
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- 7/ Place a zinc square on the moist paper.
- 8/ Place another piece of filter paper on the stack and moisten it with salt solution.
- 9/ Continue building the stack with alternate pieces of copper and zinc sheet separated by moistened filter paper.
- 10/ Turn the stack on its side holding it together with a finger of one hand on the copper base and a finger of the opposite hand on the zinc top.
- 11/ Do you feel anything?

The metal sheets should be cleaned beforehand by brief immersion in 4M sulfuric acid and rinsing in water.



Result: A slight tingling is felt holding the stack ends with the fingers of opposite hands.

Conclusion: An voltaic cell (battery) is created by two dissimilar metals linked by a salt solution. Creating a stack of alternating metals and salt solution increases the power (voltage) of the combined voltaic cells.

Risk Level: Mild Hazard: No students with bionic aids should attempt this experiment.

Aim: To determine the power efficiency of an electric motor.

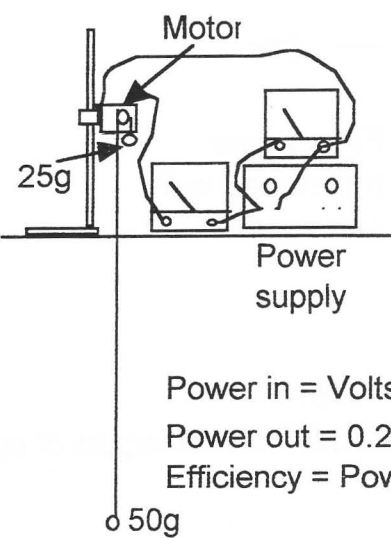
Equipment

- Retort Stand
- Clamp and boss head
- Power supply, 0-12V DC
- Electric Motor, 6V DC
- Connecting wires, five
- Multimeter (or Voltmeter 0-10V and Ammeter 0-5A)
- Stop watch
- Metre rule
- String (1.2m)
- Masses, 50g and 25g

Procedure

- 1/ Clamp the electric motor on the retort stand at the edge of the bench and one metre from the floor.
- 2/ Tie a 50g mass on one end of the string and the 25g mass at the other end.
- 3/ Drape the string over the motor spindle so the 50g mass is on the floor and the 25g mass one metre above.
- 4/ Connect the motor to the ammeter and the positive ammeter terminal to the positive DC terminal of the power supply.
- 5/ Link the negative DC power supply terminal back to the motor. Connect the voltmeter across the power terminals.
- 6/ Set the power supply to 6V. When the power is turned on record how long it takes for the 50g mass to be raised 1 metre as well as the voltage and current flowing.
- 7/ Repeat step 6 four times.

Power = work done / second, Electric Power = Volts X Amps
Mechanical Work : PE gained = mgh = 0.245 Watts,
Plus KE gained = $\frac{1}{2}mv^2 = 0.15 / t^2$



| Trial | Volts | Amps | Time |
|---------|-------|------|------|
| 1 | | | |
| 2 | | | |
| 3 | | | |
| 4 | | | |
| 5 | | | |
| Average | | | |

Power in = Volts X Amps = _____ Watts
Power out = 0.245 + $(0.15 / t^2)$ = _____ Watts
Efficiency = Power out / power in X 100 = _____ %.

Results: _____

Conclusion: _____

Topics: Electromagnetism

Energy

Machines

Aim: To determine the power efficiency of an electric motor.**Equipment**

Retort Stand
 Clamp and boss head
 Power supply, 0-12V DC
 Electric Motor, 6V DC
 Connecting wires, five
 Multimeter (or Voltmeter 0-10V and Ammeter 0-5A)
 Stop watch
 Metre rule
 String (1.2m)
 Masses, 50g and 25g

Procedure

1/ Clamp the electric motor on the retort stand at the edge of the bench and one metre from the floor.
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 7/ Repeat step 6 four times.
 Power = work done / second, Electric Power = Volts X Amps
 Mechanical Work : PE gained = $mgh = 0.245 \text{ Watts}$,
 Plus KE gained = $1/2mv^2 = 0.15 / t^2$

$$\text{KE gained} = 1/2 \times 0.075 \times v^2$$

$$v^2 = u^2 + 2as = 2a$$

$$s = ut + 1/2at^2 = 1/2at^2$$

$$a = 2/t^2$$

$$v^2 = 4/t^2$$

$$\text{KE} = 0.0375 \times 4 / t^2 = 0.15 / t^2$$

Result: Efficiencies as low as 3% are to be expected.

Conclusion: Many devices have poor power efficiencies. Most of the power loss occurs as heat. Laboratory model electric motors are designed so the internal workings can be seen and so the outer magnet windings are restricted to one side, guaranteeing a poor efficiency.

Risk Level: Low Hazard:

STUDENT: _____

247

Electric Wind

Aim: To demonstrate that a flame contains charged particles.

Equipment

Van De Graaf Generator

Connecting Wire

Bunsen

Wooden test tube peg

Procedure

- Place a connecting wire in the dome plug of the Van De Graaf generator.
- Start the generator
- Light a Bunsen within reach of the connecting wire and adjust to a yellow flame.
- Use the test tube peg to grasp the free end of the connecting wire.
- Bring the connecting wire near the side of the Bunsen flame.
- Draw the apparatus and result in the space below.

Hint: Van De Graaf generators produce large positive charges.

Results: _____

Conclusion: _____

Electric Wind

Topics: Electrostatics

Ions

Aim: To demonstrate that a flame contains charged particles.

Equipment

Van De Graaf Generator

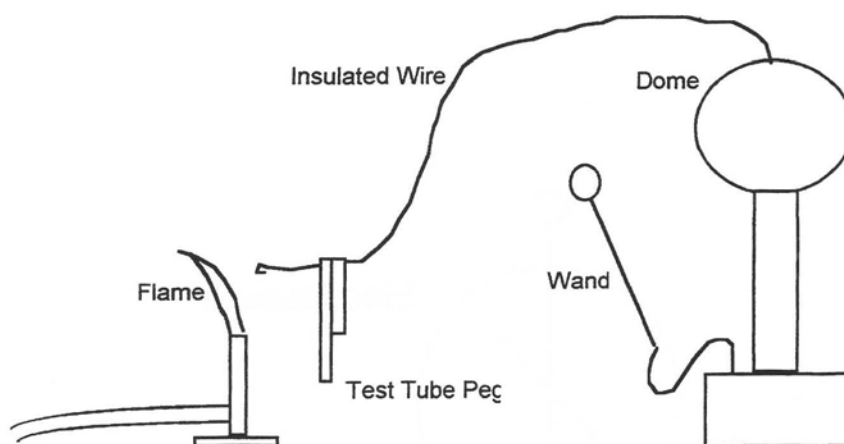
Connecting Wire

Bunsen

Wooden test tube peg

Procedure

- Place a connecting wire in the dome plug of the Van De Graaf generator.
- Start the generator
- Light a Bunsen within reach of the connecting wire and adjust to a yellow flame.
- Use the test tube peg to grasp the free end of the connecting wire.
- Bring the connecting wire near the side of the Bunsen flame.



Result: The Bunsen flame is repelled by the connecting wire as if blown by a wind.

Conclusion: The coloured part of a flame contains positive ions.

Risk Level: Moderate Hazard:

STUDENT: _____

248

Electron Attractions

Aim: To observe some reactions due to electrostatic force.

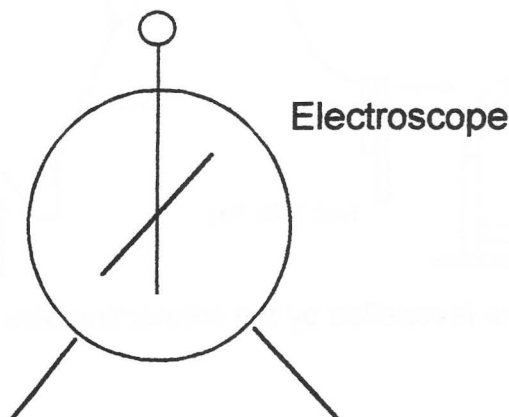
Equipment

Perspex Rod
Ebonite Rod
Dry cloths, two
Balloons, two
String, 50cm, two
Electroscope

Note: When a perspex rod is rubbed with a cloth, negatively charged electrons are lost to the cloth leaving the rod with a net positive charge.

Procedure

- 1/ Inflate two balloons and attach a peice of string to each
 - 2/ Rub the balloons vigourously with a cloth (charging).
 - 3/ Suspend the balloons by their strings and bring one balloon close to the other. Record anything unusual.
 - 4/ Bring a charged perspex rod close to the side of the balloons. Record the result.
 - 5/ Bring a charged ebonite rod (use a separate cloth) close to the balloons and record any result.
 - 6/ Charge a perspex rod and briefly bring it close the terminal of an electroscope.
 - 7/ Touch the electroscope terminal with a charged perspex rod then withdraw the rod. Record any difference from step 7.
 - 8/ Touch the charged electroscope with your finger.
- What is the charge of the balloons?
 - What is the charge of an ebonite rod?
 - What can you conclude about the forces between charges?



Results: _____

Conclusion: _____

Electron Attractions

Topics: Electrostatics

Aim: To observe some reactions due to electrostatic force.

Equipment

Perspex Rod
Ebonite Rod
Dry cloths, two
Balloons, two
String, 50cm, two
Electroscope

Note: When a perspex rod is rubbed with a cloth, negatively charged electrons are lost to the cloth leaving the rod with a net positive charge.

Electrostatic experiments require low humidity.

Procedure

- 1/ Inflate two balloons and attach a piece of string to each
 - 2/ Rub the balloons vigorously with a cloth (charging).
 - 3/ Suspend the balloons by their strings and bring one balloon close to the other. Record anything unusual.
 - 4/ Bring a charged perspex rod close to the side of the balloons. Record the result.
 - 5/ Bring a charged ebonite rod (use a silk cloth) close to the balloons and record any result.
 - 6/ Charge a perspex rod and briefly bring it close the terminal of an electroscope.
 - 7/ Touch the electroscope terminal with a charged perspex rod then withdraw the rod. Record any difference from step 7.
 - 8/ Touch the charged electroscope with your finger.
- What is the charge of the balloons?
 - What is the charge of an ebonite rod?
 - What can you conclude about the forces between charges?

Result: The balloons repel each other and are attracted to a perspex rod but repelled by the ebonite. The electroscope deflects when near a charge and stays deflected if the charge contacts. Finger contact neutralises the charge.

Conclusion: The balloons acquire a positive charge. The ebonite rod acquires a negative charge. Similar charges repel each other but opposite charges attract.

Risk Level: Low Hazard.

STUDENT: _____

249

Corrosion

Aim: To investigate the conditions affecting for the corrosion of iron.

Equipment

Test tube rack
test tubes, medium, six
Nails, 25mm, six
Steel wool
Cooking oil
Zinc Granules
Copper strip

Procedure

- 1/ Clean each nail with steel wool.
- 2/ Place the test tubes in the rack and a nail in each tube.
- 3/ Add 3cm of oil to the second tube.
- 4/ Add 3cm of water to the third tube.
- 5/ Add 3cm of water and then 1cm of oil to the fourth tube.
- 6/ Add a peice of zinc and 3cm of water to the fifth tube.
- 7/ Add a peice of copper and 3cm of water to the last tube.
- 8/ Leave the test tubes for one week
- 9/ Compare the rust formation in the tubes.

- What is the purpose of the conditions in each test tube?

| Tube | Purpose | Corrosion |
|------|---------|-----------|
| 1 | | |
| 2 | | |
| 3 | | |
| 4 | | |
| 5 | | |
| 6 | | |

Results: _____

Conclusion: _____

Corrosion

Topics: Elements Chemical Runs

Aim: To investigate the conditions affecting for the corrosion of iron.

Equipment

Test tube rack
test tubes, medium, six
Nails, 25mm, six
Steel wool
Cooking oil
Zinc Granules
Copper strip

Procedure

- 1/ Clean each nail with steel wool.
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- 5/ Add 3cm of water and then 1cm of oil to the fourth tube.
- 6/ Add a piece of zinc and 3cm of water to the fifth tube.
- 7/ Add a piece of copper and 3cm of water to the last tube.
- 8/ Leave the test tubes for one week
- 9/ Compare the rust formation in the tubes.

- What is the purpose of the conditions in each test tube?

| Tube | Purpose | Corrosion |
|------|--|-----------|
| 1 | Does rusting require air but not water? | |
| 2 | Does rusting occur without air or water? | |
| 3 | Does rusting require air and water? | |
| 4 | Does rusting require water but not air? | |
| 5 | Is rusting accelerated by a more active metal? | |
| 6 | Is rusting accelerated by a less active metal? | |

Result: No corrosion was apparent in tubes 1, 2, and 5. Only limited corrosion occurred in tube 4, tube three exhibited corrosion and tube 6 was strongly corroded.

Conclusion: Corrosion of iron (rust) is oxidation of iron to iron oxide. Oxygen is required as a reactant and water is needed as a medium for the ions to form and react. Zinc is more reactive and undergoes sacrificial corrosion, protecting the iron. Copper is less reactive than iron and stimulates sacrificial corrosion of the iron.

Risk Level: Low Hazard.

STUDENT: _____

250

Elastic Energy

Aim: To observe energy changes in an elastic band.

Equipment

Rubber band (No 11)

Procedure

There are three types of energy which can be interchanged in a rubber band: Elastic Potencial Energy, Kinetic Energy and Heat Energy

1/ Hold the rubber band between the the thumb and forefinger of both hands.

2/ Place the rubber band against your lower lip.

3/ Stretch the rubber band and while taught, place it against your lower lip.

4/ Relax the tension in the band while holding it against your lip.

- Describe the energy transformation which would take place if the rubber band were stretched over the tip of a ruler and then released.

- Describe an energy transformation taking place as the band is stretched.

- Will a band fly further if stretched and immediately released than if stretched, held a while and then released. Why?

Results: _____

Conclusion: _____

Elastic Energy

Topics: Energy

Aim: To observe energy changes in an elastic band.

Equipment

Rubber band (No 11)

Procedure

There are three types of energy which can be interchanged in a rubber band: Elastic Potential Energy, Kinetic Energy and Heat Energy

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3/ Stretch the rubber band and while taught, place it against your lower lip.

4/ Relax the tension in the band while holding it against your lip.

- Describe the energy transformation which would take place if the rubber band were stretched over the tip of a ruler and then released.

- Describe an energy transformation taking place as the band is stretched.

- Will a band fly further if stretched and immediately released than if stretched, held a while and then released. Why?

Result: The rubber band is warmer when stretched and cooler when relaxed.

Conclusion: As a rubber band is stretched it stores elastic potential energy and if the band is suddenly released the energy is converted into kinetic energy as the band flies off. However when the band is stretched, some of the elastic potential energy is converted to heat energy which can be transmitted away, so pausing before releasing the band ensures it will not fly as far.

Risk Level: Low Hazard: There is sure to be some launching of rubber bands, the trick is not to let it get out of control.

Aim: To measure the energy output of a simple circuit and so derive its power in Watts.

Equipment

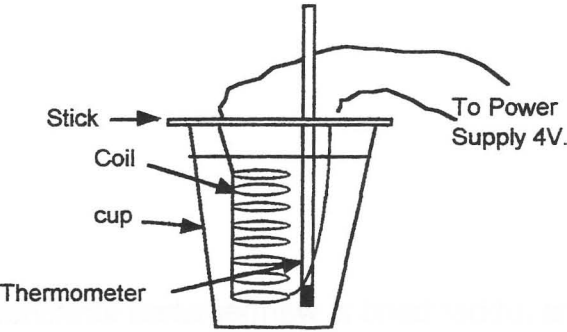
- Power Supply, DC 0 - 12V
- Connecting Leads, two
- Alligator clips, two
- Nichrome wire, 0.5m
- Foam Cup
- Thermometer, 0-100
- Paddle Pop stick
- Measuring Cylinder, 100ml

Optional: Multimeter

Procedure

- 1/ Wind the nichrome wire in even coils around a finger.
- 2/ Use an alligator clip on a to clamp one end of the coil to the paddle pop stick.
- 3/ Unwind 3 loops at the opposite end of the coil, straighten the wire and clip it back to the paddle pop stick with an alligator clip.
- 4/ Lower the nichrome coil into the foam cup.
- 5/ Measure water into the foam cup until the coil is covered.
- 6/ Record the temperature of the water.
- 7/ Use the connecting leads to join the alligator clips to the DC terminals of the power supply.
- 8/ Set the voltage to 4V and turn on the power for 300 seconds.
- 9/ Record the new temperature of the water.
- 10/ Calculate the energy released by the circuit from:
$$\text{Heat Energy (Joules)} = \text{Temperature Change} \times \text{Water Volume (ml)} \times 4.180 \text{ (Heat Capacity)}$$
- 11/ Calculate the power output from:
$$\text{Power (Watts)} = \text{Energy} / \text{Time (seconds)}$$

| Record | Figure |
|-------------|--------|
| Water Vol. | |
| First temp. | |
| Final Temp. | |
| Energy | |
| Power | |



Results: _____

Conclusion: _____

Power

Topics: Energy Electricity

Aim: To measure the energy output of a simple circuit and so derive its power in Watts.

Equipment

Power Supply, DC 0 - 12V
Connecting Leads, two
Alligator clips, two
Nichrome wire, 0.5m
Foam Cup
Thermometer, 0-100
Paddle Pop stick
Measuring Cylinder, 100ml

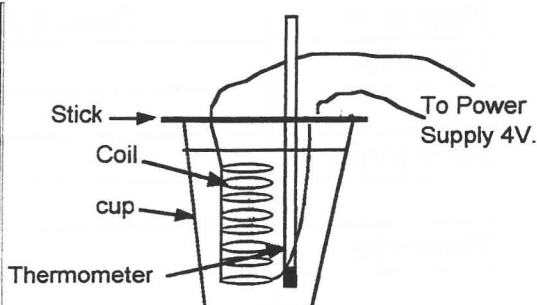
Optional: Multimeter

The Power Output can be confirmed by measuring the voltage and current of the circuit and using $P=VI$.

Procedure

- 1/ Wind the nichrome wire in even coils around a finger.
- 2/ Use an alligator clip to clamp one end of the coil to the paddle pop stick.
- 3/ Unwind 3 loops at the opposite end of the coil, straighten the wire and clip it back to the paddle pop stick with an alligator clip.
- 4/ Lower the nichrome coil into the foam cup.
- 5/ Measure water into the foam cup until the coil is covered.
- 6/ Record the temperature of the water.
- 7/ Use the connecting leads to join the alligator clips to the DC terminals of the power supply.
- 8/ Set the voltage to 4V and turn on the power for 300 seconds.
- 9/ Record the new temperature of the water.
- 10/ Calculate the energy released by the circuit from:
Heat Energy = Temperature X Water X 4.180
(Joules) Change Volume (ml) (Heat Cap.)
- 11/ Calculate the power output from:
Power (Watts) = Energy / Time (seconds)

| Record | Figure |
|-------------|--------|
| Water Vol. | |
| First temp. | |
| Final Temp. | |
| Energy | |
| Power | |



Result: The water temperature rose by about 10 degrees.

Conclusion: The Power Output of the circuit is about 15Watts, the electrical energy being converted into heat energy.

Risk Level: Low Hazard: Beware of students using high voltage settings since this will cause the circuit breakers in the power supplies to trip. Caution students not to turn on the power unless the wire is immersed otherwise burnt fingers and marked benches will result.

STUDENT: _____

252

Competing Equilibria

Aim: To demonstrate that a series of reactions can take place involving the same anion.

Equipment

Test Tube, large
Silver nitrate, 0.1M, (1.7%)
Dropper Bottles:
Ammonia, 0.1M, (0.7%)
Sodium chloride, 0.1M
(0.6%)
Sodium Bromide, 0.1M, 1%
Sodium Thiosulfate, 0.1M,
(1.6%)
Sodium Iodide, 1M, (1.5%)

Procedure

- 1/ Add 1cm of silver nitrate solution to the test tube.
- 2/ Mix sodium chloride dropwise until a change occurs.
- 3/ Mix ammonia solution dropwise until a change occurs.
- 4/ Mix sodium bromide dropwise until a change occurs.
- 5/ Mix sodium thiosulfate dropwise until a change occurs.
- 6/ Mix sodium iodide dropwise until a change occurs.

- How can a precipitate be forced to dissolve? (Think in terms of reversible reactions)
- Write an equation for each reaction (Thiosulfate = $\text{S}_2\text{O}_3^{2-}$)
- What can you conclude about the equilibrium constant of the reactions?

| Reagent | Result / Equation |
|-------------|-------------------|
| Chloride | |
| | |
| Ammonia | |
| | |
| Bromide | |
| | |
| Thiosulfate | |
| | |
| Iodide | |
| | |

Results: _____

Conclusion: _____

Competing Equilibria

Topics: Equilibrium Chem Reactions

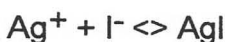
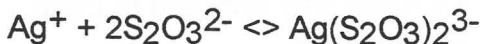
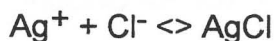
Aim: To demonstrate that a series of reactions can take place involving the same anion.

Equipment

Test Tube, large
Silver nitrate, 0.1M, (1.7%)
Dropper Bottles:
Ammonia, 0.1M, (0.7%)
Sodium chloride, 0.1M
(0.6%)
Sodium Bromide, 0.1M, 1%
Sodium Thiosulfate, 0.1M,
(1.6%)
Sodium Iodide, 1M, (1.5%)

Procedure

- 1/ Add 1cm of silver nitrate solution to the test tube.
 - 2/ Mix sodium chloride drop wise until a change occurs.
 - 3/ Mix ammonia solution drop wise until a change occurs.
 - 4/ Mix sodium bromide drop wise until a change occurs.
 - 5/ Mix sodium thiosulfate drop wise until a change occurs.
 - 6/ Mix sodium iodide drop wise until a change occurs.
- Write an equation for each reaction (Thiosulfate = $\text{S}_2\text{O}_3^{2-}$)
 - What can you conclude about the equilibrium constant of the reactions?



Result: A white precipitate forms with the chloride but dissolves with ammonia. A yellow precipitate forms with the bromide but dissolves with the thiosulfate. A bright yellow precipitate forms with the iodide.

Conclusion: In case a reaction occurs with a higher numerical equilibrium constant. The precipitates dissociate because their formation reaction is forced to reverse as the silver ions are consumed in a new reaction.

Risk Level: Moderate Hazard: Concentrated or solid silver nitrate is corrosive to skin and eyes, harmful by ingestion and stains skin. Contact is to be avoided with the other silver products in this experiment. The sodium reagents in the experiment pose minimal hazard in these concentrations.

STUDENT: _____

253

Equilibrium & Heat

Aim: To confirm that temperature effects the direction of a reaction.

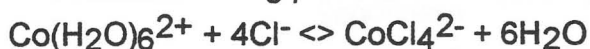
Equipment

Cobalt Chloride
Hydrochloric Acid 1M(10%)
Test Tube (20ml)
Beakers, 250ml, two
Test tube peg
Ice
Hot water
Measuring Cylinder, 10ml
Filter paper

Procedure

- 1/ Weigh 0.1g cobalt chloride on a filter paper.
- 2/ Add the cobalt chloride to a test tube and dissolve with 5mls of water. Note the colour of the solution.
- 3/ Add 5ml of 1M hydrochloric acid.
- 4/ Note the colour of the solution.
- 5/ Add Ice water to a beaker. Place the test tube in the ice water and note any colour change after five minutes.
- 6/ Place hot water in a second beaker. Place the test tube in this beaker and note any colour change after 5 minutes.

The reaction taking place is as follows:



- Which cobalt ion is responsible for red colour?
- Which cobalt ion is responsible for blue colour?
- Explain why only one of these ions predominates at a certain temperature.
- Explain any other colours observed.

Results: _____

Conclusion: _____

Equilibrium & Heat

Topics: Equilibrium Chem Reactions

Aim: To confirm that temperature effects the direction of a reaction.

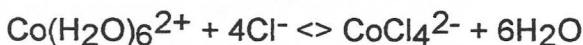
Equipment

Cobalt Chloride
Hydrochloric Acid 1M(10%)
Test Tube (20ml)
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Hot water
Measuring Cylinder, 10ml
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The reaction taking place is as follows:



- Which cobalt ion is responsible for red colour?
- Which cobalt ion is responsible for blue colour?
- Explain why only one of these ions predominates at a certain temperature.
- Explain any other colours observed.

Result: The original solution is pale red. Adding acid produces a violet colour. Cooling the reaction produces a red colour while heating results in a blue colour.

Conclusion: $\text{Co}(\text{H}_2\text{O})_6^{2+}$ is a red ion produced by dissolving the cobalt chloride.

CoCl_4^{2-} is a blue ion which predominates when the reaction with hydrochloric acid is pushed to the right by heating. Cooling forces the equilibrium back to the left. Violet is a mixed equilibrium at room temperature.

Risk Level: Moderate Hazard: Cobalt chloride is toxic if ingested and may be carcinogenic. Hydrochloric acid is corrosive and contact with the skin should be treated with prolonged washing

Paper Planes

Aim: To design and build a paper plane with the longest straight flight.

Equipment

Paper (used)
Designs (eg. Best Ever
Paper Planes by Paul
Jackson, ISBN 1-85479-
334-9)

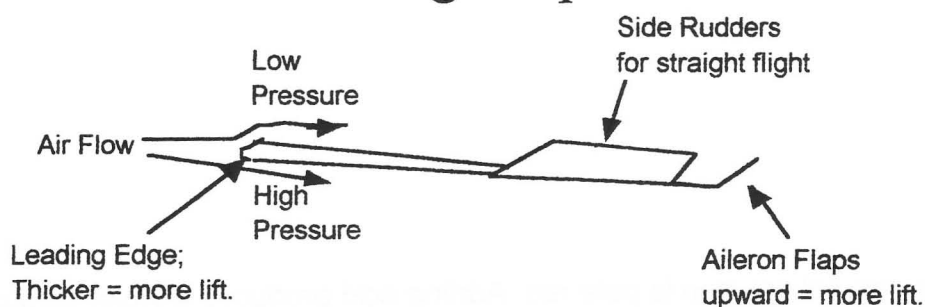
Procedure

- 1/ Fold three paper planes of different design.
- 2/ Test your planes in the school hall.
- 3/ Using the hints below modify the design that flew best.
- 4/ Test fly your design and make adjustments again.

The process above is a design procedure that mimics evolution. Firstly a random set of designs a chosen. A test is performed to select the best design. The test is similar to selection pressure in the wild ensuring only the fittest survive. Next changes are made to the selected design. These changes are similar to mutation and cross breeding in nature. The modified designs are tested and again the best are selected for further changes. In nature, success is passing your genes to your offspring.

Record the features of the best design.

Wing Shape



Results: _____

Conclusion: _____

Paper Planes

Topics: Flight

Evolution

Aim: To design and build a paper plane with the longest straight flight.

Equipment

Paper (used)

Designs (eg. Best Ever

Paper Planes by Paul

Jackson, ISBN 1-85479-334-9)

Procedure

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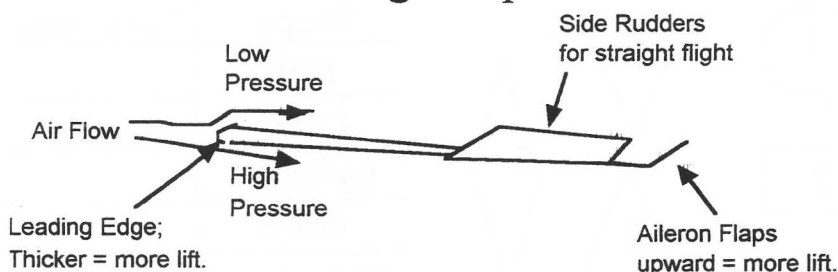
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A sweet reward will add incentive and help self regulate the competition.

Record the features of the best design.

Wing Shape



Result:

Conclusion: Rudders and balanced lift (nose to tail and across both wings) will be evident in the best design. Darts fall too fast to achieve great distance.

Risk Level: Low Hazard: Great Fun

STUDENT: _____

Streamlining

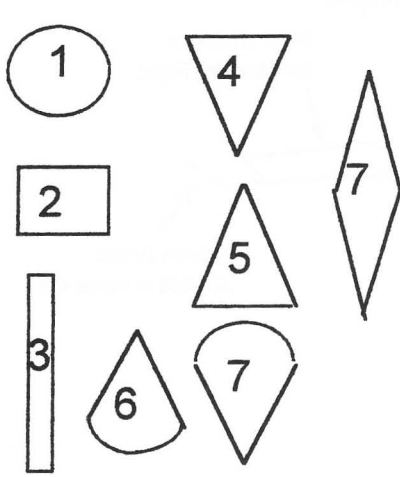
Aim: To investigate the effect of streamlining the shape of an object moving through a liquid or gas.

Equipment

- Plasticine
- Screw, 20mmX 8G
- Measuring cylinder, 100ml
- Stop watch
- Chalk

Procedure

- 1/ Role 3cm of plasticine into a ball.
- 2/ Insert the screw to give one end added weight.
- 3/ Scrape a little chalk dust into the measuring cylinder.
- 4/ Fill the measuring cylinder with water.
- 5/ Drop the ball into the water (weighted end down) and use the stop watch to measure the time for the ball to reach the bottom of the cylinder.
- 6/ Repeat the test observing the water flow around the object by watching the behaviour of the suspended chalk dust.
- 7/ Change the shape of the plasticine and repeat the test.
- 8/ Repeat until all the shapes below have been tested.



| Shape | Time (secs) |
|------------|-------------|
| ball | |
| cube | |
| cylinder | |
| cone | |
| cone (inv) | |
| Tear | |
| Tear (inv) | |
| 2X Cone | |

Results: _____

Conclusion: _____

Streamlining

Topics: Flight Density & Pressure

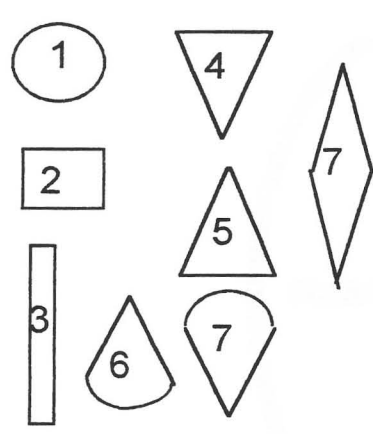
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| Shape | Time (secs) |
|------------|-------------|
| ball | |
| cube | |
| cylinder | |
| cone | |
| cone (inv) | |
| Tear | |
| Tear (inv) | |
| 2X Cone | |

Result: The double cone moved through the water fastest, the cube moving slowest. The water was observed to move rapidly around broad surfaces and to swirl behind objects without a pointed trailing end.

Conclusion: Objects pointed at the leading edge part the liquid easily and allow a smooth flow over the surface. Objects pointed at the trailing end minimise swirling turbulence. Objects pointed at both ends are streamlined, that is, shaped to allow smooth flow of the liquid or gas over the surfaces and minimising turbulence.

Risk Level: Low hazard: Surprisingly popular, perhaps because it provides an excuse for older students to play with plasticine.

STUDENT: _____

256

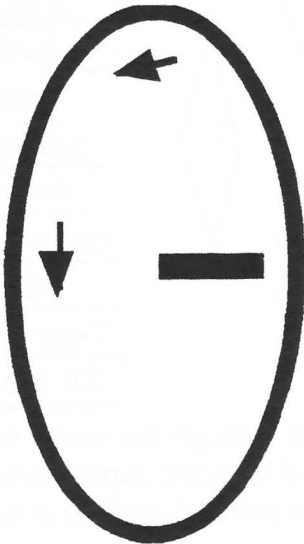
Angular Momentum

Aim: To observe angular momentum in action and its implications for the Earth spinning on its axis.

Equipment
Bicycle wheel
stunt pegs, two
Gyroscope

Procedure
A spinning object has a gyroscopic property in that it resists any torque applied to its axis, producing a resultant force which is at 90 degrees to the axis and the applied torque.
1/ Hold the bicycle wheel in front of your body by gripping the stunt pegs on either side.
2/ An assistant spins the wheel rapidly.
3/ Try turning your body to to the left or right.
4/ Try tilting the wheel.
5/ Start a gyroscope, virtically on the table and observe its behaviour.

- If the gyroscope were the Earth, what would be happening to the tilt of its axis?
- Explain the movement of the gyroscope in terms of your observation with the bicycle wheel.
- What happens to the lean of a bike if you turn to the left?
- Could a bike be steered simply by tilting the frame?



Results: _____

Conclusion: _____

Angular Momentum

Topics: Forces Earth & Space

Aim: To observe angular momentum in action and its implications for the Earth spinning on its axis.

Equipment

Bicycle wheel
stunt pegs, two
Gyroscope

Procedure

A spinning object has a gyroscopic property in that it resists any torque applied to its axis, producing a resultant force which is at 90 degrees to the axis and the applied torque.

1/ Hold the bicycle wheel in front of your body by gripping the stunt pegs on either side.

2/ An assistant spins the wheel rapidly.

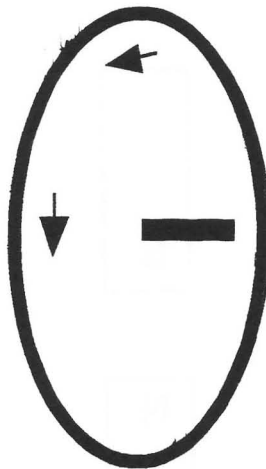
3/ Try turning your body to the left or right.

4/ Try tilting the wheel.

5/ Start a gyroscope, vertically on the table and observe its behaviour.

- If the gyroscope were the Earth, what would be happening to the tilt of its axis?
- Explain the movement of the gyroscope in terms of your observation with the bicycle wheel.
- What happens to the lean of a bike if you turn to the left?
- Could a bike be steered simply by tilting the frame?

The axis of the Earth precesses once every 10,000 years and this is thought to cause cyclic ice ages.



Result: Turning to the left results in the wheel tilting to the left. Tilting the wheel to the left results in the axis of the wheel turning to the left.

Conclusion: Since the Earth is much like a gyroscope and undergoes a torque from the gravitational pull of the sun, the axis of the Earth wobbles (precesses). The gyroscope wobbles in a circle because of the gravitational pull of the Earth. A bike turning left, tilts to the left and can actually be steered by just tilting the frame.

Risk Level: Low Hazard: Tremendous fun. The students really are intrigued. These forces are not actually noticed when the bicycle wheel is fixed in a frame. Have you noticed that car races are always in a clockwise circuit? This is because the torque on the rotating crankshaft will force the front wheels onto the ground. A clockwise circuit would reduce grip in the steering unless the engine rotation were changed. East/west engines suffer a flip torque.

STUDENT: _____

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Force Fields

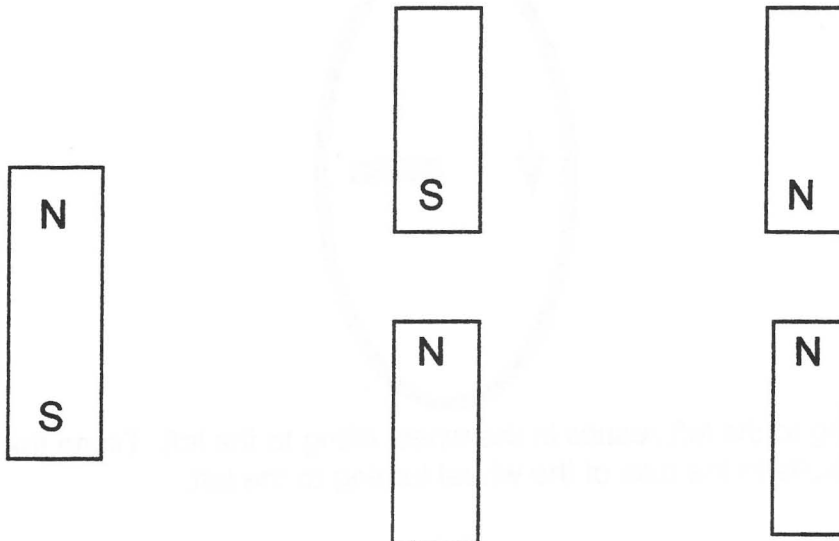
Aim: To reveal invisible lines of magnetic force.

Equipment

Bar Magnets, two
Iron filings
Paper, A4 sheet
Specimen tube, small

Procedure

- 1/ Place two spatulas of iron filings in the specimen tube.
- 2/ Place the sheet of paper over a bar magnet on the bench.
- 3/ Using the specimen tube like a salt shaker, sprinkle iron filings on the paper until a pattern forms.
- 4/ Draw the pattern in the space below.
- 5/ Lift the paper to form a U-shaped trough and pour the filings back into the specimen tube.
- 6/ Place two magnets on the bench so that the South pole of one is about 2cm from the North pole of the second magnet.
- 7/ Lay the sheet of paper over the magnets and sprinkle filings until a pattern forms.
- 8/ Draw the pattern in the space below.
- 9/ Lift the paper and return the filings to the specimen tube.
- 10/ Place two magnets on the bench so that the North pole of one is about 2cm from the North pole of the other magnet.
- 11/ Lay paper over the magnets and sprinkle iron filings until a pattern forms.
- 12/ Draw the pattern in space below.



Results: _____

Conclusion: _____

Force Fields

Topics: Forces

Aim: To reveal invisible lines of magnetic force.

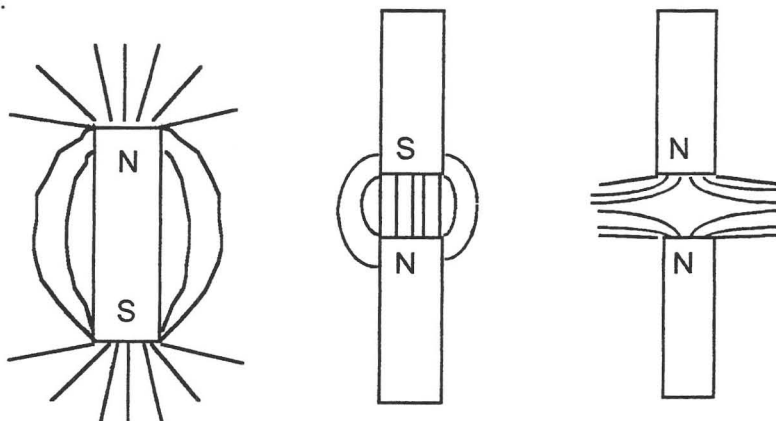
Equipment

Bar Magnets, two
Iron filings
Paper, A4 sheet
Specimen tube, small

Procedure

- 1/ Place two spatulas of iron filings in the specimen tube.
- 2/ Place the sheet of paper over a bar magnet on the bench.
- 3/ Using the specimen tube like a salt shaker, sprinkle iron filings on the paper until a pattern forms.
- 4/ Draw the pattern in the space below.
- 5/ Lift the paper to form a U-shaped trough and pour the filings back into the specimen tube.
- 6/ Place two magnets on the bench so that the South pole of one is about 2cm from the North pole of the second magnet.
- 7/ Lay the sheet of paper over the magnets and sprinkle filings until a pattern forms.
- 8/ Draw the pattern in the space below.
- 9/ Lift the paper and return the filings to the specimen tube.
- 10/ Place two magnets on the bench so that the North pole of one is about 2cm from the North pole of the other magnet.
- 11/ Lay paper over the magnets and sprinkle iron filings until a pattern forms.
- 12/ Draw the pattern in space below.

Prepare the specimen tube by making a central hole through the inner surface of the plastic cap with a hot dissecting needle.



Result: Lines of force radiate from each pole of a magnet. Lines of force link and arch around the space between a North and South pole. Lines of force clash and spread perpendicularly between two North magnetic poles.

Conclusion: Lines of force from magnetic poles have a direction. Between opposite poles the lines run into each other. Between similar poles the lines repel each other and are forced to the sides.

Risk Level: Low Hazard.

Magnetism

Aim: To investigate magnetic force.

Equipment

Bar Magnets, two
Magnetic compass
Iron Nails, 20mm, 3
Alfoil, 5cmX5cm
Copper strip

Procedure

- 1/ Investigate what materials are attracted to a magnet. Try wood, plastic, iron, copper, aluminium etc.
- Does it matter which end of the magnet you use?
- 2/ Investigate how a compass needle reacts to the North pole of a magnet.
- 3/ Investigate how a compass needle reacts to the South pole of a magnet.
- 4/ Investigate how the South pole of one magnet reacts when the South pole of another magnet is brought near.
- 5/ Investigate how the South pole of one magnet reacts when the North pole of another magnet is brought near.
- 6/ Investigate how the North pole of one magnet reacts when the North pole of another magnet is brought near.
- Write two simple rules which describe how magnetic poles react when they approach each other.
- Is a compass simply a piece of iron?
- Why do you think the poles of a magnet are named North and South?
- What is at the North Pole of the Earth to attract a compass?

Results: _____

Conclusion: _____

Magnetism

Topics: Forces

Aim: To investigate magnetic force.

Equipment

Bar Magnets, two
Magnetic compass
Iron Nails, 20mm, 3
Alfoil, 5cmX5cm
Copper strip

Procedure

- 1/ Investigate what materials are attracted to a magnet. Try wood, plastic, iron, copper, aluminium etc.
 - Does it matter which end of the magnet you use?
- 2/ Investigate how a compass needle reacts to the North pole of a magnet.
- 3/ Investigate how a compass needle reacts to the South pole of a magnet.
- 4/ Investigate how the South pole of one magnet reacts when the South pole of another magnet is brought near.
- 5/ Investigate how the South pole of one magnet reacts when the North pole of another magnet is brought near.
- 6/ Investigate how the North pole of one magnet reacts when the North pole of another magnet is brought near.
 - Write two simple rules which describe how magnetic poles react when they approach each other.
 - Is a compass simply a piece of iron?
 - Why do you think the poles of a magnet are named North and South?
 - What is at the North Pole of the Earth to attract a compass?

Result: Iron objects are attracted both ends of a magnet. Compass needles are repelled by North poles but attracted by South poles. South poles repel each other but attract North poles. North poles repel each other.

Conclusion: LIKE POLES REPEL. OPPOSITE POLES ATTRACT. A compass is a small magnet. The North pole of a compass or magnet will point to the North Pole of the Earth. The North Pole of the Earth must have a South magnetic pole.

Risk Level: Low Hazard

STUDENT: _____

259

Tension & Acceleration

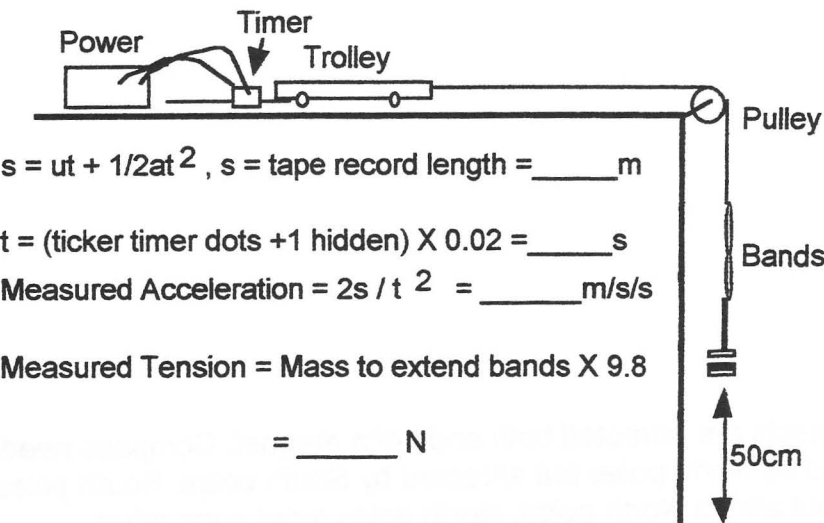
Aim: To experimentally determine the acceleration and tension in a system in motion.

Equipment

- Collision trolley
- Bench pulley
- String, 80cm
- Mass Carrier
- Masses, 50g(4),25g,5g (4)
- Ticker Timer
- Paper Tape
- Power supply, 0-12V, AC
- Connecting wires, two
- Rubber bands, two
- Adhesive tape
- Metre rule.

Procedure

- 1/ Adjust the trolley mass to 600g by adding masses.
- 2/ Tie one end of the string to the trolley. Tie the rubber bands together and tie to the other end of the string.
- 3/ Attach 60cm of paper tape to the rear of the trolley with adhesive tape. Pass the paper tape through the ticker timer.
- 4/ Set the power supply to 6V and connect the ticker timer to the AC terminals.
- 5/ Place two 50g masses on the mass carrier and hook the carrier onto the end rubber band.
- 6/ Move the trolley and ticker timer until the mass carrier is 50cm above the floor.
- 7/ Simultaneously release the trolley and turn on the power, switching the power off as the mass carrier strikes the floor.
- 8/ Measure the length of the ticker timer record. _____
- 9/ Deducting the fall height will give the change in extension of the rubber bands.
- 10/ Suspend the mass carrier and replace weights until the change in extension is duplicated. Mass = _____
- 11/ Count the dots on the ticker timer record. _____



Results: _____

Conclusion: _____

Tension & Acceleration

Topics: Forces

Aim: To experimentally determine the acceleration and tension in a system in motion.

Equipment

Collision trolley
Bench pulley
String, 80cm
Mass Carrier
Masses, 50g(4), 25g, 5g (4)
Ticker Timer
Paper Tape
Power supply, 0-12V, AC
Connecting wires, two
Rubber bands, two
Adhesive tape
Metre rule.

Procedure

- 1/ Adjust the trolley mass to 600g by adding masses.
- 2/ Tie one end of the string to the trolley. Tie the rubber bands together and tie to the other end of the string.
- 3/ Attach 60cm of paper tape to the rear of the trolley with adhesive tape. Pass the paper tape through the ticker timer.
- 4/ Set the power supply to 6V and connect the ticker timer to the AC terminals.
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- 6/ Move the trolley and ticker timer until the mass carrier is 50cm above the floor.
- 7/ Simultaneously release the trolley and turn on the power, switching the power off as the mass carrier strikes the floor.
- 8/ Measure the length of the ticker timer record. _____
- 9/ Deducting the fall height will give the change in extension of the rubber bands.
- 10/ Suspend the mass carrier and replace weights until the change in extension is duplicated. Mass = _____
- 11/ Count the dots on the timer timer record. _____

Theoretical

$$\begin{aligned}\text{Net accelerating force} &= \text{mass} \times \text{gravity} \\ &= 0.15\text{kg} \times 9.8 = 1.03\text{N}\end{aligned}$$

$$\begin{aligned}\text{Acceleration} &= \text{Force} / \text{total mass} \\ &= 1.013 / (0.6 + 0.15) = 1.35\text{m/s/s}\end{aligned}$$

$$\begin{aligned}\text{Tension} &= \text{Trolley mass} \times \text{acceleration} \\ &= 0.6 \times 1.35 = 0.81\text{N}\end{aligned}$$

Result: The ticker timer recorded the acceleration of the trolley and the change in extension of the rubber bands. The extension of the rubber bands is a measure of tension in the string.

Conclusion: Friction will ensure that the measured acceleration will be less than the theoretical value. While this would imply a reduced tension value from the theory calculation, friction actually increases the tension required to achieve the reduced acceleration. As a result the measured tension will be close to the theoretical value.

Risk Level: Low Hazard:

STUDENT: _____

260

Pressure Ignition

Aim: To demonstrate that if the pressure applied to a given volume of gas is sharply increased, the temperature will rise even to ignition point.

Equipment

"Pressure Ignition"
apparatus.

Cotton wool

Alcohol

Vacuum grease

Procedure

1/ Insure the plunger is greased and pliable.

2/ Moisten a small wad of cotton wool with alcohol.

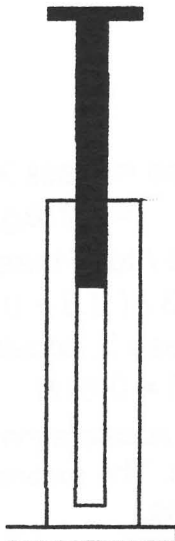
3/ Squeeze out the excess alcohol.

4/ Insert the cotton wool into the base of the cylinder.

5/ Position the plunger.

6/ Ram the plunger toward the base.

This principle is used in Diesel Engines where the compression by the piston provides ignition instead of a spark plug.



Results: _____

Conclusion: _____

Pressure Ignition

Topics: Gas Laws Density/Pressure

Aim: To demonstrate that if the pressure applied to a given volume of gas is sharply increased, the temperature will rise even to ignition point.

Equipment

"Pressure Ignition"
apparatus.

Cotton wool

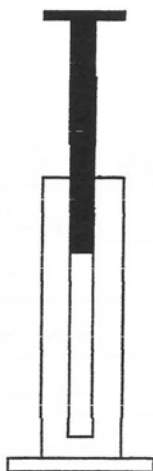
Alcohol

Vacuum grease

Procedure

- 1/ Insure the plunger is greased and pliable.
- 2/ Moisten a small wad of cotton wool with alcohol.
- 3/ Squeeze out the excess alcohol.
- 4/ Insert the cotton wool into the base of the cylinder.
- 5/ Position the plunger.
- 6/ Ram the plunger toward the base.

This principle is used in Diesel Engines where the compression by the piston provides ignition instead of a spark plug.



Result: The cotton wool ignites

Conclusion: As pressure increases sharply for a given volume, temperature increases.

$$PV = kT$$

Risk Level: Low Hazard: However the apparatus will not last long in the hands of students.
Teacher demonstration only.

STUDENT: _____

Genetic Differences

Aim: To investigate dominant and recessive genes in humans.

Equipment

Procedure

Dominant genes are expressed even if only one chromosome in a pair carries the gene. Recessive genes are only expressed if both chromosomes in a pair to carry the gene. In a multicultural society such as Australia, dominant genes will be much more common than recessive genes.

Survey the class for the characteristics listed in the table below, recording how many students have those characteristics.

From the results try to conclude which genes are likely to be dominant and which are likely to be recessive.

| Trait | Number | Trait | Number |
|-------------------|--------|-------------------|--------|
| Brown Eyes | | Black/Brown hair | |
| Blue eyes | | Fair hair | |
| Hasel eyes | | Red Hair | |
| Darwins Ear Point | | Peaked Hair line | |
| No Ear Point | | Smooth Hair Line | |
| Tongue Rolling | | No Tongue Rolling | |

Results: _____

Conclusion: _____

Genetic Differences

Topics: Genetics

Aim: To investigate dominant and recessive genes in humans.

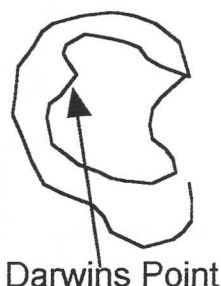
Equipment

Procedure

Dominant genes are expressed even if only one chromosome in a pair carries the gene. Recessive genes are only expressed if both chromosomes in a pair carry the gene. In a multicultural society such as Australia, dominant genes will be much more common than recessive genes.

Survey the class for the characteristics listed in the table below, recording how many students have those characteristics.

From the results try to conclude which genes are likely to be dominant and which are likely to be recessive.



Result: Black/ Brown hair, brown eyes, Darwins Point, tongue rolling and peaked hair line are all common traits. Red hair, blue eyes, no ear point, no tongue rolling and smooth hair line are much less common.

Conclusion: Black/ Brown hair, brown eyes, Darwins Point, tongue rolling and peaked hair line are all due to dominant genes. Red hair, blue eyes, no ear point, no tongue rolling and smooth hair line are due to recessive genes. Fair hair and Hazel eyes are dominant genes which share expression with other colour genes.

Risk Level: Low Hazard: Be careful about making statements like " parents with blue eyes will not have brown eyed offspring". Few human traits are governed by a single gene locus and so simple Mendelian rules do not apply. You do not want to have a student doubting their parents.

STUDENT: _____

262

Zeibops

Aim: To investigate genetic recombination through generations of the imaginary animals, Zeibops.

Equipment

Coloured Pop sticks

(chromosome pairs) :

Two, red, 10cm, "A" and "a"

Two, red, 8.5cm, "T" and "t"

Two, red, 7cm, "Q" and "q"

Two, red, 5.5cm, "D" and "d"

Two, red, 4cm, "E" and "e"

Two, red, 2.5cm, "G" and "g"

Repeat set in green

Marsh mellow, four

Match sticks, eight

Drawing pins, three

Procedure

1/ Take the red male chromosomes and randomly take one from each pair (meiosis producing sperm).

2/ Take the green female chromosomes and randomly take one from each pair (meiosis producing ova).

3/ Match the remaining red chromosomes with the remaining green chromosomes (fertilisation).

4/ Construct the progeny according to:

AA or Aa - three body segments - marsh mellow

aa - two body segments

TT or Tt - six legs (match sticks), tt - 4 legs

QQ - long neck (match sticks), Qq short neck, qq - no neck

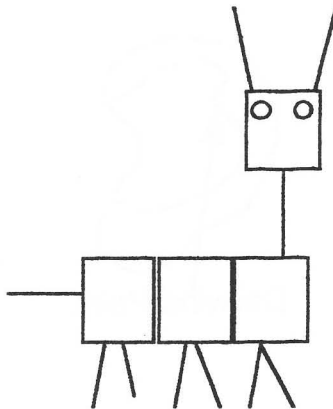
Head is a marsh mellow.

DD - 2 antennae (match sticks), Dd 1 antennae, dd - none

EE or Ee - tail (match stick), ee - no tail

GG or Gg - two eyes (drawing pins), gg - three eyes

5/ Copy the table of class results drawn up by the teacher.



A parent Zeibop

Results: _____

Conclusion: _____

Topics: Genetics

Aim: To investigate genetic recombination through generations of the imaginary animals, Zeibops.

Equipment

Coloured Pop sticks
(chromosome pairs) :
Two,red,10cm,“A” and “a”
Two,red, 8.5cm, “T” and “t”
Two,red, 7cm, “Q” and “q”
Two,red,5.5cm, “D” and “d”
Two,red,4cm, “E” and “e”
Two,red,2.5cm, “G” and “g”
Repeat set in green
Marsh mellows, four
Match sticks, eight
Drawing pins, three

Procedure

- 1/ Take the red male chromosomes and randomly take one from each pair (meiosis producing sperm).
- 2/ Take the green female chromosomes and randomly take one from each pair (meiosis producing ova).
- 3/ Match the remaining red chromosomes with the remaining green chromosomes (fertilisation).
- 4/ Construct the progeny according to:
 - AA or Aa - three body segments - marsh mellows
 - aa - two body segments
 - TT or Tt - six legs (match sticks), tt - 4 legs
 - QQ - long neck (match sticks), Qq short neck, qq - no neck
 - Head is a marsh mellow.
 - DD - 2 antennae (match sticks),Dd 1 antennae, dd - none
 - EE or Ee - tail (match stick), ee - no tail
 - GG or Gg - two eyes (drawing pins), gg - three eyes
- 5/ Copy the table of class results drawn up by the teacher.

| Phenotype | Number | % | Genotype | Number | % |
|------------|--------|---|----------|--------|---|
| 3 segments | | | AA | | |
| | | | Aa | | |
| 2 segments | | | aa | | |
| 6 legs | | | TT | | |
| | | | Tt | | |
| 4 legs | | | tt | | |
| Long Neck | | | QQ | | |
| Short Neck | | | Qq | | |
| No Neck | | | qq | | |

Result: Dominant phenotypes occur in 75% of offspring while the recessive phenotype occurs in 25% of offspring but the genotypes are 25% pure dominant, 50% hybrid dominant and 25% pure recessive.

Conclusion: Meiosis halves the chromosome pairs randomly producing a variety of characteristics in the offspring when the chromosome pairs are reformed in fertilisation.

Risk Level: Low Hazard: promise to give them clean marsh mellows to eat at the end of the practical.

STUDENT: _____

263

Antisepsis

Aim: To compare various substances designed to kill microbes.

Equipment

Petri Dish

Culture Solution : (1 beef stock cube, 10g Agar Agar per 500ml water, heated to boiling).

Cork borer

Marker pen

Paper tissues

Forceps

Filtered rise water from a garbage bin.

Dropper Bottles of:

Sodium Chloride 3%

Bleach (10% dilution)

Phenol Antiseptic(10% dil.)

Ethanol (10% dilution)

Procedure

1/ Pour hot culture solution into the petri dish and allow to gel (about ten minutes).

2/Rinse the surface of the culture gel with some of the garbage bin rinse.

3/ Use the cork borer tool to make four holes in the gel as shown below.

4/ Mark the outside edge of the petri dish adjacent to one hole, now hole number 1, progressing 2, 3, 4, clockwise.

5/ Fill hole 1 with sodium chloride solution.

6/ Fill hole number 2 with diluted bleach.

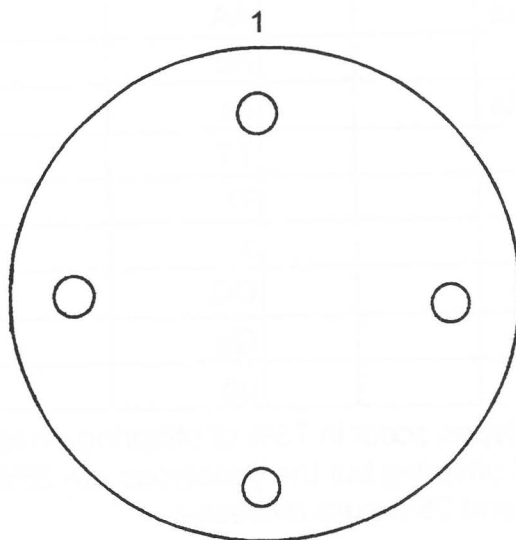
7/ Fill hole number 3 with diluted antiseptic.

8/ Fill hole number 4 with diluted ethanol.

9/ Place the cover on the petri dish and leave in a warm, dark place for 48 hours.

10/ IN THE FUME HOOD - The petri dish lid is probably opaque due to internal condensation. Remove the lid and mop up the condensation with a tissue held with forceps. Replace the lid and inspect the culture at your work bench.

11/ Draw the appearance of the culture plate on the diagram below.



Results: _____

Conclusion: _____

Antisepsis

Topics: Inside Life

Aim: To compare various substances designed to kill microbes.

Equipment

Petri Dish

Culture Solution : (1 beef stock cube, 10g Agar Agar per 500ml water, heated to boiling).

Cork borer

Marker pen

Paper tissues

Forceps

Filtered rise water from a garbage bin.

Dropper Bottles of:

Sodium Chloride 3%

Bleach (10% dilution)

Phenol Antiseptic(10% dil.)

Ethanol (10% dilution)

Procedure

1/ Pour hot culture solution into the petri dish and allow to gel (about ten minutes).

2/Rinse the surface of the culture gel with some of the garbage bin rinse.

3/ Use the cork borer tool to make four holes in the gel as shown below.

4/ Mark the outside edge of the petri dish adjacent to one hole, now hole number 1, progressing 2, 3, 4, clockwise.

5/ Fill hole 1 with sodium chloride solution.

6/ Fill hole number 2 with diluted bleach.

7/ Fill hole number 3 with diluted antiseptic.

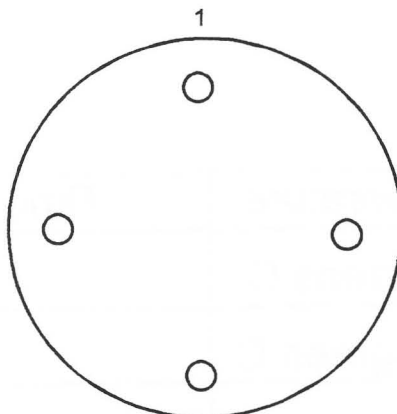
8/ Fill hole number 4 with diluted ethanol.

9/ Place the cover on the petri dish and leave in a warm, dark place for 48 hours.

10/ IN THE FUME HOOD -Remove the lid and mop up any condensation with a tissue held with forceps. Replace the lid and inspect the culture at your work bench.

11/ Draw the appearance of the culture plate on the diagram below.

Waste bucket in fume hood



Result: Bacteria and fungi grew prolifically over the plate but left clear rings around holes 2,3, and 4. Some grow was evident close to hole 1.

Conclusion: Bleach, phenol antiseptic and ethanol were all effective at preventing microbe growth. Salty water was partially effective at preventing microbe growth.

Risk Level: Moderate Hazard: Biohazard risk. All cultures should be treated as a potential source of pathogens. Tissue wipes and the cultures should be collected in a bucket in the fume hood before being sealed in a garbage bag for disposal. Students should be warned not to open the cultures except in the fume hood. All students should wash their hands before leaving the room.

Enzymes 1

Aim: To examine the temperature dependence of the catalytic action of an enzyme.

Equipment

- Rennin tablets
- Stirring rod
- Beakers, 250ml, four
- Dropper bottle
- Test tubes, medium, four
- Hot water
- Thermometer
- Glass stirring rod

Procedure

Rennin (chymase) is an enzyme secreted in the stomach of young animals to aid the digestion of milk. Rennin converts milk from a liquid into a gel.

1/ Prepare four beakers, one containing ice water, one containing tap water adjusted to 20 °C, one containing tepid water at 35 °C and one containing hot water at 50 °C.

2/ Place a medium test tube containing 5mls of milk into each beaker and allow 2 minutes for the milk to reach temperature.

3/ Dissolve one rennin tablet in 20mls of water in a dropper bottle then add 10 drops to each test tube, stirring briefly.

4/ Record the time taken in each beaker for a clear layer to begin to develop at the base of the milk.

| Temperature | Time |
|--------------|------|
| 4 degrees C | |
| 20 Degrees C | |
| 35 degrees C | |
| 50 degrees C | |

Results: _____

Conclusion: _____

Enzymes 1

Topics: Inside Life Biological Chem

Aim: To examine the temperature dependence of the catalytic action of an enzyme.

Equipment

Rennin tablets
Stirring rod
Beakers, 250ml, four
Dropper bottle
Test tubes, medium, four
Hot water
Thermometer
Glass stirring rod

Procedure

Rennin (chymase) is an enzyme secreted in the stomach of young animals to aid the digestion of milk. Rennin converts milk from a liquid into a gel.

- 1/ Prepare four beakers, one containing ice water, one containing tap water adjusted to 20 °C, one containing tepid water at 35 °C and one containing hot water at 50 °C.
- 2/ Place a medium test tube containing 5mls of milk into each beaker and allow 2 minutes for the milk to reach temperature.
- 3/ Dissolve one rennin tablet in 20mls of water in a dropper bottle then add 10 drops to each test tube, stirring briefly.
- 4/ Every minute, check the milk in each tube by touching a glass stirring rod to the surface. If a droplet of milk adheres, the milk is still liquid.
- 5/ Record the time taken in each beaker for the milk to gel.

The teacher must decide when to cease the experiment. Tepid milk will gel in 10minutes, iced milk may take an hour.

Result: The rennin acts fastest at 35 degrees, slower at 20 degrees and much slower in ice or hot water.

Conclusion: Rennin is a biological catalyst which works best close to normal body temperatures.

Risk Level: Low Hazard



Enzymes 2

Aim: To examine the effect of denaturation on an enzyme.

Equipment

- Thermometer
- Rennin tablets
- Dropper bottle
- Beakers, 250ml, two
- Test tubes, medium, four
- Tripod
- Bunsen
- Milk
- Measuring cylinder, 10ml
- Marking pen
- Hot water
- Dropper Bottles:
- Hydrochloric acid, 1M, 10%
- Sodium Hydroxide, 1M, 4%
- Glass stirring rod

Procedure

- three dimensional structure of enzymes (or any protein) can be permanently altered.
- 1/ Heat 50mls of water in a beaker to 60 degrees centigrade.
 - 2/ Dissolve a rennin tablet in 20ml of water in a dropper bottle and add 5 drops to each of four test tubes.
 - 3/ Add 10 drops of water to two of the tubes and markr these tubes as 1 and 2.
 - 4/ Place tube number 2 in the beaker of hot water.
 - 5/ Take another tube of rennin, mark it as 3 and add 5 drops of hydrochloric acid.
 - 6/ Mark the last tube of rennin as 4 and add 5 drops of sodium hydroxide.
 - 7/ After 2 minutes, nuetralise the acid in tube 3 by adding 5 drops of sodium hydroxide and nuetralise the base in tube 4 by adding 5 drops of acid.
 - 8/ Place all four tubes in a beaker containing 100mls of tepid water(about 35 degrees), add 5mls of milk to each and record the time needed for a clear layer to develop.

| Tube | Treatment | Time |
|------|---------------------|------|
| 1 | None | |
| 2 | 2 mins 60 degrees C | |
| 3 | 2 mins Acid | |
| 4 | 2 mins Base | |

Results: _____

Conclusion: _____

Enzymes 2

Topics: Inside Life Biological Chem

Aim: To examine the effect of denaturation on an enzyme.

Equipment

Thermometer
 Rennin tablets
 Dropper bottle
 Beakers, 250ml, two
 Test tubes, medium, four
 Tripod
 Bunsen
 Milk
 Measuring cylinder, 10ml
 Marking pen
 Hot water
 Dropper Bottles:
 Hydrochloric acid, 1M, 10%
 Sodium Hydroxide, 1M, 4%
 Glass stirring rod

Teacher will have to decide when to cease the experiment.

Procedure

Denaturation: Under extremes of temperature and pH, the three dimensional structure of enzymes (or any protein) can be permanently altered.

- 1/ Heat 50mls of water in a beaker to 60 degrees centigrade.
- 2/ Dissolve a rennin tablet in 20ml of water in a dropper bottle and add 10 drops to each of four test tubes.
- 3/ Add 10 drops of water to two of the tubes and mark these tubes as 1 and 2.
- 4/ Place tube number 2 in the beaker of hot water.
- 5/ Take another tube of rennin, mark it as 3 and add 5 drops of hydrochloric acid.
- 6/ Mark the last tube of rennin as 4 and add 5 drops of sodium hydroxide.
- 7/ After 2 minutes, neutralise the acid in tube 3 by adding 5 drops of sodium hydroxide and neutralise the base in tube 4 by adding 5 drops of acid.
- 8/ Place all four tubes in a beaker containing 100mls of tepid water(about 35 degrees), add 5mls of milk to each and test every minute whether surface milk will cling to a stirring rod.

| Tube | Treatment | Time |
|------|---------------------|------|
| 1 | None | |
| 2 | 2 mins 60 degrees C | |
| 3 | 2 mins Acid | |
| 4 | 2 mins Base | |

Result: Tube 1 formed into a gel in about 10 minutes, indicating activity of the rennin enzyme, however tubes 2, 3 and 4 showed little or no enzyme activity.

Conclusion: Rennin can have its enzymic activity permanently impaired through denaturation caused by high temperatures (above 40 degrees centigrade) or extremes of pH, either acid or alkaline.

Risk Level: Low Hazard: Sodium hydroxide is caustic and Hydrochloric acid is corrosive. Any skin contact with these reagents should be treated with prolonged washing in water.

STUDENT: _____

266

Pasteurisation

Aim: To demonstrate that decay is not spontaneous but requires invisible microbes from the air.

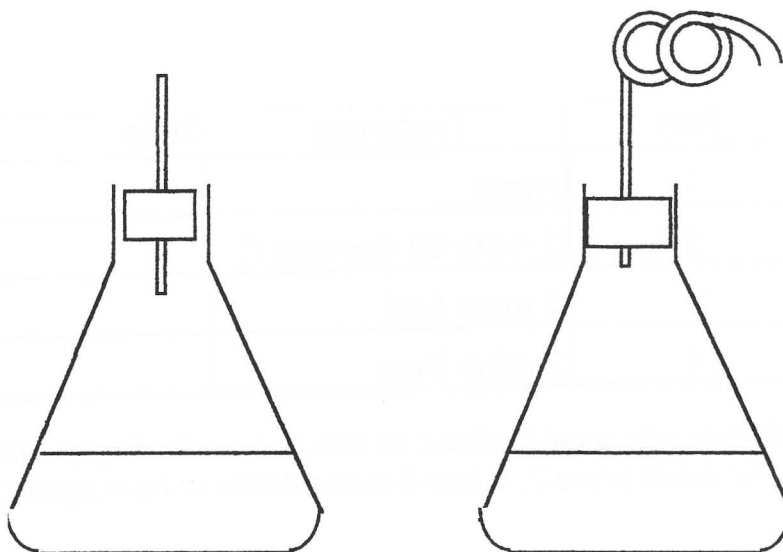
Equipment

Minced Beef
Conical Flasks, 250ml, two
Rubber Stoppers, two (with tubing hole)
Glass tubing 8mm X 8cm and 8mm X 25cm
Bunsen
Tripod
Vacuum grease

Pasteur pipettes, two
Microscope slides, two
Cover slips, two
Microscope

Procedure

- 1/ Add a teaspoon of mince to each conical flask.
- 2/ Add 100ml of water to each flask.
- 3/ Heat the long glass tubing over a Bunsen, rolling and moving a 10cm section over the flame until the glass softens. Bend the glass into a loop.
- 4/ Heat the next 10cm section of tubing to create a second loop. When both loops are complete the ends of the tube should be pointing in the same direction.
- 5/ Lightly grease one end of the tube and carefully insert it into the hole in the rubber stopper.
- 6/ Insert the stoppers into the flasks.
- 7/ Grease and insert the short tubing into the other stopper.
- 7/ Heat each flask until the mixture boils for two minutes.
- 8/ Place both flasks on a shelf for 3 days.
- 9/ Use a Pasteur pipette to draw liquid from one flask and prepare a microscope slide. Repeat for the other flask.
- 10/ Examine both slides under a microscope at 400X magnification.



Results: _____

Conclusion: _____

Pasteurisation

Topics: Inside Life Consumer Science

Aim: To demonstrate that decay is not spontaneous but requires invisible microbes from the air.

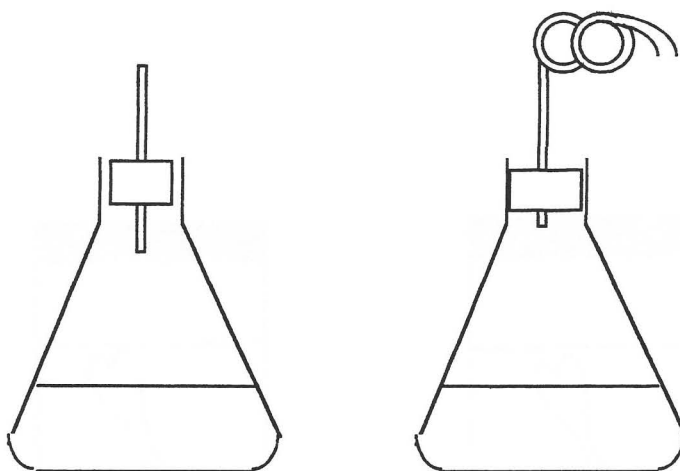
Equipment

Minced Beef
Conical Flasks, 250ml, two
Rubber Stoppers, two (with tubing hole)
Glass tubing 8mm X 8cm and 8mm X 25cm
Bunsen
Tripod
Vacuum grease

Pasteur pipettes, two
Microscope slides, two
Cover slips, two
Microscope

Procedure

- 1/ Add a teaspoon of mince to each conical flask.
- 2/ Add 100ml of water to each flask.
- 3/ Heat the long glass tubing over a Bunsen, rolling and moving a 10cm section over the flame until the glass softens. Bend the glass into a loop.
- 4/ Heat the next 10cm section of tubing to create a second loop. When both loops are complete the ends of the tube should be pointing in the same direction.
- 5/ Lightly grease one end of the tube and carefully insert it into the hole in the rubber stopper.
- 6/ Insert the stoppers into the flasks.
- 7/ Grease and insert the short tubing into the other stopper.
- 7/ Heat each flask until the mixture boils for two minutes.
- 8/ Place both flasks on a shelf for 3 days.
- 9/ Use a Pasteur pipette to draw liquid from one flask and prepare a microscope slide. Repeat for the other flask.
- 10/ Examine both slides under a microscope at 400X magnification.



Result: The fluid in the flask with straight tubing was full of bacteria while the fluid in the other flask was sterile.

Conclusion: Boiling the flasks killed any microbes already inside. The flask with twisted glass tubing was sealed by condensed steam which formed a water loop. After three days microbes had entered the flask with straight tubing and multiplied. The mixture in the sealed flask was free of microbes and preserved by pasteurisation (boiling and sealing).

Risk Level: Moderate Hazard: Caution students only to handle the ends of the heated glass tubing and not to pick up any tubing which may still be hot. Great care should be used while inserting glass tubing into stoppers since the tubing easily shatters. Heavy garden gloves may be advisable.

STUDENT: _____

267

Rubber Bones

Aim: To observe the collagen matrix of bones by dissolving the carbonate component of the bones.

Equipment

Chicken bones (leg or wish bones)

Vinegar

Jars, 200ml, two

Procedure

1/ Place one chicken bone in a jar of water.

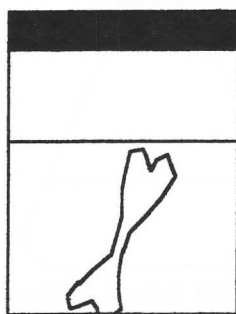
2/ Place another chicken bone in a jar of vinegar.

3/ Replace the vinegar every 2 days for a week.

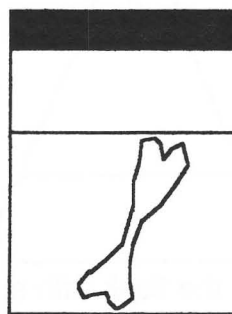
4/ Remove and wash the bones then compare their rigidity.

5/ Compare the strength of the water soaked bone to the strength of a stick of chalk.

- write a general equation for the reaction of the acid vinegar with the carbonates in the bone.



Water



Vinegar

Results: _____

Conclusion: _____

Rubber Bones

Topics: Inside Life Acids & bases

Aim: To observe the collagen matrix of bones by dissolving the carbonate component of the bones.

Equipment

Chicken bones (leg or wish bones)

Vinegar

Jars, 200ml, two

Procedure

1/ Place one chicken bone in a jar of water.

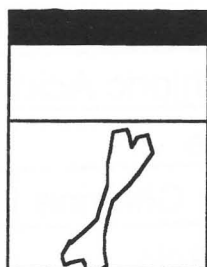
2/ Place another chicken bone in a jar of vinegar.

3/ Replace the vinegar every 2 days for a week.

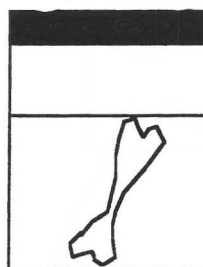
4/ Remove and wash the bones then compare their rigidity.

5/ Compare the strength of the water soaked bone to the strength of a stick of chalk.

- write a general equation for the reaction of the acid vinegar with the carbonates in the bone.



Water



Vinegar

Result: The bone soaked in vinegar was flexible enough to tie in a knot.

Conclusion: The acid in the vinegar reacted with the carbonate deposits in the bone (acid + carbonate > salt + water + carbon dioxide). The bone still looked whole due to the remaining collagen matrix, however this matrix is flexible. Carbonate deposits make the bone rigid but the collagen matrix gives the bone much greater strength than a stick of chalk.

Risk Level: Low Hazard:

Electrolytes

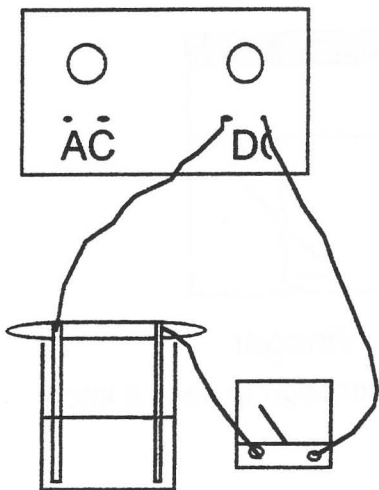
Aim: To measure the conductivity of various solutions of the same concentration as a means of comparing the degree of ionisation.

Equipment

- Paddle pop stick
- Alligator clips,two
- Beaker 250 ml
- Connecting wires,three
- Milli ammeter (or multimeter)
- Copper electrodes,two
- Steel wool
- Power supply, 0-12V DC
- Test Solutions in beakers:
- Hydrochloric acid 0.1M,1%
- Sucrose, 0.1M, 3.4%
- Acetic acid, 0.1M, 0.6%
- Sodium Chloride,0.1M 0.6%
- Ethanol, 0.1M, 0.6%
- Ammonia, 0.1M, 0.7%
- Nitric Acid ,0.1M, 0.6%

Procedure

- Substances dissolve in water by reducing to individual molecules and ions. Not all substances completely dissociate into ions. Since ions conduct electricity in a solution, the amount of current flow should reflect the degree of ionisation.
- 1/ Clean two copper electrodes with steel wool.
 - 3/ Insert the electrodes into the solution on opposite sides of a beaker of test solution. Hold the electrodes in place by fastenning to a paddle pop stick with alligator clips.
 - 4/ Connect one alligator clip to the positive terminal of the ammeter.
 - 5/ Connect the negative ammeter terminal to the negative DC terminal of the power supply.
 - 6/ Connect the remaining alligator clip to the positive DC terminal of the power supply.
 - 7/ Set the power supply to 2V and record the meter reading.
 - 8/ Rinse the electrodes in a beaker of water then measure the current flow in another test solution.
 - 9/ Repeat step 8 until all the solutions have been tested.



| Solution | Current |
|-------------------|---------|
| Hydrochloric Acid | |
| Sucrose | |
| Sodium Chloride | |
| Acetic Acid | |
| Ethanol | |
| Ammonia | |
| Nitric Acid | |

Results: _____

Conclusion: _____

Electrolytes

Topics:

Ions

Acids & Bases

Aim: To measure the conductivity of various solutions of the same concentration as a means of comparing the degree of ionisation.

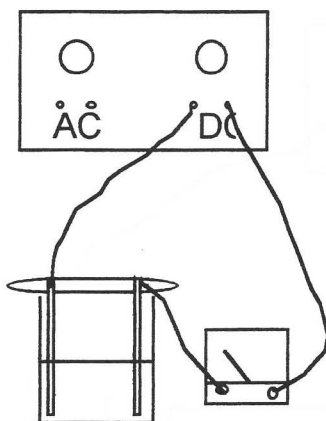
Equipment

Paddle pop stick
 Alligator clips, two
 Beaker 250 ml
 Connecting wires, three
 Milli ammeter (or multimeter)
 Copper electrodes, two
 Steel wool
 Power supply, 0-12V DC
 Test Solutions in beakers:
 Hydrochloric acid 0.1M, 1%
 Sucrose, 0.1M, 3.4%
 Acetic acid, 0.1M, 0.6%
 Sodium Chloride, 0.1M 0.6%
 Ethanol, 0.1M, 0.6%
 Ammonia, 0.1M, 0.7%
 Nitric Acid, 0.1M, 0.6%

Procedure

Substances dissolve in water by reducing to individual molecules and ions. Not all substances completely dissociate into ions. Since ions conduct electricity in a solution, the amount of current flow should reflect the degree of ionisation.

- 1/ Clean two copper electrodes with steel wool.
- 3/ Insert the electrodes into the solution on opposite sides of a beaker of test solution. Hold the electrodes in place by fastening to a paddle pop stick with alligator clips.
- 4/ Connect one clip to the positive terminal of the ammeter.
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- 8/ Rinse the electrodes in a beaker of water then measure the current flow in another test solution.
- 9/ Repeat step 8 until all the solutions have been tested.



| Solution | Current |
|-------------------|---------|
| Hydrochloric Acid | |
| Sucrose | |
| Sodium Chloride | |
| Acetic Acid | |
| Ethanol | |
| Ammonia | |
| Nitric Acid | |

Result: Hydrochloric acid, sodium chloride and nitric acid gave high conductivity readings. Acetic acid and ammonia passed a reduced current while sucrose and ethanol passed very little current.

Conclusion: Hydrochloric acid, sodium chloride and nitric acid completely dissociate into ions. Ammonia and acetic acid only partially dissociate into ions. Ethanol and sucrose remain mostly in a molecular form in solution.

Note : All the test solutions are monoprotic.

Risk Level: Low Hazard: All the test solutions are of low toxicity and minimally corrosive at these concentrations.

STUDENT: _____

269

Electrophoresis

Aim: To demonstrate that ions move under the influence of an electric current.

Equipment

Microscope slides, two

Filter papers, two

scissors

Lead Nitrate, 1M (33%) in a dropper bottle.

Potassium Iodide, 1M (17%) in dropper bottle.

Connecting leads, two

Power supply, 0-12V DC

Procedure

1/ Trace the outline of a glass slide on each filter paper.

2/ Cut out the traced rectangle (removing the drawn lines).

3/ Place a paper rectangle on each slide.

4/ Moisten the paper rectangles with a few drops of water.

5/ Place a drop of Potassium iodide solution on the right hand side of each slide.

6/ Place a drop of lead nitrate on the left hand side of each slide.

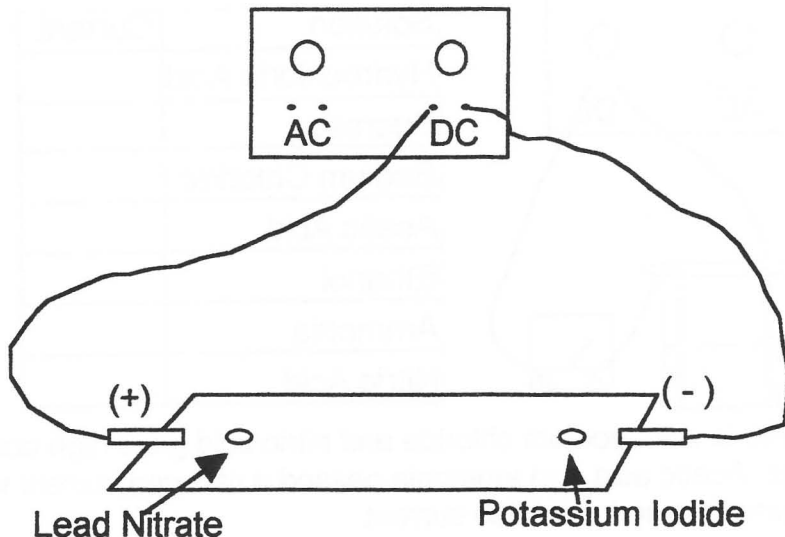
7/ Fasten alligator clips to each end of **one** slide.

8/ Connect the left clip to the positive DC terminal of the power supply.

9/ Connect the right clip to the negative DC terminal of the power supply.

10/ Set the voltage to 12V DC and turn on the power for about 40 minutes.

Lead iodide is a yellow and insoluble.



Results: _____

Conclusion: _____

Electrophoresis

Topics: Ions

Aim: To demonstrate that ions move under the influence of an electric current.

Equipment

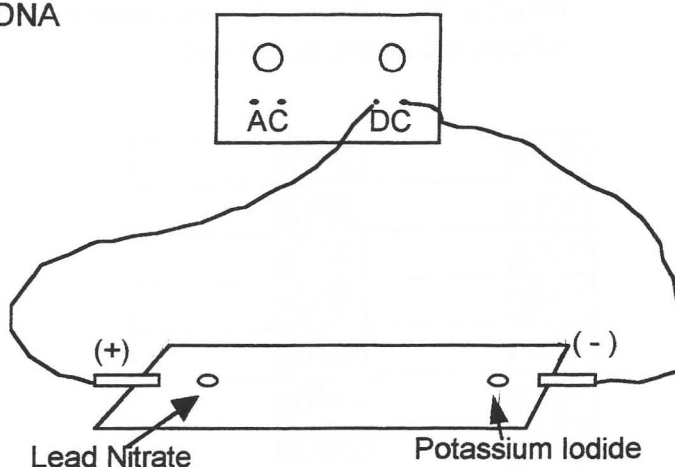
Microscope slides, two
Filter papers, two
scissors
Lead Nitrate, 1M (33%) in a dropper bottle.
Potassium Iodide, 1M (17%) in dropper bottle.
Connecting leads, two
Power supply, 0-12V DC

Procedure

- 1/ Trace the outline of a glass slide on each filter paper.
- 2/ Cut out the traced rectangle (removing the drawn lines).
- 3/ Place a paper rectangle on each slide.
- 4/ Moisten the paper rectangles with a few drops of water.
- 5/ Place a drop of Potassium iodide solution on the right hand side of each slide.
- 6/ Place a drop of lead nitrate on the left hand side of each slide.
- 7/ Fasten alligator clips to each end of **one** slide.
- 8/ Connect the left clip to the positive DC terminal of the power supply.
- 9/ Connect the right clip to the negative DC terminal of the power supply.
- 10/ Set the voltage to 12V DC and turn on the power for about 40 minutes.

Moving ions in a conducting strip is called electrophoresis and is a central technique in DNA finger printing.

Lead iodide is a yellow and insoluble.



Result: No change was seen in the slide without applied current. A yellow strip appeared near the centre of the slide with applied current.

Conclusion: Under the influence of an electric potential negative iodine ions moved toward the positive clip (anode) while positive lead ions moved toward the negative electrode (cathode). Where the lead and iodine ions met they formed a band of yellow lead iodide.

Risk Level: Mild Hazard: Lead nitrate is a cumulative toxin and ingestion or skin contact should be avoided. Potassium iodide is harmful if continually ingested.

STUDENT: _____

270

Dissolving

Aim: To determine how much sugar can be dissolved in 80ml of water.

Equipment

Sugar
Beaker, 150ml
Stirring Rod
Balance, 0.1g
Measuring cylinder
Paper cup

Procedure

- 1/ Measure 80mls of water into the beaker.
- 2/ Weigh 100gms of sugar in a paper cup and add it to the water.
- 3/ Stir until the sugar has dissolved
- 4/ Weigh 20g of sugar and add it to the water.
- 5/ Stir the sugar until it dissolves.
- 6/ If the sugar dissolves make a tally mark in the table below.
- 7/ Keep repeating the last three steps until no more sugar will dissolve.
- 8/ Pour the solution into the measuring cylinder and record

its volume.

- The sugar has dissolved forming a SOLUTION of sugar and water. In this solution the sugar is called the SOLUTE and the water is called the SOLVENT (a liquid which can dissolve substances).

- How much solute is in this solution?
- How much solvent is in the solution?
- What is the volume of the solution?
- Where did all the sugar go?

| | |
|-------|------|
| Total | 240g |
| 100g | 260g |
| 120g | 280g |
| 140g | 300g |
| 160g | 320g |
| 180g | 340g |
| 200g | 360g |
| 220g | 380g |

| | |
|--------------|--|
| Final Volume | |
|--------------|--|

Results: _____

Conclusion: _____

Dissolving

Topics: Kinetic Theory

Matter

Aim: To determine how much sugar can be dissolved in 80ml of water.**Equipment**

Sugar
Beaker, 150ml
Stirring Rod
Balance, 0.1g
Measuring cylinder
Paper cup

Procedure

- 1/ Measure 80mls of water into the beaker.
- 2/ Weigh 100gms of sugar in a paper cup.
- 3/ Add the sugar to the water, stirring until it all has dissolved.
- 4/ Weigh 10g of sugar and add it to the water.
- 5/ Stir the sugar until it dissolves.
- 6/ If the sugar dissolves make a tally mark in the table below.
- 7/ Keep repeating the last three steps until no more sugar will dissolve.

8/ Pour the solution into the measuring cylinder and record its volume.

- The sugar has dissolved forming a SOLUTION of sugar and water. In this solution the sugar is called the SOLUTE and the water is called the SOLVENT (a liquid which can dissolve substances).

- How much solute is in this solution?
- How much solvent is in the solution?
- What is the volume of the solution?
- Where did all the sugar go?

A surprisingly good experiment for year 7.

Result: Over 200 g of sugar dissolved in the water yet the volume of the solution only increased slightly.

Conclusion: Sugar is the solute and water is the solvent. Solid crystals of sugar are composed of sugar molecules in a regular pattern. In water the sugar molecules separate mixing among and between the water molecules. The volume of the solution increases only slightly because the solid sugar is mostly empty space and there is much empty space in the water.

Risk Level: Low Hazard:

STUDENT: _____

271

Hot Cooling

Aim: To observe a latent energy release during a change of state.

Equipment

Test Tube, medium
Sodium Thiosulfate
Thermometer, 0 -110 C,
Bunsen
Retort stand
Clamp and boss head

Procedure

- 1/ Place two cm of sodium thiosulfate (hypo) in the test tube.
- 2/ Mount the test tube in a clamp on the retort stand above the bunsen.
- 3/ Adjust the bunsen to a clear flame and heat the test tube gently until the sodium thiosulfate melts. **DO NOT OVER HEAT.**
- 4/ Carefully place the thermometer into the molten thiosulfate.
- 5/ Record the temperture every 15 seconds as the compound cools.
- 6/ Convert the results into a graph of temperature versus time.

- Could your results explain why people say it is warmer while snow is falling?
- Gently submerge the test tubes and thermometers in a washing tray to dissolve the thiosulfate.

| Time (mins) | Temp. | Time (mins) | Temp |
|-------------|-------|--------------|------|
| 0.00 | | 2.50 | |
| 0.25 | | 2.75 | |
| 0.50 | | 3.00 | |
| 0.75 | | 3.25 | |
| 1.00 | | 3.50 | |
| 1.25 | | 3.75 | |
| 1.50 | | 4.00 | |
| 1.75 | | 4.25 | |
| 2.00 | | 4.50 | |
| 2.25 | | 4.75 | |

Results: _____

Conclusion: _____

Hot Cooling

Topics: Kinetic Theory

Aim: To observe a latent energy release during a change of state.

Equipment

Test Tube, medium
Sodium Thiosulfate
Thermometer, 0 -110 C,
Bunsen
Retort stand
Clamp and boss head

Procedure

- 1/ Place two cm of sodium thiosulfate (hypo) in the test tube.
- 2/ Mount the test tube in a clamp on the retort stand above the Bunsen.
- 3/ Adjust the Bunsen to a clear flame and heat the test tube gently until the sodium thiosulfate melts. DO NOT OVER HEAT.
- 4/ Carefully place the thermometer into the molten thiosulfate.
- 5/ Record the temperature every 30 seconds as the compound cools.
- 6/ Convert the results into a graph of temperature versus time.

- Could your results explain why people say it is warmer while snow is falling?
- Gently submerge the test tubes and thermometers in a washing tray to dissolve the thiosulfate.

Provide a long washing tray of water to dissolve the solidified hypo.

Result: Initially the temperature fell, then rose as crystals formed then finally began to fall once again.

Conclusion: As a liquid changes state into a solid it releases energy called latent heat. As water vapour freezes into snow flakes it also releases latent heat which temporarily warms the surrounding air. Latent heat is a major energy source in cyclones and hurricanes. Such storms weaken over dry land where there is less water vapour to condense into clouds producing latent heat.

Risk Level: Moderate Hazard: Sodium thiosulfate can react violently with oxidising agents. Overheating sodium thiosulfate may produce corrosive sulfur dioxide fumes.

STUDENT: _____

272

Concave Lens

Aim: To observe and record the effect of a concave lens on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Concave lens(flat edge,
vertically standing)
Concave lens, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged lens on the piece of paper.
- 2/ Trace the outline of the lens.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the lens.
- 6/ Make some pencil dots to mark the path of the rays toward the lens and the refracted rays emerging.
- 7/ Join the dots with a ruler.
- 8/ Measure from the lens to where the rays converge (focus)
- 9/ Move the lens slightly below the first drawing
- 10/ Aim the light rays onto the lens but at an angle then repeat steps 2, 6 and 7.
- 11/ Paste both drawings in the space below.
- 12/ Write the word "Normal" on a strip of paper.
- 13/ Draw how the letters appear through the circular lens.
- 14/ Move the paper further away and draw the image.

Results: _____

Conclusion: _____

Concave Lens

Topics: Light

Waves

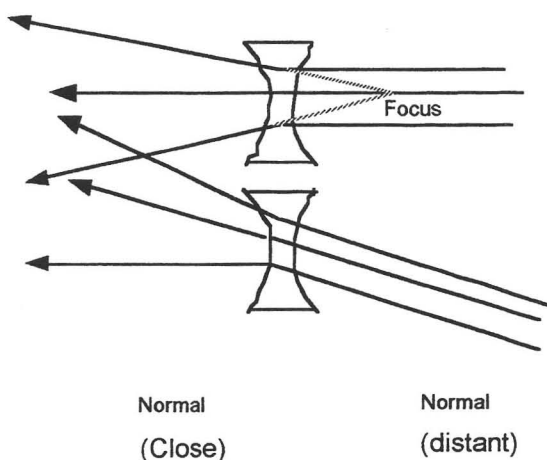
Aim: To observe and record the effect of a concave lens on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Concave lens(flat edge, vertically standing)
Concave lens, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged lens on the piece of paper.
- 2/ Trace the outline of the lens.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the lens.
- 6/ Make some pencil dots to mark the path of the rays toward the lens and the refracted rays emerging.
- 7/ Join the dots with a ruler.
- 8/ Measure from the lens to where the rays converge (focus)
- 9/ Move the lens slightly below the first drawing
- 10/ Aim the light rays onto the lens but at an angle then repeat steps 2, 6 and 7.
- 11/ Paste both drawings in the space below.
- 12/ Write the word "Normal" on a strip of paper.
- 13/ Draw how the letters appear through the circular lens.
- 14/ Move the paper further away and draw the image.



Result: Light rays passing through the lens are spread apart. Letters close to the lens appear smaller. Letters further from the lens appear very small.

Conclusion: Concave lenses refract parallel light rays so they spread apart as if they originated from the same point. Images seen through these lenses appear smaller and further away.

Risk Level: Low Hazard

STUDENT: _____

273

Concave Mirror

Aim: To observe and record the effect of a concave mirror on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Concave mirror (flat edge,
vertically standing)
Concave mirror, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged mirror on the piece of paper.
- 2/ Trace the outline of the mirror.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the mirror.
- 6/ Make some pencil dots to mark the path of the rays toward the mirror (Incident Rays) and the Reflected Rays.
- 7/ Join the dots with a ruler.
- 8/ Move the mirror slightly below the first drawing
- 9/ Aim the light rays onto the mirror but at an angle then repeat steps 6 and 7.
- 10/ Measure from the mirror to where the lines meet . Record this on the diagram as the "focal length."
- 10/ Paste both drawings in the space below.
- 11/ Write the word "Normal" on a strip of paper.
- 12/ Place the paper a few centimetres from the round mirror.
- 13/ Draw how the letters appear in the mirror.
- 14/ Move the paper one metre away and draw the image.

Results: _____

Conclusion: _____

Concave Mirror

Topics: Light

Waves

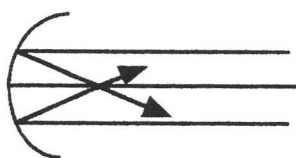
Aim: To observe and record the effect of a concave mirror on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Concave mirror (flat edge, vertically standing)
Concave mirror, circular
Sheet of Paper
Paper glue

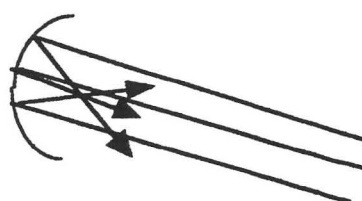
Procedure

- 1/ Place the flat edged mirror on the piece of paper.
- 2/ Trace the outline of the mirror.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the mirror.
- 6/ Make some pencil dots to mark the path of the rays toward the mirror (Incident Rays) and the Reflected Rays.
- 7/ Join the dots with a ruler.
- 8/ Move the mirror slightly below the first drawing
- 9/ Aim the light rays onto the mirror but at an angle then repeat steps 6 and 7.
- 10/ Measure from the mirror to where the lines meet . Record this on the diagram as the "focal length."
- 10/ Paste both drawings in the space below.
- 11/ Write the word "Normal" on a strip of paper.
- 12/ Place the paper a few centimetres from the round mirror.
- 13/ Draw how the letters appear in the mirror.
- 14/ Move the paper one metre away and draw the image.



Normal

(Close)



Normal

(1 metre)

Result: The parallel rays are always reflected to meet at a point. Letters close to the mirror appear normal but magnified. Letters one metre from the mirror appear magnified, reversed and inverted.

Conclusion: Convex mirrors reflect parallel light rays to converge at a focus point. Images seen in these mirrors are magnified. Objects beyond the focal point of the mirror are still magnified but inverted both horizontally and vertically. Large images are warped toward the mirror edge.

Risk Level: Low Hazard.

STUDENT: _____

274

Convex Lens

Aim: To observe and record the effect of a convex lens on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Convex mirror (flat edge,
vertically standing)
Convex mirror, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged lens on the piece of paper.
- 2/ Trace the outline of the lens.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the lens.
- 6/ Make some pencil dots to mark the path of the rays toward the lens and the refracted rays emerging.
- 7/ Join the dots with a ruler.
- 8/ Measure from the lens to where the rays meet (focus)
- 9/ Move the lens slightly below the first drawing
- 10/ Aim the light rays onto the lens but at an angle then repeat steps 2, 6 and 7.
- 11/ Paste both drawings in the space below.
- 12/ Write the word "Normal" on a strip of paper.
- 13/ Draw how the letters appear through the circular lens.
- 14/ Move the paper further away and draw the image.

Results: _____

Conclusion: _____

Convex Lens

Topics: Light

Waves

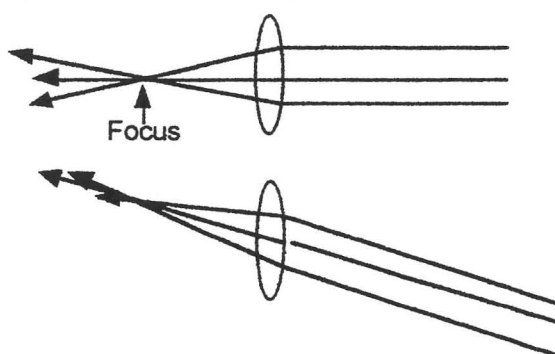
Aim: To observe and record the effect of a convex lens on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Convex mirror (flat edge, vertically standing)
Convex mirror, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged lens on the piece of paper.
- 2/ Trace the outline of the lens.
- 3/ Connect the ray box to the power supply (set at 6V)
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- 11/ Paste both drawings in the space below.
- 12/ Write the word "Normal" on a strip of paper.
- 13/ Draw how the letters appear through the circular lens.
- 14/ Move the paper further away and draw the image.



Normal

(Close)

Normal

(distant)

Result: The rays passing through the lens converge at a point. Letters close to the lens appear magnified. Letters further from the lens are inverted and reversed. At a point close to the focus the letters cannot be seen.

Conclusion: Convex lenses refract parallel light rays to converge at a focus point. Objects close to the lens are magnified. Objects further away appear inverted, magnified at first but less so with increasing distance from the lens.

Risk Level: Low Hazard

STUDENT: _____

275

Convex Mirror

Aim: To observe and record the effect of a convex mirror on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Convex mirror (flat edge,
vertically standing)
Convex mirror, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged mirror on the piece of paper.
- 2/ Trace the outline of the mirror.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the mirror.
- 6/ Make some pencil dots to mark the path of the rays toward the mirror (Incident Rays) and the Reflected Rays.
- 7/ Join the dots with a ruler.
- 8/ Move the mirror slightly below the first drawing
- 9/ Aim the light rays onto the mirror but at an angle then repeat steps 6 and 7.
- 10/ Draw dotted lines extending the reflected rays back beyond the mirror. Wher the lines meet is called the focus.
- 11/ Paste both drawings in the space below.
- 12/ Write the word "Normal" on a strip of paper.
- 13/ Place the paper a few centimetres from the round mirror.
- 14/ Draw how the letters appear in the mirror.
- 15/ Move the paper one metre away and draw the image.

Results: _____

Conclusion: _____

Convex Mirror

Topics: Light

Waves

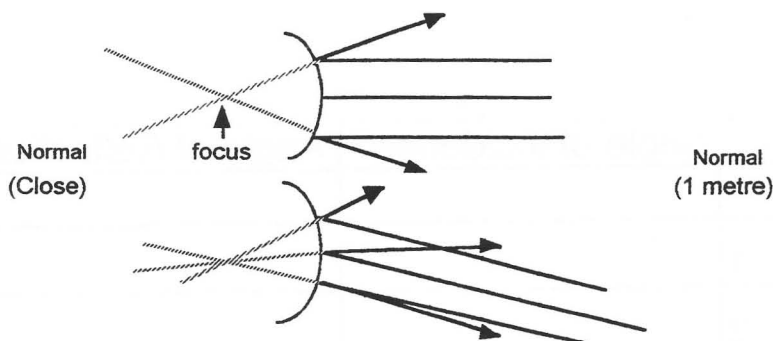
Aim: To observe and record the effect of a convex mirror on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Convex mirror (flat edge, vertically standing)
Convex mirror, circular
Sheet of Paper
Paper glue

Procedure

- 1/ Place the flat edged mirror on the piece of paper.
- 2/ Trace the outline of the mirror.
- 3/ Connect the ray box to the power supply (set at 6V)
- 4/ Insert the triple slit slide in front of the light to produce three parallel rays of light.
- 5/ Aim the rays straight onto the mirror.
- 6/ Make some pencil dots to mark the path of the rays toward the mirror (Incident Rays) and the Reflected Rays.
- 7/ Join the dots with a ruler.
- 8/ Move the mirror slightly below the first drawing
- 9/ Aim the light rays onto the mirror but at an angle then repeat steps 6 and 7.
- 10/ Draw dotted lines extending the reflected rays back beyond the mirror. Where the lines meet is called the focus.
- 11/ Paste both drawings in the space below.
- 12/ Write the word "Normal" on a strip of paper.
- 13/ Place the paper a few centimetres from the round mirror.
- 14/ Draw how the letters appear in the mirror.
- 15/ Move the paper one metre away and draw the image.



Result: Light rays reflected from a convex mirror are scattered. Images seen in the mirrors appear properly oriented but small.

Conclusion: Convex mirrors cause parallel light rays to diverge as if they originated from a point behind the mirror. The images produced are not inverted but are reduced in size. Convex mirrors are useful for a wide angle view eg. security mirrors in stores.

Risk Level: Low Hazard.

Flat Mirror

Aim: To observe and record the effect of a flat mirror on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Flat Mirror
Sheet of Paper
Protractor

Procedure

- 1/ Place the flat mirror on one side of the piece of paper.
- 2/ Trace the outline of the mirror.
- 3/ Rule a line perpendicular to the centre of the mirror face. This line is called the "normal".
- 4/ Connect the ray box to the power supply (set at 6V)
- 5/ Insert the single slit slide in front of the light to produce a ray of light.
- 6/ Aim the light ray where the Normal meets the mirror.
- 7/ Make some pencil dots to mark the path of the ray toward the mirror (Incident Ray) and the Reflected Ray.
- 8/ Join the dots with a ruler then use a protractor to measure the angle between Incident Ray and the normal.
- 9/ Measure the angle between the Reflected ray and the normal. Record both angles in the table below.
- 10/ Repeat the last four steps for two new angles.
- 11/ Write the word "Normal" on a strip of paper.
- 12/ Place the paper a few centimetres from the mirror.
- 13/ Observe how the letters appear in the mirror.

| Line | Angle of Incidence | Angle of Reflection |
|------|--------------------|---------------------|
| | | |
| 1 | | |
| 2 | | |
| 3 | | |

Results: _____

Conclusion: _____

Flat Mirror

Topics: Light

Waves

Aim: To observe and record the effect of a flat mirror on light rays and images.

Equipment

Hudson Ray Box
Power Supply 2-12V, DC
Flat Mirror
Sheet of Paper
Protractor

Procedure

- 1/ Place the flat mirror on one side of the piece of paper.
- 2/ Trace the outline of the mirror.
- 3/ Rule a line perpendicular to the centre of the mirror face. This line is called the "normal".
- 4/ Connect the ray box to the power supply (set at 6V)
- 5/ Insert the single slit slide in front of the light to produce a ray of light.
- 6/ Aim the light ray where the Normal meets the mirror.
- 7/ Make some pencil dots to mark the path of the ray toward the mirror (Incident Ray) and the Reflected Ray.
- 8/ Join the dots with a ruler then use a protractor to measure the angle between Incident Ray and the normal.
- 9/ Measure the angle between the Reflected ray and the normal. Record both angles in the table below.
- 10/ Repeat the last four steps for two new angles.
- 11/ Write the word "Normal" on a strip of paper.
- 12/ Place the paper a few centimetres from the mirror.
- 13/ Draw how the letters appear in the mirror.

16m10M

Result: A flat mirror reflects a light ray at the same angle as it strikes the mirror.
Letters viewed in a mirror appear reversed (but not inverted).

Conclusion: The Law of Reflection: The angle of reflection equals the angle of incidence.
Flat mirrors invert an image only in the horizontal plane.

Risk Level: Low Hazard

STUDENT: _____

277

Arthropod Art

Aim: To classify and preserve a small arthropod in in a resin block.

Equipment

Casting Resin (two batches
of 50ml each)
Plastic egg cup
Floor Polish
Capture Jar
Methanol
Pins, 12
styrofoam 5cm X 5cm
Turpentine
Wet & dry abrasive paper
Brass Polish
Packing Tape (labels)
Match sticks

Procedure

- 1/ Collect a small insect or spider (about 3cm max. diam.)
- 2/ Dehydrate in methanol for at least 1hour.
- 3/ Use pins to mount the specimen in a realistic pose on a labelled styrofoam base and allow to dry for one week.
- 4/ Rinse an egg cup with floor polish and wipe clean.
- 5/ Half fill a plastic egg cup with slow setting casting resin.
- 6/ Dip the dried specimen in turpentine and allow to drain on tissue paper.
- 7/ Use forceps to place the specimen on the soft resin. Use a match stick to gently immerse the specimen.
- 8/ After about 1.5hrs the resin will begin to gel. Pour fresh resin over the specimen to fill the egg cup.
- 9/ Allow the resin to set overnight. Remove the plastic block . Smooth and polish with wet & dry abrasive paper as needed.
- 10/ Final buff with brass polish.

- Use an identification key to fully classify your specimen.
- Write a report on the life cycle, food, habitat, special behaviours and features of your specimen.

Results: _____

Conclusion: _____

Arthropod Art

Topics: Living Things

Aim: To classify and preserve a small arthropod in in a resin block.

Equipment

Casting Resin (two
batches of 50ml each)
Plastic egg cup
Floor Polish
Capture Jar
Methanol
Pins, 12
styrofoam 5cm X 5cm
Turpentine
Wet & dry abrasive paper
Brass Polish
Packing Tape (labels)
Match sticks

Casting Resin: 0.5ml
hardener to 100ml resin.
Use evaporating basins or
spray can lids for larger
specimens.

Procedure

- 1/ Collect a small insect or spider (about 3cm max. diam.)
 - 2/ Dehydrate in methanol for at least 1hour.
 - 3/ Use pins to mount the specimen in a realistic pose on a labelled styrofoam base and allow to dry for one week.
 - 4/ Rinse an egg cup with floor polish and wipe clean.
 - 5/ Half fill a plastic egg cup with slow setting casting resin.
 - 6/ Dip the dried specimen in turpentine and allow to drain on tissue paper.
 - 7/ Use forceps to place the specimen on the soft resin. Use a match stick to gently immerse the specimen.
 - 8/ After about 1.5hrs the resin will begin to gel. Pour fresh resin over the specimen to fill the egg cup.
 - 9/ Allow the resin to set overnight. Remove the plastic block . Smooth and polish with wet & dry abrasive paper as needed.
 - 10/ Final buff with brass polish.
- Use an identification key to fully classify your specimen.
 - Write a report on the life cycle, food, habitat, special behaviours and features of your specimen.

Result: The specimen is preserved in a clear magnifying dome of polymer resin

Conclusion: This activity is time consuming but adds real student interest to the study of invertebrates. Specimen catalogues show a well presented beetle to be valuable. Students can purchase their completed specimen at cost on condition of completing the classification and research assignment.

Risk Level: Mild Hazard: Good ventilation is essential throughout. Students should not mix or handle the liquid resin. Methanol and Turpentine are flammable and harmful if ingested. Skin contact with the liquid resin should be avoided. Use turpentine to clean resin spills. Slow setting resin is used to avoid damage to the specimen from the heat of polymerisation.

STUDENT: _____

278

Gears and Torque

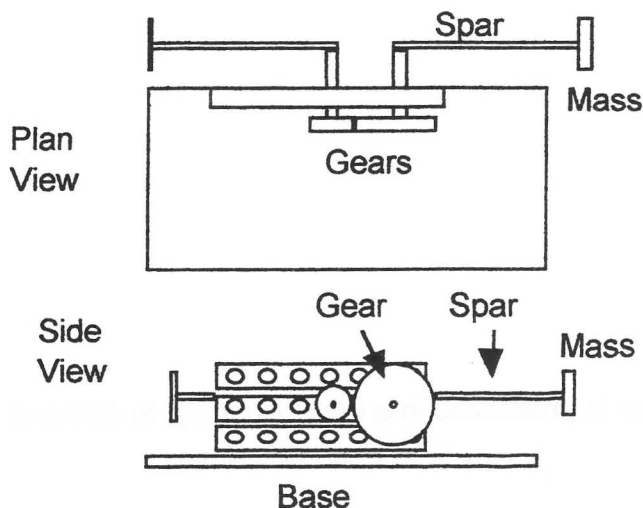
Aim: To examine the mechanical advantage of gears.

Equipment

Lego Technics Set
Masses, 25g (one),
5g (three), 50g (one)

Procedure

- 1/ Build the simple Lego construction drawn below.
- 2/ Count the number of teeth on each gear and record the result in the table.
 - How many times will the small gear wheel turn for each turn of the large gear wheel ?
- 3/ Adjust the axle spars so both are of equal length, horizontal and opposite.
- 4/ Slide the 25g mass onto the end of the spar connected to the axle of the large gear.
- 5/ Add 5g masses to the end of spar connected to the axle of the small gear until the forces balance.
- 6/ Record the result in the table
- 7/ Change the spar on the axle of the small gear so it is twice as long as the spar connected to the large gear.
- 8/ Change the mass on the large gear to 50g and add 5g masses to the small gear spar until the forces balance.
 - Torque is a measure of force applied at an axle. What is the effect of doubling the spar length on torque?



Results: _____

Conclusion: _____

Gears and Torque

Topics: Machines

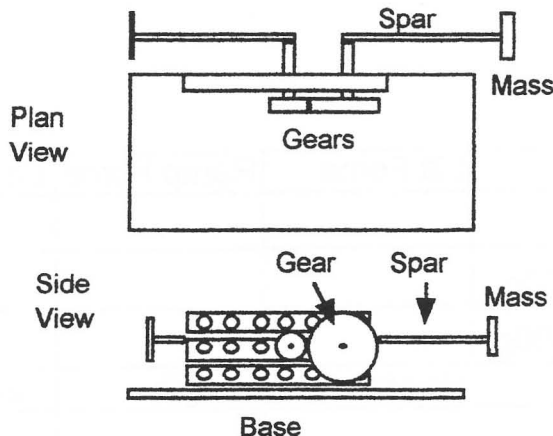
Aim: To examine the mechanical advantage of gears.

Equipment

Lego Technics Set
Masses, 25g (one),
5g (three), 50g (one)

Procedure

- 1/ Build the simple Lego construction drawn below.
- 2/ Count the number of teeth on each gear and record the result in the table.
- How many times will the small gear wheel turn for each turn of the large gear wheel ?
- 3/ Adjust the axle spars so both are equal and opposite.
- 4/ Slide the 25g mass onto the end of the spar connected to the axle of the large gear.
- 5/ Add 5g masses to the end of spar connected to the axle of the small gear until the forces balance.
- 6/ Record the result in the table
- 7/ Change the spar on the axle of the small gear so it is twice as long as the spar connected to the large gear.
- 8/ Change the mass on the large gear to 50g and add 5g masses to the small gear spar until the forces balance.
- Torque is a measure of force applied at an axle. What is the effect of doubling the spar length on torque?



Result: Far less mass needs to be applied to the small gear to balance the large mass on the large gear.

Conclusion: The ratio of teeth numbers on gears is equal to the mechanical advantage supplied by the gears. A force is applied to an axle via a wheel or extended spar is called Torque. Doubling the radius of the wheel or the length of the spar doubles the torque at the axle.

Risk Level: Low Hazard

Inclined Plane

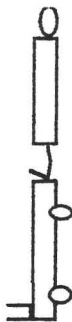
Aim: To measure the force required to raise an object with and without the use of an inclined plane.

Equipment

- Newton Cart
- Spring Balance, 10 N
- Masses, 250g, two
- Metre rule
- Ramp, 1m long

Procedure

- 1/ Use the spring balance to measure the force required to lift the Newton Cart. Record all measurements in the table below.
- 2/ Position the ramp to run from the floor to a chair seat.
- 3/ Measure the height of the chair seat.
- 4/ Measure the length of the ramp.
- 5/ Measure the force required to pull the cart slowly up the ramp.
- 6/ Add one 250g mass to the cart.
- 7/ Measure the force need to pull the cart up the ramp and the force needed to directly lift the cart.
- 8/ Add another 250g mass to the cart and repeat step 7.
- 9/ Complete the calculations in the table.



| Trial | Lift Force | Ramp Force | Lift F./Ramp F. |
|-------------|------------|------------|-----------------|
| Cart | | | |
| Cart +250g | | | |
| Cart + 500g | | | |
| | Height | Length | Height/Length |
| Ramp | | | |

Results: _____

Conclusion: _____

Inclined Plane

Topics: Machines Vectors

Aim: To measure the force required to raise an object with and without the use of an inclined plane.

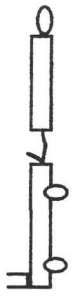
Equipment

Newton Cart
Spring Balance, 10 N
Masses, 250g, two
Metre rule
Ramp, 1m long

Procedure

- 1/ Use the spring balance to measure the force required to lift the Newton Cart. Record all measurements in the table below.
- 2/ Position the ramp to run from the floor to a chair seat.
- 3/ Measure the height of the chair seat.
- 4/ Measure the length of the ramp.
- 5/ Measure the force required to pull the cart slowly up the ramp.
- 6/ Add one 250g mass to the cart.
- 7/ Measure the force need to pull the cart up the ramp and the force needed to directly lift the cart.
- 8/ Add another 250g mass to the cart and repeat step 7.
- 9/ Complete the calculations in the table.

String may be needed if the cart lacks a hook eye for the spring balance.



| Trial | Lift Force | Ramp Force | Lift F./Ramp F. |
|-------------|------------|------------|-----------------|
| Cart | | | |
| Cart +250g | | | |
| Cart + 500g | | | |
| | Height | Length | Height/Length |
| Ramp | | | |

Result: The force required to pull the cart up the ramp was much less than the force needed to lift the cart directly. The ratio of these to forces was equal to the ratio of the length and height of the ramp.

Conclusion: A ramp is a simple machine that reduces the force required to raise an object. The force required to raise the object is reduced in proportion to the length of the ramp divided by its height.

Risk Level: Low Hazard:

Pulleys

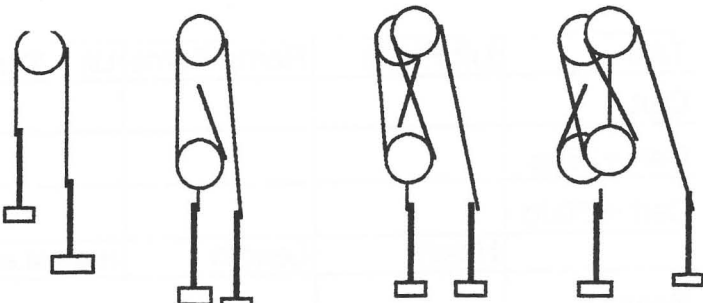
Aim: To examine pulleys as force magnifying machines.

Equipment

Single pulleys, two
Double wheel pulleys, two
Retort stand,
Clamp and boss head,
Pulley string, 1.5m
Mass Carriers, two
Masses, 50g (seven)
25g(two), 5g (two)

Procedure

- 1/ Use the retort stand to support a single pulley over the edge of the bench.
- 2/ Pass the string through the pulley.
- 3/ Loosely tie a mass carrier at one end and add four 50g masses ie. total mass 250g
- 4/ Loosely tie the other mass carrier to the opposite end of the string. Add masses until the carriers balance. Record the balancing mass in the table below.
- 5/ Untie the string and pass it through two single pulleys as shown below.
- 6/ Record the mass required to balance 250g.
- 7/ Untie the string and pass it through one single pulley and one double pulley as shown.
- 8/ Record the mass required to balance 250g.
- 9/ Untie the string and pass it through two double pulleys as shown.
- 10/ Record the mass required to balance 250g.
- 11/ Complete the table.



| | | | | |
|-------------|------|------|------|------|
| Pulleys | 1 | 2 | 3 | 4 |
| Load | 250g | 250g | 250g | 250g |
| Mass | | | | |
| Load / Mass | | | | |

Results: _____

Conclusion: _____

Pulleys

Topics: Machines

Work

Aim: To examine pulleys as force magnifying machines.

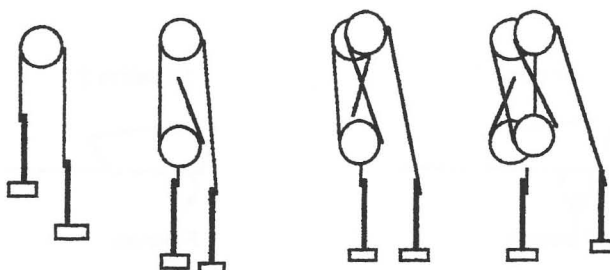
Equipment

Single pulleys, two
Double wheel pulleys, two
Retort stand,
Clamp and boss head,
Pulley string, 1.5m
Mass Carriers, two
Masses, 50g (seven)
25g(two), 5g (two)

Procedure

- 1/ Use the retort stand to support a single pulley over the edge of the bench.
- 2/ Pass the string through the pulley.
- 3/ Loosely tie a mass carrier at one end and add four 50g masses ie. total mass 250g
- 4/ Loosely tie the other mass carrier to the opposite end of the string. Add masses until the carriers balance. Record the balancing mass in the table below.
- 5/ Untie the string and pass it through two single pulleys as shown below.
- 6/ Record the mass required to balance 250g.
- 7/ Untie the string and pass it through one single pulley and one double pulley as shown.
- 8/ Record the mass required to balance 250g.
- 9/ Untie the string and pass it through two double pulleys.
- 10/ Record the mass required to balance 250g.
- 11/ Complete the table.

If equipment is in short supply have some of the class do "Gears" or "Inclined Plane".



| Pulleys | 1 | 2 | 3 | 4 |
|-------------|------|------|------|------|
| Load | 250g | 250g | 250g | 250g |
| Mass | | | | |
| Load / Mass | | | | |

Result: One pulley did not reduce the mass required to balance 250g. Increasing the number of pulleys decreased the mass required to balance 250g.

Conclusion: Pulleys as simple force magnifying machines, the force applied being magnified by the number of pulley wheels. That is the mechanical advantage of a pulley system is equal to the number of pulley wheels. Point out to the students that for a mechanical advantage of 4, the applied force must move four times further than the weight.

Risk Level: Low Hazard:

STUDENT: _____

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Wheel barrow

Aim: To examine the properties of a wheel barrow as a second order lever.

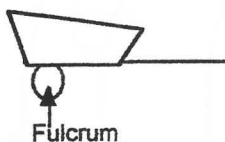
Equipment

Wheel barrow
Two concrete blocks or 8
house bricks

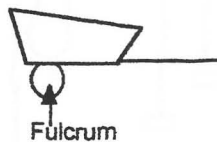
Procedure

- 1/ Place the bricks in the barrow over the wheel .
- 2/ Lift the barrow by the end of the handles.
- 3/ On the diagram provided draw the position of the load with a downward arrow and the position of your hands with an upward arrow.
- 4/ On the line below the diagram record whether the lift is easy or hard.
- 5/ Lower the barrow and slide your hands along the handles to the bucket. Lift the barrow again.
- 6/ Record the load and effort positions on the next diagram and whether the lift is easier or harder.
- 7/ Move the bricks close to the handles.
- 8/ Lift with your hands close to the bucket.
- 9/ Repeat step 7.
- 10/ Lift by the end of the handles and repeat step 7.
- 11/ Compare the relative distance of the load and effort from the fulcrum in each diagram.

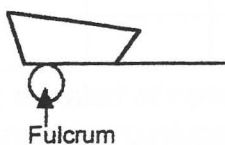
Situation 1



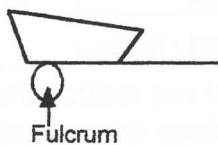
Situation 2



Situation 3



Situation 4



Results: _____

Conclusion: _____

Wheel barrow

Topics: Machines

Aim: To examine the properties of a wheel barrow as a second order lever.

Equipment

Wheel barrow
Two concrete blocks or 8
house bricks

Procedure

- 1/ Place the bricks in the barrow over the wheel .
- 2/ Lift the barrow by the end of the handles.
- 3/ On the diagram provided draw the position of the load with a downward arrow and the position of your hands with an upward arrow.
- 4/ On the line below the diagram record whether the lift is easy or hard.
- 5/ Lower the barrow and slide your hands along the handles to the bucket. Lift the barrow again.
- 6/ Record the load and effort positions on the next diagram and whether the lift is easier or harder.
- 7/ Move the bricks close to the handles.
- 8/ Lift with your hands close to the bucket.
- 9/ Repeat step 7.
- 10/ Lift by the end of the handles and repeat step 7.
- 11/ Compare the relative distance of the load and effort from the fulcrum in each diagram.

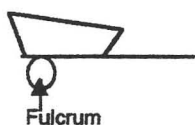
Situation 1



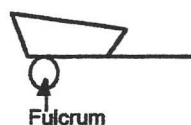
Situation 2



Situation 3



Situation 4



Result: Lifting is easiest in situation 1, harder in situation 2, harder again in situation 3 and then a bit easier in situation 4.

Conclusion: The least effort is required to lift the barrow when the distance of the effort to the fulcrum is much greater than the distance of the load to the fulcrum.

Risk Level: Low Hazard but supervision is necessary to stop showing off and to ensure a load is not spilled onto a student.

STUDENT: _____

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Electromagnets 1

Aim: To produce a magnetic field from current in a coil of wire.

Equipment

- 150mm Iron nail or bolt
- Insulated wire, 1m
- Power supply 2-12V DC
- Connecting wires, 2
- Steel paper clips
- Magnetic Compass
- Adhesive Tape

Procedure

- 1/ Wind the insulated wire around the iron nail in even, tight coils, from one end to the other and then back in a second layer.
- 2/ Use adhesive tape to bind the coil, leaving the two ends free.
- 3/ Use the connecting wires to link the positive terminal of the power supply to one free end of the wire coil and the negative terminal to the other free end.
- 4/ Set the power supply to 2V DC
- 5/ Place a compass near one end of the nail
- 6/ Turn on the power briefly and record the behaviour of the magnet.
- 7/ Swap the power supply connections and repeat step 6.
- 8/ Place the paper clips 1cm from the end of the nail, briefly turn on the power and record what happens.
- 9/ Carefully remove the iron nail from the coil and repeat the last step using the hollow coil instead

Results: _____

Conclusion: _____

Electromagnets 1

Topics: Magnetism Forces Electromagnetism

Aim: To produce a magnetic field from current in a coil of wire.

Equipment

150mm Iron nail or bolt
Insulated wire, 1m
Power supply 2-12V DC
Connecting wires, 2
Steel paper clips
Magnetic Compass
Adhesive Tape

Procedure

- 1/ Wind the insulated wire around the iron nail in even, tight coils, from one end to the other and then back in a second layer.
- 2/ Use adhesive tape to bind the coil, leaving the two ends free.
- 3/ Use the connecting wires to link the positive terminal of the power supply to one free end of the wire coil and the negative terminal to the other free end.
- 4/ Set the power supply to 2V DC
- 5/ Place a compass near one end of the nail
- 6/ Turn on the power briefly and record the behaviour of the magnet.
- 7/ Swap the power supply connections and repeat step 6.
- 8/ Place the paper clips 1cm from the end of the nail, briefly turn on the power and record what happens.
- 9/ Carefully remove the iron nail from the coil and repeat the last step using the hollow coil instead.

Result: The compass needle swung toward the electromagnet. When the current was reversed the compass needle swung to the opposite direction. When the nail was removed the electromagnet was much weaker.

Conclusion: A magnetic field is created when electric current flows in a coil of wire. The field is reversed if the current is reversed. The field is much stronger if the coil is wound around an iron core.

Risk Level: Low Hazard; Higher voltage settings will overload the power supplies. The units may automatically reset after a few seconds or internal fuses may need to be replaced.

Electromagnets 2

Aim: To investigate the variables affecting the magnetic field produced by a current in a coil of wire.

Equipment

- 150mm Iron nail or bolt
- Insulated wire, 1m
- Power supply 2-12V DC
- Connecting wires, 3
- Retort Stand, boss head
- Spring Balance, 5N
- Adhesive Tape

Procedure

- 1/ Wind the insulated wire around the iron nail in even, tight coils, from one end to the other and then back in a second layer. Bind the coil with tape, leaving the two ends free.
- 2/ Use the connecting wires to link the terminals of the power supply to one free ends of the wire coil.
- 3/ Tape a bar magnet to the hook of the spring balance (North Pole downward). Suspend the balance on the stand.
- 4/ Record the balance reading in the table provided.
- 5/ Turn on the power at 2 V DC and bring one pole of the electromagnet to 2cm below the bar magnet.
- 6/ Record the balance reading.
- 7/ Repeat the readings using 4V DC.
- 8/ Repeat the readings at 2V DC, power terminals reversed.
- 9/ Unwind the out layer of windings and repeat the reading.
- 10/ Draw a diagram of your electromagnet in the space at left, indicating the direction of current flow and direction of the magnetic field produced.

| Trial | Balance | Force |
|------------------------------------|---------|------------|
| | Reading | (Newtons) |
| Bar magnet alone | | 0 |
| Electromagnet at 2V | | |
| Electromagnet at 4V | | |
| Electromagnet at 2V reversed | | |
| Electromagnet at 2V, half windings | | |

Results: _____

Conclusion: _____

Electromagnets 2

Topics: Magnetism

Forces

Electromagnetism

Aim: To investigate the variables affecting the magnetic field produced by a current in a coil of wire.

Equipment

150mm Iron nail or bolt
Insulated wire, 1m
Power supply 2-12V DC
Connecting wires, 3
Retort Stand, boss head
Spring Balance, 5N
Adhesive Tape

Procedure

1/ Wind the insulated wire around the iron nail in even, tight coils, from one end to the other and then back in a second layer. Bind the coil with tape, leaving the two ends free.
2/ Use the connecting wires to link the terminals of the power supply to one free ends of the wire coil.
3/ Tape a bar magnet to the hook of the spring balance (North Pole downward). Suspend the balance on the stand.
4/ Record the balance reading in the table provided.
5/ Turn on the power at 2 V DC and bring one pole of the electromagnet to 2cm below the bar magnet.
6/ Record the balance reading.
7/ Repeat the readings using 4V DC.
8/ Repeat the readings at 2V DC, power terminals reversed.
9/ Unwind the out layer of windings and repeat the reading.
10/ Draw a diagram of your electromagnet in the space at left, indicating the direction of current flow and direction of the magnetic field produced.

| Trial | Balance Reading | Force (Newtons) |
|------------------------------------|--------------------|---------------------|
| Bar magnet alone | | 0 |
| Electromagnet at 2V | | |
| Electromagnet at 4V | | |
| Electromagnet at 2V reversed | | |
| Electromagnet at 2V, half windings | | |

Result: Increasing voltage increased the magnetic field produced. Reversing voltage reversed the magnetic field. Reducing the number of windings reduced the magnetic field produced.

Conclusion: A magnetic field is created when electric current flows in a coil of wire. The direction of the field is dependent on the direction of the current (right hand rule). The strength of the field is proportional to the number of coils and to the size of the current (current is proportional to applied voltage).

Risk Level: Low Hazard; Higher voltage settings will overload the power supplies. The units may automatically reset after a few seconds or internal fuses may need to be replaced.

Invisible Force

Aim: To determine whether invisible magnetic force lines can be blocked by various materials.

Equipment

- Spring Balance ,5N
- Bar Magnets, 2
- Adhesive tape
- Copper sheet
- glass slide
- Retort stand, boss head
- heat mat
- plastic sheet

Procedure

- 1/ Suspend the spring balance from the boss head on the retort stand.
- 2/ Tape a bar magnet to the balance hook, North Pole facing down.
- 3/ Record the reading on the balance in the table provided.
- 4/ Position the South Pole of the second magnet so it is 2cm below the suspended magnet and record the new balance reading.
- 5/ Place a sample of the following materials between the magnets, recording the balance reading each time:
 - copper sheet
 - glass slide
 - wooden ruler
 - pages of your book
 - plastic sheet
 - heat mat
- 6/ Complete the table

| Trial | Balance Reading | Force (Newtons) |
|----------------------------|--------------------|---------------------|
| One Magnet | | 0 |
| Both Magnets | | |
| Both Magnets, Copper sheet | | |
| Both Magnets, glass slide | | |
| Both Magnets, wooden ruler | | |
| Both Magnets, paper | | |
| Both Magnets, heat mat | | |
| Both Magnets, plastic | | |

Results: _____

Conclusion: _____

Invisible Force

Topics: Magnetism

Forces

Aim: To determine whether invisible magnetic force lines can be blocked by various materials.

Equipment

Spring Balance ,5N

Bar Magnets, 2

Adhesive tape

Copper sheet

glass slide

Retort stand, boss head

heat mat

plastic sheet

Procedure

1/ Suspend the spring balance from the boss head on the retort stand.

2/ Tape a bar magnet to the balance hook, North Pole facing down.

3/ Record the reading on the balance in the table provided.

4/ Position the South Pole of the second magnet so it is 2cm below the suspended magnet and record the new balance reading.

5/ Place a sample of the following materials between the magnets, recording the balance reading each time:

copper sheet

glass slide

wooden ruler

pages of your book

plastic sheet

heat mat

6/ Complete the table

| Trial | Balance Reading | Force (Newtons) |
|----------------------------|--------------------|---------------------|
| One Magnet | | 0 |
| Both Magnets | | |
| Both Magnets, Copper sheet | | |
| Both Magnets, glass slide | | |
| Both Magnets, wooden ruler | | |
| Both Magnets, paper | | |
| Both Magnets, heat mat | | |
| Both Magnets, plastic | | |

Result: Positioning the second magnet increased the balance reading but interposing materials between the magnets caused no change.

Conclusion: Invisible lines of magnetic force between the magnets could not be blocked with any of the materials tested.

Risk Level: Low Hazard

STUDENT: _____

285

Alloys

Aim: To make a metallic alloy.

Equipment

Tin granules
Lead Shot
Crucible
Pipe clay triangle
Tripod
Bunsen

Procedure

An alloy is formed when two metallic elements are mixed together to form a substance which has the combined properties of the original metals. No chemical reaction has taken place.

Silver coloured coins are an alloy of nickel and tin.

Brass is an alloy of copper and tin.

Golden coins are an alloy of nickel, copper and aluminium.

1/ Add a ten grains of lead shot to a crucible. Lead is a soft, grey metal.

2/ Add two granules of tin to the crucible. Tin is a hard, shiny metal.

3/ Place the crucible in a pipe clay triangle on top of a tripod.

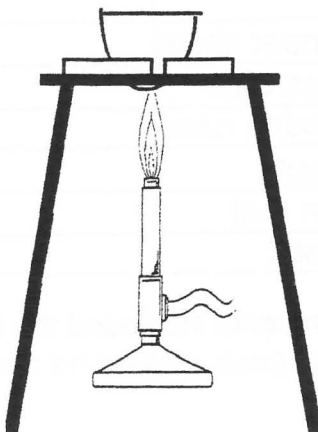
4/ Heat the crucible with tip of a blue bunsen flame until both metals have melted and mixed.

DO NOT BREATHE ANY FUMES FROM THE CRUCIBLE.

- Which metal melted first?

5/ Allow the crucible to cool for at least five minutes.

6/ Remove the disk of metal which has formed and describe its properties.



Results: _____

Conclusion: _____

Alloys

Topics: Matter Elements

Aim: To make a metallic alloy.

Equipment

Tin granules
Lead Shot
Crucible
Pipe clay triangle
Tripod
Bunsen

Procedure

An alloy is formed when two metallic elements are mixed together to form a substance which has the combined properties of the original metals. No chemical reaction has taken place.

Silver coloured coins are an alloy of nickel and tin.

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3/ Place the crucible in a pipe clay triangle on top of a tripod.

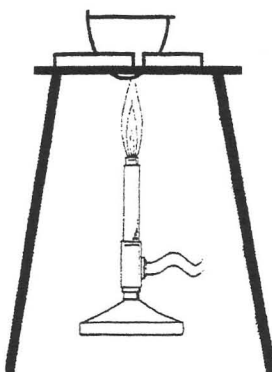
4/ Heat the crucible with tip of a blue Bunsen flame until both metals have melted and mixed.

DO NOT BREATH ANY FUMES FROM THE CRUCIBLE.

- Which metal melted first?

5/ Allow the crucible to cool for at least five minutes.

6/ Remove the disk of metal which has formed and describe its properties.



Result: The lead shot melted first, followed by the tin granules. The combined metals cooled to form a dull silver disk which was relatively easy to bend.

Conclusion: The lead and tin melted together to form an alloy with properties from both of the original metals.

Risk Level: Mild Hazard: PERFORM THIS EXPERIMENT IN A WELL VENTILATED ROOM.

STUDENT: _____

286

Indelible Chalk

Aim: To make chalk which won't rub off the board.

Equipment

Sugar
Beaker, 250ml
Stirring rod
Chalk
Paper hand towel

Procedure

- 1/ Add 200g of sugar to 200ml water, stirring until dissolved.
- 2/ Add several sticks of chalk.
- 3/ When the chalk ceases to release air bubbles, remove it from the water and allow to drain on the hand towel.
- 4/ Allow the chalk to dry for at least two days.
- 5/ Try the treated chalk on the black board.

Results: _____

Conclusion: _____

Indelible Chalk

Topics: Matter

Aim: To make chalk which won't rub off the board.

Equipment

Sugar
Beaker, 250ml
Stirring rod
Chalk
Paper hand towel

Procedure

- 1/ Add 200g of sugar to 200ml water, stirring until dissolved.
- 2/ Add several sticks of chalk.
- 3/ When the chalk ceases to release air bubbles, remove it from the water and allow to drain on the hand towel.
- 4/ Allow the chalk to dry for at least two days.
- 5/ Try the treated chalk on the black board.

For those inclined to practical jokes.

The chalk marks may be removed from the board with a wet cloth.



Result: The chalk writing on the board will not rub off with a duster, much like when you use the wrong pen on a white board.

Conclusion: The sugar forms crystals within the chalk. When drawn on the board the chalk has a higher adhesion strength and resists efforts to be removed with a duster.

Risk Level: Low Hazard: Except from the maths teacher when he finds out who doctored his chalk.

STUDENT: _____

287

Kinetic Corn

Aim: To demonstrate the Kinetic Theory of Matter.

Equipment

Pop Corn

Butter

Beakers, 250ml, two

Tripod

Bunsen

Evaporating basin, large

Fry pan

Preparation: Melt 50g of butter and pour over 100ml of pop corn kernels in an evaporating basin. Leave in a refrigerator until the butter hardens.

Procedure

1/ Watch the demonstration performed by the teacher, making notes on the properties of solids, liquids and gases and the behaviour of the particles in each of these states.

SOLID

Properties: _____

Particles are _____

LIQUIDS

Properties: _____

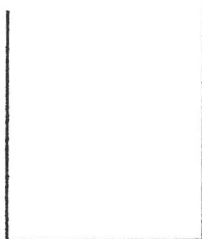
Particles are _____

GASES

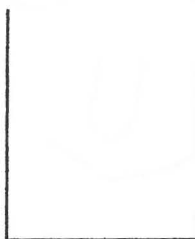
Properties: _____

Particles are _____

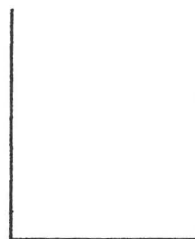
2/ In the containers drawn below draw particles in the solid, liquid and gaseous state.



Solid



Liquid



Gas

Results: _____

Conclusion: _____

Kinetic Corn

Topics: Matter Kinetic Theory

Aim: To demonstrate the Kinetic Theory of Matter.

Equipment

Pop Corn
Butter
Beakers, 250ml, two
Tripod
Bunsen
Evaporating basin, large
Fry pan

Preparation: Melt 50g of butter and pour over 100ml of pop corn kernels in an evaporating basin. Leave in a refrigerator until the butter hardens.

Good Idea to have a large pan ready to make a batch of pop corn for eating.

Procedure

1/ Remove the pop corn "solid" from the evaporating basin. Point out that the particles of corn are fixed in place and so the solid holds its shape and is incompressible.
2/ Heat the "solid" in the evaporating dish until the butter melts. Demonstrate that the particles are no longer fixed and the pop corn is now a "liquid" which can be poured to and from a beaker.
3/ Pour the pop corn "liquid" into a fry pan and continue heating. Students will note the behaviour of the corn particles as they pop forming a "gas".

Result: In solids, molecules are fixed in place by intermolecular bonds. In liquids, molecules are free to move past and around each other. In gases the molecules are freely flying apart.

Conclusion: The basic properties of solids, liquids and gases can be explained in terms of the motion of particles.

Risk Level: Mild Hazard: Beware of hot butter droplets as the corn pops (clear the bench first).

STUDENT: _____

288

Atomic Mass

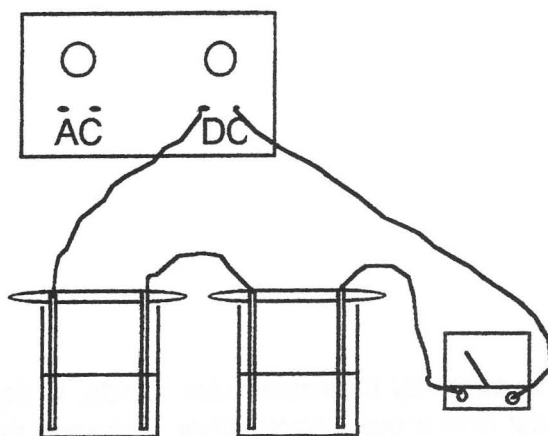
Aim: To determine the atomic mass of two elements.

Equipment

Power supply, 0-12V
Connecting wires, 4
Ammeter, 0-1A
Alligator clips, four
Tin electrodes, two
Copper electrodes, two
Balance, 1mg sensitivity
Copper Sulphate, 0.05M
(1.25%), 50ml
Tin Chloride, 0.05M, 50ml
Paddle pop sticks, two
Test tube peg
Steel wool

Procedure

- 1/ Clean the electrodes with steel wool.
- 2/ Carefully weigh one copper electrode. _____
- 3/ Fasten the electrode to a paddle pop stick with an alligator clip which is then linked to the positive DC power terminal.
- 4/ Clip the second copper electrode to the stick and place both electrodes in a beaker with 50ml of copper sulfate.
- 5/ Carefully weigh a tin electrode. _____
- 6/ Clip the tin electrode to a second stick. Connect the alligator clip to the second copper electrode.
- 7/ Clip a second tin electrode to the stick and place both electrodes in a beaker with 50ml of tin chloride.
- 8/ Connect the second tin electrode to the positive terminal of the ammeter and connect the negative terminal to the negative terminal of the power supply.
- 9/ Adjust the voltage until about 0.5Amps are flowing.
- 10/ After exactly 30mins, remove the first copper electrode, dry it 20cm above a bunsen and measure its weight. _____
- 10/ Repeat step 9 for the first tin electrode. _____



Atomic Mass of Copper

= Mass lost X current

0.00233

= _____

Atomic Mass of Tin

= Mass loss X current

0.00233

= _____

Results: _____

Conclusion: _____

Atomic Mass

Topics:

Molarity

Ions

Electricity

Aim: To determine the atomic mass of two elements.**Equipment**

Power supply, 0-12V

Connecting wires, 4

Ammeter, 0-1A

Alligator clips, four

Tin electrodes, two

Copper electrodes, two

Balance, 1mg sensitivity

Copper Sulphate, 0.05M

(1.25%), 50ml

Tin Chloride, 0.05M, 50ml

Paddle pop sticks, two

Test tube peg

Steel wool

Dissolve 11.3g of tin chloride in 50ml conc. hydrochloric acid then dilute to 1 litre.

Procedure

See the student copy for procedure. While the experiment proceeds take the students through the following calculation.

1 amp = 1 coulomb of charge per second.

In 30mins at 0.5 amps, $30 \times 60 \times 0.5$ coulombs of electrons have flowed ie 900 coulombs.

1 electron has a charge of 1.602×10^{-19} C

Therefore $900 \text{ C} = 900 / 1.602 \times 10^{-19} = 5.62 \times 10^{21}$ electrons

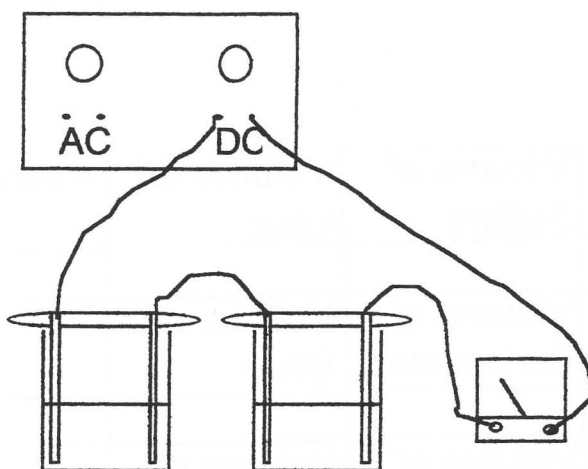
2 electrons are needed to ionise each metal atom so

2.81×10^{21} atoms have dissolved.

2.81×10^{21} atoms divided by Avogadro's number (6.02×10^{23}) gives 0.00466 moles.

Dividing the mass lost from each electrode by 0.00466 moles will give the mass per mole ie. Atomic Mass

(To adjust for a different current multiply by, current / 0.5)



Result: The copper anode lost about 0.3g weight while the tin electrode lost 0.6g.

Conclusion: Since the same number of electrons passed through both cells the same number of atoms must have dissolved into ions and so tin must have an atomic mass twice that of copper. The calculations of mass per mole give figures within 10% of the actual atomic mass, 63.5 for copper and 118.7 for tin.

Risk Level: Mild Hazard: 1M hydrochloric acid is mildly corrosive. Crystalline tin chloride may react violently with metal nitrates. Tin Chloride and copper sulfate may irritate the skin and are harmful if ingested.

STUDENT: _____

Valency

Aim: To investigate the valency (or combining power) of some elements.

Equipment

Test Tube Rack
Filter Funnel
Test tubes, large, four
Filter paper
Sodium Sulfate, 0.1M, 1.4%
Lead Nitrate, 0.1M, 3.3%
Silver Nitrate, 0.1M, 1.5 %
Potassium Carbonate, 0.1M, 1.4%
(use anhydrous salts)
Forceps (do not handle used filter papers)
Measuring cylinder, 10ml
(rinse thoroughly between uses).

Procedure

Atoms react in simple whole number ratios to form molecules. This ratio is called valency. If compounds of equal concentration are used, they too will react in simple ratios of volume.

1/ Add 10ml of sodium sulfate solution to a test tube and then add 5ml of lead nitrate solution.
2/ Filter the precipitate, collecting the filtrate into a clean tube.
3/ Add another 5ml of lead nitrate. If a strong precipitate develops, repeat steps 2 and 3.
4/ Add 5ml of potassium carbonate solution to a test tube and then add 5ml of silver nitrate.
5/ Filter the precipitate, collecting the filtrate into a clean tube.
6/ Add another 5ml of silver nitrate. If a strong precipitate develops repeat steps 5 and 6.
7/ Add 5ml of potassium carbonate solution to a test tube and then add 5ml of lead nitrate solution.
8/ Filter the precipitate, collecting the filtrate into a clean tube.
9/ Add another 5ml of lead nitrate. If a strong precipitate develops repeat steps 8 and 9.

| Reaction | Volume of Sulfate | Volume of Lead | Ratio |
|----------|---------------------|------------------|-------|
| 1 | | | |
| | Volume of Carbonate | Volume of Silver | |
| 2 | | | |
| | Volume of Carbonate | Volume of Lead | |
| 3 | | | |

Results: _____

Conclusion: _____

Valency

Topics: Molarity How Atoms Join Making Chemicals

Aim: To investigate the valency (or combining power) of some elements.

Equipment

Test Tube Rack
Filter Funnel
Test tubes, large, four
Filter paper
Sodium Sulfate, 0.1M, 1.4%
Lead Nitrate, 0.1M, 3.3%
Silver Nitrate, 0.1M, 1.5 %
Potassium Carbonate, 0.1M, 1.4%
(use anhydrous salts)
Forceps (do not handle used filter papers)
Measuring cylinder, 10ml (rinse thoroughly between uses).

Procedure

Atoms react in simple whole number ratios to form molecules. This ratio is called valency. If compounds of equal concentration are used, they too will react in simple ratios of volume.

1/ Add 10ml of sodium sulfate solution to a test tube and then add 5ml of lead nitrate solution.

2/ Filter the precipitate, collecting the filtrate into a clean tube.

3/ Add another 5ml of lead nitrate. If a strong precipitate develops, repeat steps 2 and 3.

4/ Add 5ml of potassium carbonate solution to a test tube and then add 5ml of silver nitrate.

5/ Filter the precipitate, collecting the filtrate into a clean tube.

6/ Add another 5ml of silver nitrate. If a strong precipitate develops repeat steps 5 and 6.

7/ Add 5ml of potassium carbonate solution to a test tube and then add 5ml of lead nitrate solution.

8/ Filter the precipitate, collecting the filtrate into a clean tube.

9/ Add another 5ml of lead nitrate. If a strong precipitate develops repeat steps 8 and 9.

| Reaction | Volume of Sulfate | Volume of Lead | Ratio |
|----------|---------------------|------------------|-------|
| 1 | 5 | 5 | 1:1 |
| | Volume of Carbonate | Volume of Silver | |
| 2 | 5 | 10 | 1:2 |
| | Volume of Carbonate | Volume of Lead | |
| 3 | 5 | 5 | 1:1 |

Result: Sodium sulfate and lead nitrate react in a 1:1 ratio. Potassium carbonate and silver nitrate react in a 1:2 ratio. Potassium carbonate and lead nitrate react in a 1:1 ratio.

Conclusion: Lead and sulfate have equal valency. Carbonate has twice the valency of silver but an equal valency to lead. Therefore silver has a valency of 1, lead a valency of 2, carbonate a valency of -2 and sulfate a valency of -2.

Risk Level: Moderate Hazard: Lead nitrate is a cumulative toxin and exposure should be minimised. Silver nitrate is toxic and may cause black staining of skin.

Elastic Collisions

Aim: To compare the elastic efficiency of several balls.

Equipment

- Tennis Ball
- Cricket Ball
- Baseball
- Golf Ball
- Squash Ball
- Metre Rule

Procedure

1/ Drop each ball from a height of one metre onto a concrete surface and record the height to which it rebounds.

Each ball possesses Potential Energy being equal to:
Mass X gravity X height.

As the ball falls, the potential energy is converted to kinetic energy. On impact the kinetic energy is stored as elastic potential energy then released on the rebound as kinetic energy again. The ball will rise until all its kinetic energy is converted back into potential energy. Since gravity and the mass of the ball are constant only the original height and the rebound height need be compared.

- Which ball best stores elastic potential energy ?
- Which ball has the poorest storage of elastic potential?
- Where does the energy go that is lost?

| Ball | Rebound Height |
|--------------|----------------|
| Cricket Ball | |
| Baseball | |
| Tennis Ball | |
| Golf Ball | |
| Squash Ball | |

Results: _____

Conclusion: _____

Elastic Collisions

Topics: Momentum

Energy

Aim: To compare the elastic efficiency of several balls.**Equipment**

Tennis Ball

Cricket Ball

Baseball

Golf Ball

Squash Ball

Metre Rule

Procedure

1/ Drop each ball from a height of one metre onto a concrete surface and record the height to which it rebounds.

Each ball possesses Potential Energy being equal to:

$\text{Mass} \times \text{gravity} \times \text{height}$.

As the ball falls, the potential energy is converted to kinetic energy. On impact the kinetic energy is stored as elastic potential energy then released on the rebound as kinetic energy again. The ball will rise until all its kinetic energy is converted back into potential energy. Since gravity and the mass of the ball are constant only the original height and the rebound height need be compared.

- Which ball best stores elastic potential energy ?
- Which ball has the poorest storage of elastic potential?
- Where does the energy go that is lost?

For a more advanced treatment of impact see Impulse 92, Volume 1

Result: The golf ball had the highest rebound while the squash ball had the lowest rebound.

Conclusion: The golf ball is best at storing elastic energy. The squash ball has the poorest elastic potential. The energy lost in the rebound is converted into heat and sound.

Risk Level: Low Hazard:

STUDENT: _____

291

Chain reaction

Aim: To demonstrate how a chain reaction occurs.

Equipment

Matches

Putty or play dough.

Procedure

Observe the demonstration carried out by the teacher.
Remember, the demonstration is a representation in two dimensions. Most reactions occur in three dimensions.

- What is a chain reaction?

- What is the critical factor in the demonstration which determines whether a chemical chain reaction takes place?

- Nuclear reactions are triggered by decay particles (eg neutrons) colliding with a nucleus. Even in solids where the atoms are closely packed, most of each atom is space and a decay particle may pass through many atoms before striking a nucleus. What factor would increase the likelihood of a decay neutron striking another nucleus?

Results: _____

Conclusion: _____

Chain reaction

Topics: Nuclear Physics

Aim: To demonstrate how a chain reaction occurs.

Equipment

Matches

Putty or play dough.

Procedure

IN A WELL VENTILATED ROOM

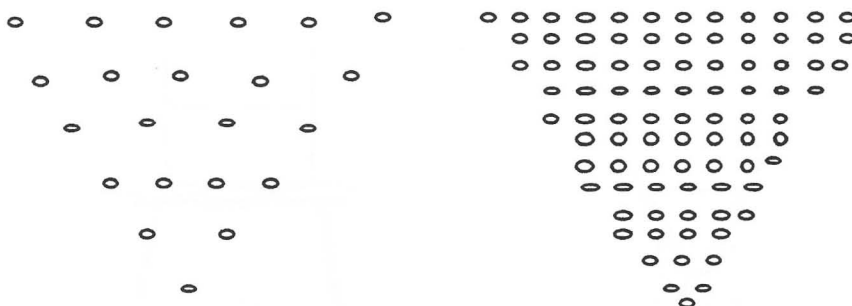
1/ Make two flat disks of putty about 8cm across.

2/ In the first disk, insert unused matches to make a triangular formation, each match head being about 1cm from its neighbour. Each side of the formation should be rank of 6 matches.

3/ In the second disk, insert unused matches to make a triangular formation, each match head being about 0.5cm from its neighbour. Each side of the formation should be rank of 12 matches.

4/ Light a match at the apex of the first formation.

5/ Light a match at the apex of the second formation.



Result: When matches are about 1cm apart the heat of one igniting is not sufficient to ignite a neighbour. When the matches are 0.5cm apart the first match ignites two which then ignite 4 and so on in an escalating chain reaction.

Conclusion: The decisive factor determining whether a chain reaction will take place is the density of reacting particles. Since much of an atom is space, decay particles can still escape small volumes of solid. In nuclear reactions there also has to be enough atoms so that decay particles are likely to strike another nucleus. This minimum number of atoms is called **Critical Mass**.

Risk Level: Moderate Hazard: TEACHER DEMONSTRATION ONLY. Phosphorus fumes from large numbers of burning matches is harmful if inhaled.



STUDENT: _____

292

Casein

Aim: To make a simple plastic.

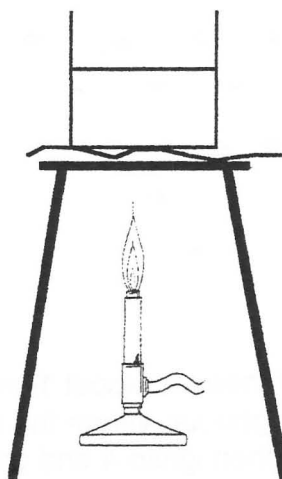
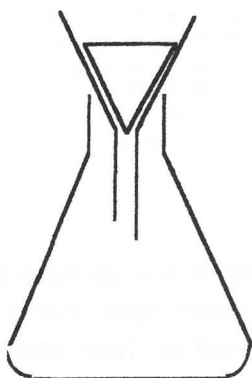
Equipment

Milk, 100ml
Vinegar
Beaker, 250ml
Conical flask, 250ml
Stirring Rod
Filter Funnel
Bunsen
Tripod
Gauze Square
Crepe Bandage, 10X10cm
Thermometer, 0-110C
Measuring cylinder, 10ml

Procedure

- 1/ Pour 100ml of milk into the beaker.
- 2/ Place the beaker on the gauze square and tripod.
- 3/ Heat with a bunsen, periodically measuring the temperature until the milk reaches 50 degrees centigrade.
- 4/ Slowly add 5ml of vinegar while stirring.
- 5/ Place the crepe material in the filter funnel and the funnel in the conical flask.
- 6/ Filter the milk. "Curds" will remain in the filter while the "Whey" passes through into conical flask.
- 7/ Remove the curds in the crepe and squeeze dry over a sink.
- 8/ Mold the curds into a shape and allow to dry for a few days.

- Your shape has now hardened into a plastic called casein. Is the shape flexible?
- Hold a tip of your shape in a bunsen flame. If the plastic melts it is called a thermoplastic. If the plastic hardens it is called a thermosetting plastic.



Results: _____

Conclusion: _____

Casein

Topics: Organic Chem

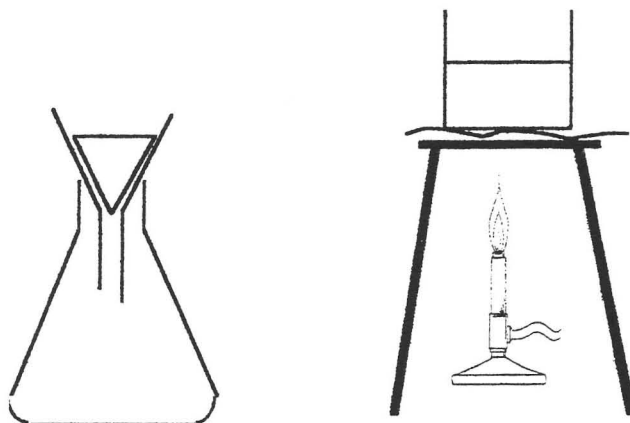
Aim: To make a simple plastic.

Equipment

Milk, 100ml
Vinegar
Beaker, 250ml
Conical flask, 250ml
Stirring Rod
Filter Funnel
Bunsen
Tripod
Gauze Square
Crepe Bandage, 10X10cm
Thermometer, 0-110C
Measuring cylinder, 10ml

Procedure

- 1/ Pour 100ml of milk into the beaker.
 - 2/ Place the beaker on the gauze square and tripod.
 - 3/ Heat with a Bunsen, periodically measuring the temperature until the milk reaches 50 degrees centigrade.
 - 4/ Slowly add 5ml of vinegar while stirring.
 - 5/ Place the crepe material in the filter funnel and the funnel in the conical flask.
 - 6/ Filter the milk. "Curds" will remain in the filter while the "Whey" passes through into conical flask.
 - 7/ Remove the curds in the crepe and squeeze dry over a sink.
 - 8/ Mould the curds into a shape and allow to dry for a few days.
- Your shape has now hardened into a plastic called casein. Is the shape flexible?
 - Hold a tip of your shape in a Bunsen flame. If the plastic melts it is called a thermoplastic. If the plastic hardens it is called a thermosetting plastic.



Result: The curds harden into a flexible, rubber like plastic which hardens when heated.

Conclusion: Casein is a thermosetting plastic. Unlike other plastics casein is not a true polymer (long chains of molecules). Milk contains large molecules called proteins. Heat and the acid in vinegar denature the proteins so they are no longer soluble in water. The protein molecules tangle and clump forming curds.

Risk Level: Low Hazard: Lots of fun making rubber balls.

STUDENT: _____

293

Nylon

Aim: To produce the polymer fibre, Nylon.

Equipment

Sebacoyl Chloride 2.2g
(or Adipoyl Chloride 1.7g)
1,6 Diaminohexane 2.2g
Dichloromethane, 100ml
Sodium Carbonate, 6g
Beakers, 250ml, two
Beaker, 400ml
Glass stirring rod
Forceps
Retort stand and clamp
Tripod.

Procedure

A solution of diaminohexane is poured over a solution of sebacoyl chloride to form a two layered liquid.

- Draw and label the apparatus used by the teacher.
- Describe how nylon is produced from the interface of the two liquids.
- How would you chemically describe Nylon?

Results: _____

Conclusion: _____

Nylon

Topics: Organic Chem

Aim: To produce the polymer fibre, Nylon.

Equipment

Sebacoyl Chloride 2.2g
(or Adipoyl Chloride 1.7g)
1,6 Diaminohexane 2.2g
Dichloromethane, 100ml
Sodium Carbonate, 6g
Beakers, 250ml, two
Beaker, 400ml
Glass stirring rod
Forceps
Retort stand and clamp
Tripod.

Procedure

Perform Steps 1 to 3 in a fume hood.

1/ Dissolve 2.2g sebacoyl chloride in 100ml dichloromethane with vigorous stirring in a 250ml beaker.

2/ Dissolve 2.2g diaminohexane and 6g sodium carbonate in 50ml of water.

3/ Carefully pour the diaminohexane down the side of the beaker containing the sebacoyl chloride solution to form a two layered solution.

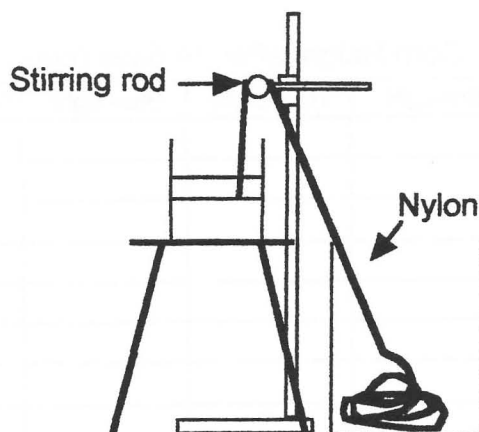
4/ Fix the glass rod horizontally on the retort stand about 30cm above the bench.

5/ Place the layered solution on the tripod positioned below the horizontal stirring rod.

6/ Place the 400ml beaker beside the tripod.

7/ Use forceps to grasp the nylon at the interface of the two layers and carefully draw it over the stirring rod and down into the large beaker.

The reagents in this experiment are expensive.



Result: A continuous strand of nylon is drawn from the interface and accumulates in the large beaker.

Conclusion: Nylon is a linear addition polymer formed from sebacoyl chloride and diaminohexane.

Risk Level: HAZARDOUS: TEACHER DEMONSTRATION ONLY. Sebacoyl chloride is corrosive and reacts with moisture in the air to form hydrogen chloride fumes (old bottles can explode from internal pressure). Diaminohexane is caustic and inflammable. Dichloromethane is flammable.

STUDENT: _____

Green Plants?

Aim: To investigate why plants are green rather than red or blue.

Equipment

Corn seeds (40)
Plastic punnets, four
Clear wrapping plastic
Tinted plastic, 30X40cm,
red, green and blue.

potting mix
beaker, 250ml
Sticky tape

Procedure

- 1/ Fill each plastic punnet with moist potting mix.
- 2/ Plant ten corn seeds in each punnet, 1cm deep and evenly dispersed.
- 3/ Pour 200ml of water slowly over each punnet.
- 4/ Wrap one punnet in a tube of clear wrapping plastic.
 Wrap one punnet in a tube of red tinted plastic.
 Wrap one punnet in a tube of blue tinted plastic.
 Wrap one punnet in a tube of green tinted plastic.
- 5/ Place the punnets in a sunny position, giving each punnet 100ml of water every two days.
- 6/ After 14 days examine the corn plants, measuring the height of each.

Corn Height after 14 days (cm)

| Plants | Sunlight | Red Light | Blue light | Green light |
|---------|----------|-----------|------------|-------------|
| 1 | | | | |
| 2 | | | | |
| 3 | | | | |
| 4 | | | | |
| 5 | | | | |
| 6 | | | | |
| 7 | | | | |
| 8 | | | | |
| 9 | | | | |
| 10 | | | | |
| Average | | | | |

Results: _____

Conclusion: _____

Green Plants?

Topics: Plants Scientific method

Aim: To investigate why plants are green rather than red or blue.

Equipment

Corn seeds (40)
Plastic punnets, four
Clear wrapping plastic
Tinted plastic, 30X40cm,
red, green and blue.

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| 6 | | | | |
| 7 | | | | |
| 8 | | | | |
| 9 | | | | |
| 10 | | | | |
| Average | | | | |

Result: Plants in the clear plastic tube were green and tall. Plants in the red tube and blue tube were less vigorous. Plants in the green tube were very pale and stunted.

Conclusion: Plants receiving a full sunlight spectrum grew well while plants receiving predominantly red or blue light grew, but not as well. Plants receiving green light did not grow well. Therefore the pigment in plants which absorbs and uses light for photosynthesis does not absorb green light. Plants appear green because they reflect green light while absorbing other frequencies.

Risk Level: Low Hazard:



STUDENT: _____

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Leaf Section

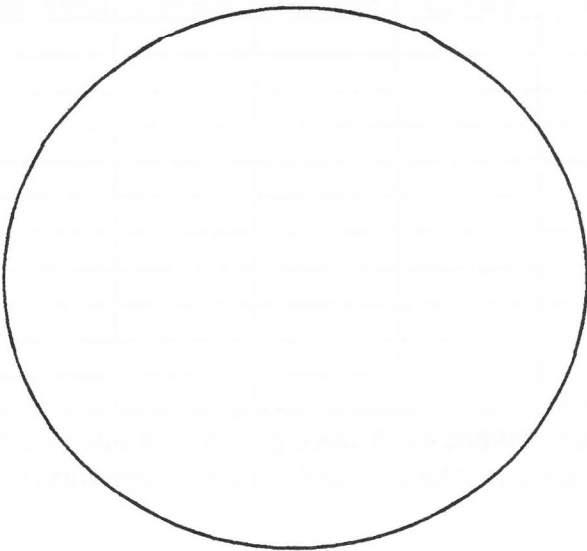
Aim: To prepare and examine a leaf cross-section under a microscope.

Equipment

Safety razor blades, two
Glass slide
coverslip
Dropper bottle of water
Microscope
Fresh plant leaf.

Procedure

1. Place the leaf over the glass slide.
2. Hold both razor blades together and with light pressure draw the corner of the blades across the leaf. A thin leaf section remains between the severed parts of the leaf.
3. Remove the main parts of the leaf.
4. Trim the leaf section to about 1cm in length.
5. Carefully turn the leaf section on its side with one razor.
6. Gently place a cover slip over the leaf section.
7. Place a droplet of water at the edge of the coverslip. The water will be draw around the leaf section by capilarity.
8. Examine the leaf section under the microscope at 100X magnification.
9. In the space below, carefully draw the internal structure of the leaf section.



Results: _____

Conclusion: _____

Leaf Section

Topics: Plants

Aim: To prepare and examine a leaf cross-section under a microscope.

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Glass slide
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Dropper bottle of water
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9. In the space below, carefully draw the internal structure of the leaf section.

Begonia or other indoor plants are best, being softer and thicker.

Result: Several different layers can be discerned within the leaf section.

Conclusion: Cuticle and palisade layers should be easily discerned.

Risk Level: Mild Hazard: Count and issue the safety razors just before the students are ready to commence. Collect and count the razors as soon as the sections have been prepared.

Leaf Section

Topics: Plants

Aim: To prepare and examine a leaf cross-section under a microscope.

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Glass slide
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STUDENT: _____

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Falling Target

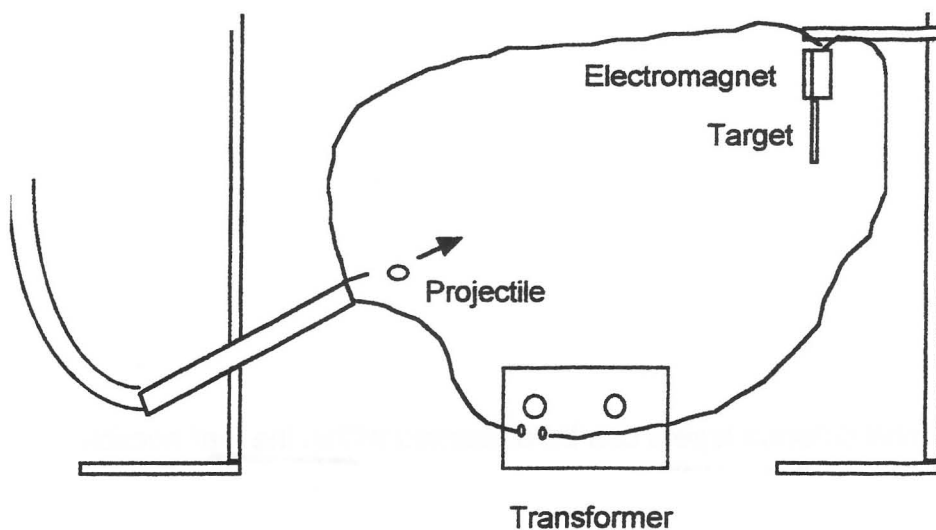
Aim: To demonstrate that objects fall at the same rate thereby eliminating the vertical component of motion.

Equipment

Retort Stands, 2
Clamps & Boss Heads, 2
Electromagnet
Power supply 2-12V
Connecting wires, 3
Steel Target plate, 5x5cm
Flexible tubing, 40cm
Rigid tube, 20cm
Marble projectile
Adhesive tape

Procedure

- 1/ Connect the flexible tube to the rigid tube
- 2/ Use the retort stands to set up the apparatus as shown in the diagram below.
- 3/ Use a connecting wire to link one terminal of the power supply to the electromagnet. Link the other magnet terminal to the mouth of the rigid tube with tape. Link the other power terminal to the mouth of the rigid tube. Link the two wires at the tube without blocking the tube mouth.
- 4/ Turn on the power to 6 Volts DC and suspend the target plate beneath the electromagnet.
- 5/ Place the marble in the rigid tube and aim at the target
- 6/ Aim the rigid tube at the target.
- 7/ Blow rapidly through the flexible tube
- 8/ Change the height of the rigid tube, aim and fire again.
- 9/ Place the tube level with the target, aim and fire again.
- 10/ Repeat steps 5 to 9 with the wires touching across the the tube mouth so the marble will break the contact .



Results: _____

Conclusion: _____

Falling Target

Topics: Projectiles

Gravity

Vectors

Aim: To demonstrate that objects fall at the same rate thereby eliminating the vertical component of motion.

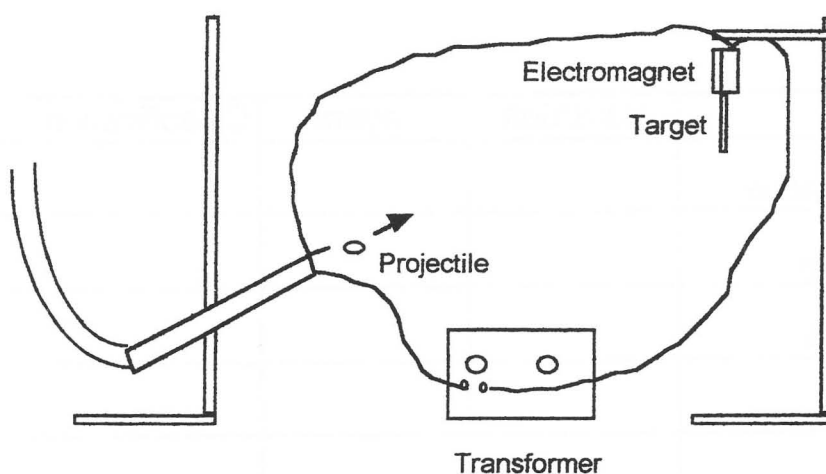
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Clamps & Boss Heads, 2
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Power supply 2-12V
Connecting wires, 3
Steel Target plate, 5x5cm
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Rigid tube, 20cm
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- 8/ Change the height of the rigid tube, aim and fire again.
- 9/ Place the tube level with the target, aim and fire again.
- 10/ Repeat steps 5 to 9 with the wires touching across the the tube mouth so the marble will break the contact .

Hint ; Check that the marble and the rig tube are a close fit.



Result: The marble follows a curved path missing the target held by the magnet. When the marble breaks the contact wires the target plate falls. The marble always hits the falling target regardless of the firing height or angle.

Conclusion: The marble misses the fixed target since its path is altered by gravity. When the marble breaks the wire contacts at the mouth of the tube the marble and the target fall with equal acceleration due to gravity. When the marble reaches the horizontal position of the target it will also be at the same height as the target and so they strike.

Risk Level: Low Hazard; Use a projectile which does not gather too much velocity or the students will be tempted to shoot each other ie. a marble is better than a dried pea.

Rock Types

Aim: To classify rock samples as either sedimentary, igneous or metamorphic.

Equipment

- Rock Samples:
Basalt
Shale
Sandstone
Schist
Granite
Gneiss

Procedure

Examine each rock sample.
Record in the table below whether the sample contains layers.
Record whether sample is hard or soft depending on if it is easily marked or grains removed by a finger nail.
Classify the samples as sedimentary, metamorphic or igneous based on the following data:
- Igneous rocks form from molten magma or lava and so are hard and lack straight layers.
- Sedimentary rocks are formed by weathered particles settling into layers which gradually compact. Sedimentary rocks have layers and are usually soft.
- Metamorphic rocks are sedimentary rocks which have been changed by extreme heat and pressure. These rocks often have thin compressed layers and though harder than the original rock, metamorphic rocks are not as hard as igneous rocks.

| Rock | Hard/Soft | Layers | Classification |
|-----------|-----------|--------|----------------|
| Sandstone | | | |
| Granite | | | |
| Gneiss | | | |
| Basalt | | | |
| Schist | | | |
| Shale | | | |

Results: _____

Conclusion: _____

Rock Types

Topics: Rocks & Minerals

Aim: To classify rock samples as either sedimentary, igneous or metamorphic.

Equipment

Rock Samples:

Basalt

Shale

Sandstone

Schist

Granite

Gneiss

Procedure

Examine each rock sample.

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| Rock | Hard/Soft | Layers | Classification |
|-----------|-----------|--------|----------------|
| Sandstone | | | |
| Granite | | | |
| Gneiss | | | |
| Basalt | | | |
| Schist | | | |
| Shale | | | |

Result: Shale and sandstone are soft with layers. Schist and gneiss have fine layers and are harder. Granite and basalt have no layers and are very hard.

Conclusion: Shale and Sandstone are sedimentary. Schist and gneiss are metamorphic rocks while basalt and granite are igneous.

Risk Level: Low Hazard.

STUDENT: _____

298

Weathering

Aim: To simulate weathering by heating and cooling.

Equipment

Rock Sample (granite)

Tongs

Bunsen

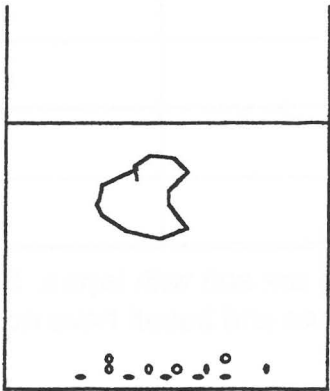
Beaker, 250ml

Protective Eye Glasses

Procedure

- 1/ Fill a beaker with water.
- 2/ Light the Bunsen and put on the eye glasses.
- 3/ Grasp the rock with the tongs and heat it over the Bunsen.
- 4/ Plunge the rock into the water.
- 5/ Repeat the last two steps several times.

- What is happening to the rock?
- What is happening to the fragments produced?



Results: _____

Conclusion: _____

Weathering

Topics: Rocks & Minerals

Aim: To simulate weathering by heating and cooling.

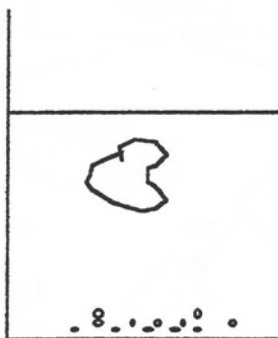
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Rock Sample (granite)
Tongs
Bunsen
Beaker, 250ml
Protective Eye Glasses

Procedure

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- 4/ Plunge the rock into the water.
- 5/ Repeat the last two steps several times.

- What is happening to the rock?
- What is happening to the fragments produced?



Result: Fragments of the rock break off in the water and settle to the bottom.

Conclusion: The hard granite rock expands during heating. When plunged into water the outer layers contract faster than the inner parts of the rock. As a result fragments of rock break off. The rock is being weathered by heating and cooling. The fragments are sedimenting in the water. Over time the sediments may build up and become compressed into sedimentary rock.

Risk Level: Moderate Hazard: Hot fragments of rock may fly off during heating and so eye protection is necessary.

STUDENT: _____

299

Mystery Solution

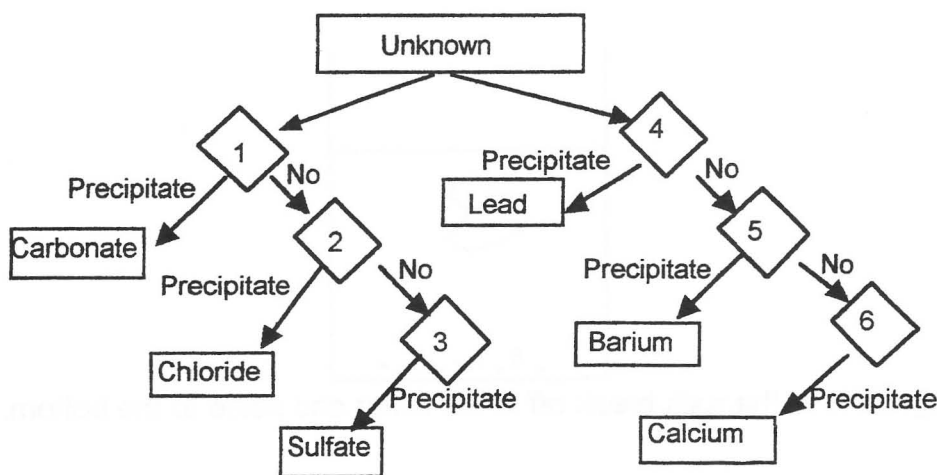
Aim: To identify the compound in unidentified solutions using a flow chart of tests.

Equipment

Test Tube rack
Test tube, medium, two
Labelled dropper bottles :
Copper Nitrate, 1%
Silver Nitrate, 1%
Lead Nitrate, 1%
Sodium Sulfate, 1%
Sodium Chloride, 10%
Sodium Hydroxide, 1%
Unknown 1
Unknown 2
Unknown 3

Procedure

1/ Add 20 drops of Unknown 1 to both test tubes
2/ Carry out the tests as indicated by the flow chart using the following key:
1/ Add 3 drops of copper nitrate.
2/ Add 3 drops of silver nitrate.
3/ Add 3 drops of lead nitrate.
4/ Add 3 drops of sodium chloride.
5/ Add 3 drops of sodium sulfate.
6/ Add 3 drops of sodium hydroxide.
3/ Having identified Unknown 1, thoroughly rinse the test tubes and repeat the procedure to identify Unknown 2.
4/ Thoroughly rinse the test tubes and repeat the procedure to identify unknown 3.



Results: _____

Conclusion: _____

Mystery Solution

Topics: Solubility

Scientific Method

Aim: To identify the compound in unidentified solutions using a flow chart of tests.

Equipment

Test Tube rack
 Test tube, medium, two
 Labelled dropper bottles :
 Copper Nitrate, 1%
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 Unknown 3

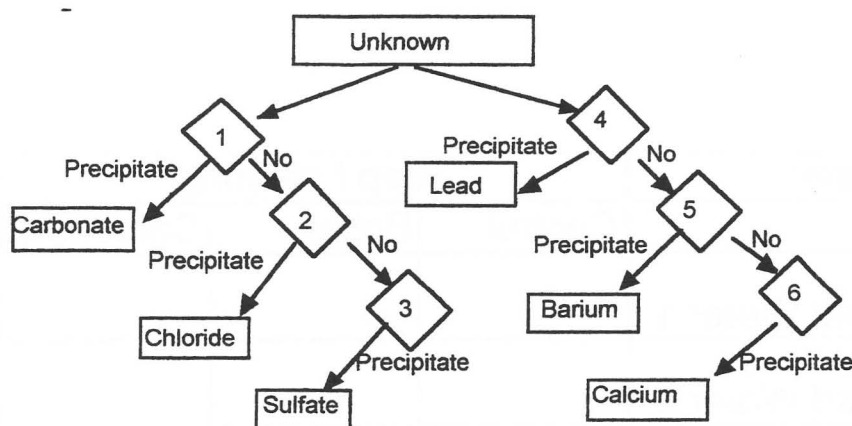
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3/ Having identified Unknown 1, thoroughly rinse the test tubes and repeat the procedure to identify Unknown 2.
 4/ Thoroughly rinse the test tubes and repeat the procedure to identify unknown 3.

- 1/ Calcium Chloride, 1%
- 2/ Barium Chloride, 1%
- 3/ Calcium Sulfate, 1%



Result: Unknown 1 produces a precipitate in reactions 2 and 6. Unknown 2 produces a precipitate in reactions 2 and 5. Unknown 3 produces a precipitate in reactions 3 and 6.

Conclusion: Unknown 1 is calcium chloride.
 Unknown 2 is barium chloride.
 Unknown 3 is calcium sulfate.

Risk Level: Moderate Hazard: Lead and silver nitrate are toxic and should be handled with care. Silver nitrate causes black skin stains while lead nitrate is a cumulative toxin. All precipitates of the lead or silver reagents should be considered harmful if ingested. Sodium hydroxide is caustic and any contact with the skin or eyes should be treated by prolonged washing with water.

STUDENT: _____

300

Softening Water

Aim: To investigate “Hard” water and how it may be softened.

Equipment

Test tubes, medium, three
Test tube stoppers, three
Test tube rack
Hard water 1(1.1g calcium chloride, 1.7g sodium bicarbonate, 1 litre water).
Hard water 2 (1.7g calcium sulfate, 1 L water)
Hard water 3 (1.1g calcium chloride, 1 L water)
Sodium Carbonate, 0.01M (0.1%) in dropper bottle.
Soap Solution (10g soap flakes dissolved in 1L water) in dropper bottle.
Measuring Cylinder, 10ml
Bunsen
Test tube peg.

Procedure

Water drawn from bores and wells often contains salts dissolved from rocks such as limestone. Calcium salts interfere with the lathering of soap. Water containing calcium is called “hard water”. Hard water may be softened by various methods depending on the type of calcium salt.

1/ Add 5ml of Hard Water 1 (Calcium Bicarbonate) to each test tube.
2/ Heat the second test tube until it boils gently for 1 minute.
3/ Add Sodium Carbonate dropwise to the third test tube until no more precipitate forms.
4/ Add ten drops of soap solution noting any change. Shake the the tube to see if a lather will form.
5/ Thoroughly rinse the test tubes.
6/ Repeat steps 1 to 5 for Hard Water 2 (calcium sulfate)
7/ Repeat steps 1 to 5 for Hard Water 3 (calcium chloride)

| Water | Soap Lathering | | |
|--------------|----------------|--------|-----------|
| | Control | Boiled | Carbonate |
| Hard Water 1 | | | |
| Hard Water 2 | | | |
| Hard Water 3 | | | |

Results: _____

Conclusion: _____

Softening Water

Topics:

Water

Ions

Organic Chem

Aim: To investigate "Hard" water and how it may be softened.**Equipment**

Test tubes, medium, three
 Test tube stoppers, three
 Test tube rack
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| Water | Soap Lathering | | |
|--------------|----------------|--------|-----------|
| | Control | Boiled | Carbonate |
| Hard Water 1 | | | |
| Hard Water 2 | | | |
| Hard Water 3 | | | |

Result: Boiling can soften water containing calcium bicarbonate. Sodium Carbonate precipitates calcium ions and softens all three types of hard water.

Conclusion: Boiling converts bicarbonate ions to carbonate ions causing precipitation of calcium carbonate. Sodium Carbonate is also known as washing soda since it softens hard water by precipitating calcium carbonate. Calcium ions can also be chelated in solution by complex ions such as EDTA. Gels containing anion exchange resins can also remove calcium ions.

Risk Level: Low Hazard:

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Reaction Rates

- 66 Exploding Bubbles
- 99 Iodate Clock
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- 8 Alcohols
- 20 Breathalyser
- 70 Food Tests
- 162 Soap
- 300 Softenning Water

- 62 Esters
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- 232 Cosmetics 2
- 233 Cosmetics 3

- 292 Casein

- 293 Nylon

- 151 Rayon

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- 141 Photosynthesis 2
- 158 Respiration 2
- 41 Colorimetry
- 264 Enzymes 1
- 265 Enzymes 2

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- 115 Mass of Air 2
- 72 Fountain Expt
- 170 Suction Fiction
- 184 Vacuum Boiling
- 133 Oxygen
- 134 Oxygen in Air
- 24 Carbon Dioxide
- 79 Green House Effect
- 89 Hydrogen Balloons
- 107 Liquid Air
- 112 Making Clouds
- 116 Measuring Clouds
- 224 Tin Canometer
- 209 Graphs 2
- 16 Bernoulli Effect
- 171 Sunset Expt.

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- 117 Measuring the Earth
- 179 The Seasons
- 106 Latitude
- 172 Sunspots
- 177 Tectonics
- 119 Metho Rockets
- 256 Angular Momentum

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- 46 Crystal Forms 1
- 49 Crystal Size
- 113 Making Rocks
- 298 Weathering
- 297 Rock Types

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- 49 Crystal Size
- 57 Earthquake Waves
- 177 Tectonics
- 190 Volcano

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- 175 Surface Tension
- 15 Bending Water
- 84 Heat & Temp
- 178 Temp vs Heat
- 270 Dissolving
- 185 Vacuum Boiling
- 300 Softenning Water
- 206 Controls
- 227 Quantitative Assay

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- 74 Fuses
- 87 Human Power
- 178 Temp versus Heat
- 250 Elastic Energy
- 246 Power
- 290 Elastic Collisions

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- 43 Convection
- 65 Expansion in Solids
- 91 Ice Cream
- 180 Thermocouples
- 83 Heat Absorbtion
- 84 Heat/Temp 1
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- 258 Magnetism
- 284 Invisible Force
- 257 Force Fields
- 296 Falling Target
- 200 "g" and an Air Track
- 194 Water Rocket
- 188 Video Expt 1
- 136 Pendulum 1
- 17 Big Lift
- 28 Centripetal Force
- 73 Friction
- 181 Ticker Timer 1
- 182 Ticker Timer 2
- 282 Electromagnets 1
- 283 Electromagnets 2
- 279 Inclined Plane
- 11 Attwoods Machine
- 71 Force Table
- 12 Ballistic Arrow
- 137 Pendulum 2
- 176 Suspension Bridge
- 187 Vector Tension
- 259 Tension and Acceleration
- 216 Residual Error

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- 181 Ticker Timer 1
- 182 Ticker Timer 2
- 73 Friction
- 148 Projectiles
- 188 Video Expt 1
- 296 Falling Target
- 12 Ballistic Arrow

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- 40 Collisions 1
- 92 Impulse
- 189 Video Expt 2
- 290 Elastic Collisions

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- 17 Big Lift
- 124 Moments
- 279 Inclined Plane
- 281 Wheel barrow
- 280 Pulleys
- 280 Pulleys
- 278 Gears and Torque
- 87 Human Power
- 246 Power

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- 72 Fountain Expt
- 22 Buoyancy
- 26 Cartesian Diver
- 69 Floating Iron
- 93 Instant Hydrometer
- 177 Tectonics
- 86 Hot air Balloon
- 114 Mass of Air 1
- 115 Mass of Air 2
- 89 Hydrogen Balloons
- 194 Water Rocket
- 240 Antigravity
- 170 Suction Fiction
- 173 Super Balloons
- 43 Convection
- 112 Making Clouds
- 184 Vacuum Boiling
- 260 Pressure Ignition

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- 16 Bernoulli Effect
- 255 Streamlining

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- 6 Aerodynamics
- 16 Bernoulli Effect
- 86 Hot air Balloon
- 207 Variables & Constants
- 254 Paper Planes
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- 248 Electron Attractions
- 15 Bending Water
- 185 Van de Graaf 1
- 186 Van De Graaf 2

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- 186 Van De Graaf 2
- 191 Water & Electrons
- 125 Morse Code
- 241 Fire Alarm
- 243 Series and Parallel
- 74 Fuses
- 13 Batteries 1
- 244 Shock Stack
- 14 Batteries 2
- 59 Electrolytic Plating
- 95 Internal Resistance
- 160 Seeing Ions
- 180 Thermocouples
- 242 Fuel Cell
- 198 Wheatstone Bridge
- 150 Radio Waves
- 48 Crystal Set
- 51 The Bug
- 245 Physics Prac
- 251 Power

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- 9 Alfoil Attractions
- 50 Current balance
- 60 Electron Beams
- 174 Super Induction
- 282 Electromagnets 1
- 283 Electromagnets 2
- 246 Power

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- 276 Flat Mirror
- 273 Concave Mirror
- 275 Convex Mirror
- 274 Convex Lens
- 272 Concave Lens
- 156 Refractive Index
- 155 Red is Black
- 94 Internal Reflection
- 127 Optical Illusions
- 237 Blind Spot
- 144 Polarisation
- 238 Depth Perception
- 139 Photochemical Reaction
- 83 Heat Absorbtion
- 48 Crystal Set
- 1 A Hairs Width

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- 80 Guitar Science
- 102 Kazoo
- 110 Long Springs
- 165 Sound Cannon
- 167 Speed of Sound
- 56 Dogs and Bats
- 161 Senses - Hearing
- 135 Particle Refraction
- 156 Refractive Index
- 195 Wave Tank
- 196 Waves in Strings
- 37 Coat Hanger Bell
- 57 Earthquake Waves
- 36 Closed Resonance Pipes
- 82 Harmonic Bunsen
- 144 Polarisation
- 150 Radio Waves
- 171 Sunset Expt.
- 197 Waves in Strings 2

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- 195 Wave Tank
- 197 Waves In Strings 2
- 94 Internal Reflection
- 135 Particle Refraction
- 104 Laser Diffraction
- 144 Polarisation
- 171 Sunset Expt.

1 A Hairs Width

Nuclear Physics

- 52 Discharge Tubes
- 60 Electron Beams
- 97 Invisible beams
- 149 Quantum Leaps
- 291 Chain reaction

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- 203 Procedure
- 235 Window Cleaners
- 234 Laundry Detergents
- 231 Cosmetics 1
- 232 Cosmetics 2
- 233 Cosmetics 3
- 266 Pasteurisation

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- 109 Living Fire
- 120 Microbes 2
- 126 Nematodes
- 7 Air Borne Microbes
- 277 Arthropod Art
- 212 Replicates
- 96 Invertebrates 2

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- 2 A Leaf is a Leaf
- 33 Chlorophyll Types
- 128 Osmosis
- 129 Osmosis 2
- 140 Photosynthesis 1
- 141 Photosynthesis 2
- 142 Plant Trophisms
- 145 Pollen Tubes
- 159 Seed Needs
- 169 Stomates
- 199 Xylem Tubes
- 294 Green Plants?
- 295 Leaf Section
- 217 Statistical Analysis
- 221 Plant Indicators

Inside Life

- 267 Rubber Bones
- 41 Colorimetry
- 140 Photosynthesis 1
- 141 Photosynthesis 2
- 157 Respiration
- 158 Respiration 2
- 8 Alcohols
- 266 Pasteurisation
- 162 Soap
- 263 Antisepsis
- 7 Air Borne Microbes

- 33 Chlorophyll Types
- 70 Food Tests
- 87 Human Power
- 264 Enzymes 1
- 265 Enzymes 2
- 42 Coloured Fire

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- 35 Chromosomes
- 145 Pollen Tubes
- 39 Code of Life
- 261 Genetic Differences
- 262 Zeibops
- 236 Forensic Science

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- 10 Are You Quick?
- 19 Body Language
- 55 Divers Response
- 56 Dogs and Bats
- 125 Morse Code
- 127 Optical Illusions
- 142 Plant Trophisms
- 154 Rebreathing
- 161 Senses - Hearing
- 183 Touch Sense
- 237 Blind Spot
- 238 Depth Perception
- 239 Learning
- 214 Blind Trial
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- 2 A Leaf is a Leaf
- 4 Adaptations 1
- 5 Adaptations 2
- 23 Camouflage
- 146 Pond Ecosystem
- 126 Nematodes

Evolution

- 23 Camouflage
- 254 Paper Planes
- 2 A Leaf is a Leaf
- 4 Adaptations 1
- 5 Adaptations 2

Scientific Method

- 201 A Scientists Eyes
- 202 Aims
- 203 Procedure
- 204 Results
- 205 Conclusion
- 206 Controls
- 207 Variables & Constants
- 208 Graphs 1
- 209 Graphs 2
- 210 Hypothesis
- 211 Parameters
- 212 Replicates
- 213 Sampling
- 214 Blind Trial
- 215 Double Blind Trial
- 216 Residual Error
- 217 Statistical Analysis
- 294 Green Plants?
- 234 Laundry Detergents

SCIENTIFIC METHOD

Scientific Method

201 A Scientists Eyes
202 Aims
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205 Conclusion
206 Controls
207 Variables & Constants
208 Graphs 1
209 Graphs 2
210 Hypothesis
211 Parameters
212 Replicates
213 Sampling
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215 Double Blind Trial
216 Residual Error
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234 Laundry Detergents
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299 Mystery Solution

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8 Alcohols
14 Batteries 2
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114 Mass of Air 1
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128 Osmosis
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134 Oxygen in Air
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174 Super Induction
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180 Thermocoples
199 Xylem Tubes
206 Controls
207 Variables & Constants
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264 Enzymes 1
265 Enzymes 2
266 Pasteurisation
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268 Electrolytes
269 Electrophoresis
282 Electromagnets 1
283 Electromagnets 2
284 Invisible Force
294 Green Plants?
300 Softenning Water

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- 201 A Scientists Eyes
- 202 Aims
- 205 Conclusion
- 206 Controls
- 208 Graphs 1
- 234 Laundry Detergents
- 203 Procedure
- 204 Results
- 207 Variables & Constants

4.3 Models- Atomic

- 53 Displacing Copper
- 270 Dissolving
- 118 Metho and Water

4.6.1 Energy

- 290 Elastic Collisions
- 250 Elastic Energy
- 136 Pendulum 1

4.6.2 Forces

- 257 Force Fields
- 258 Magnetism

4.6.3 Electrical Energy

- 74 Fuses
- 251 Power

4.6.4 Electrical Energy

- 294 Green Plants?
- 80 Guitar Science
- 82 Harmonic Bunsen
- 102 Kazoo
- 110 Long Springs

4.6.5 Spectrum

- 195 Wave Tank
- Colour Wheel

4.6.6 Heat Energy

- 43 Convection
- 65 Expansion in Solids
- 83 Heat Absorption

4.6.7 Friction

- 73 Friction

4.6.8 Electrostatics

- 15 Bending Water
- 247 Electric Wind
- 248 Electron Attractions
- 185 Van de Graaf 1
- 186 Van De Graaf 2

4.6.9 Magnetism

- 284 Invisible Force
- 258 Magnetism

4.7.1 Kinetic Theory

- 75 Gas Diffusion 1
- 76 Gas Diffusion 2
- 84 Heat/Temp 1
- 108 Liquid Diffusion
- 180 Thermocouples

4.7.2 Matter/Density/Pressure

- 6 Aerodynamics
- 16 Bernoulli Effect
- 26 Cartesian Diver
- 45 Crystal Forms
- 69 Floating Iron
- 72 Fountain Expt
- 85 Heat/Temp 2
- 86 Hot air Balloon
- 89 Hydrogen Balloons
- 93 Instant Hydrometer

- 255 Streamlining

- 170 Suction Fiction

- 173 Super Balloons

- 178 Temp versus Heat

- 194 Water Rocket

4.7.3 Changes of State

- 268 Electrolytes
- 68 Fire without Burning
- 271 Hot Cooling
- 91 Ice Cream
- 287 Kinetic Corn
- 105 Latent Heat
- 107 Liquid Air
- 168 States of Iodine

4.7.4 Elements

- 285 Alloys
- 33 Chlorophyll Types
- 58 Electrolysis
- 88 Hydrogen
- 133 Oxygen

4.7.5 Mixtures

- 3 Adsorbtion
- 34 Chromatography
- 231 Cosmetics 1
- 232 Cosmetics 2
- 233 Cosmetics 3
- 47 Crystal Garden
- 54 DISTILLATION
- 67 Filtration
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- 32 Chlorine
- 230 Consumer Science
- 61 Empirical Formula
- 98 Invisible Ink
- 101 Iron Sulfide
- 4.8.1 Cells**
- 120 Microbes
- 140 Photosynthesis 1

4.8.2 Living Things

- 2 A Leaf is a Leaf
- 277 Arthropod Art
- 126 Nematodes
- 109 Living Fire

4.8.3 Plant Tissues

- 295 Leaf Section
- 128 Osmosis
- 129 Osmosis 2
- 159 Seed Needs
- 169 Stomates
- 199 Xylem Tubes

4.8.4 Diversity

- 96 Invertebrates 2
- 157 Respiration

4.8.5 Body Systems

- 42 Coloured Fire
- 264 Enzymes 1
- 70 Food Tests
- 87 Human Power
- 267 Rubber Bones

4.9.1 Earth and Planets

- 106 Latitude
- 117 Measuring the Earth
- 172 Sunspots
- 179 The Seasons
- 256 Angular Momentum

4.9.4 Atmosphere

- 79 Green House Effect
- 112 Making Clouds
- 116 Measuring Clouds
- 134 Oxygen in Air
- 224 Tin Canometer

4.9.5 Hydrosphere

- 300 Softening Water
- 184 Vacuum Boiling
- 90 Hydrophilic/ phobic
- 175 Surface Tension Boats

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4.9.6 Rocks & Minerals

- 46 Crystal Forms 1
- 49 Crystal Size
- 113 Making Rocks
- 297 Rock Types
- 298 Weathering

5.2 Scientific method

- 214 Blind Trial
- 215 Double Blind Trial
- 209 Graphs 2
- 210 Hypothesis
- 299 Mystery Solution
- 211 Parameters
- 212 Replicates
- 213 Sampling
- 235 Window Cleaners

5.6.1 Waves

- 37 Coat Hanger Bell
- 161 Senses - Hearing
- 165 Sound Cannon
- 167 Speed of Sound
- 196 Waves In Strings

5.6.2 Motion

- 244 Series and Parallel
- 181 Ticker Timer 1
- 188 Video Expt 1
- 191 Water & Electrons

5.6.3 Circuits

- 48 Crystal Set
- 95 Internal Resistance
- 125 Morse Code
- 241 Fire Alarm
- 246 Power
- 282 Electromagnets 1

5.6.4 Light

- 272 Concave Lens
- 273 Concave Mirror
- 274 Convex Lens
- 275 Convex Mirror
- 276 Flat Mirror
- 155 Red is Black
- 156 Refractive Index

5.6.5 Nuclear Physics

- 291 Chain reaction
- 97 Invisible beams
- 149 Quantum Leaps

5.6.6 Mass and Weight

- 240 Antigravity
- 22 Buoyancy
- 114 Mass of Air 1
- 115 Mass of Air 2
- 89 Hydrogen balloons

5.7.1 Atomic Theory

- 52 Discharge Tubes
- 60 Electron Beams
- 78 Green Fire

5.7.2 Periodic Table

- 100 Ions
- 132 Oxides/pH
- 163 Solubilities
- 219 Amphoteric salts
- 220 Metallic Order
- 288 Atomic Mass
- 289 Valency

5.7.3 Chemical Rns

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- 249 Corrosion
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- 160 Seeing Ions
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- 166 Spectrum Clock
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- 222 Titration 1
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- 13 Batteries 1
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5.8.1 Cells 2

- 35 Chromosomes
- 141 Photosynthesis 2

5.8.2 Genetics

- 39 Code of Life
- 236 Forensic Science
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- 262 Zeibops

5.8.3 Evolution

- 4 Adaptations 1
- 23 Camouflage
- 254 Paper planes

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- 7 Air Borne Microbes
- 263 Antisepsis
- 10 Are You Quick?
- 237 Blind Spot
- 19 Body Language
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- 120 Microbes 2
- 127 Optical Illusions
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- 142 Plant Tropisms
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- 183 Touch Sense

5.9.1 Astronomy

- 150 Radio Waves

5.9.2 Tectonics

- 57 Earthquake Waves
- 177 Tectonics
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5.10 Ecology

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- 146 Pond Ecosystem

5.11.2 Pollution

- 228 Quantitative Assay
- 79 Greenhouse Effect

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- 17 Big Lift
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- 124 Moments
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- 281 Wheel barrow

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| 9 Afoil Attractions | 21 Buffers | 254 Paper Planes |
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| 242 Fuel Cell | 44 Copper Complexes | 33 Chlorophyll types |
| 92 Impulse | 269 Electrophoresis | 35 Chromosomes |
| 279 Inclined Plane | 265 Enzymes 2 | 39 Code of life |
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| 104 Laser Diffraction | 62 Esters | 142 Plant tropisms |
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| 137 Pendulum 2 | 81 Halogen Ions | 213 Sampling |
| 243 Physics Prac | 227 Ignition | 214 Blind trial |
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| 148 Projectiles 1 | 123 Molecular Weight | |
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| 171 Sunset Expt. | 130 Oxidation & Reduction | |
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